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Ion transfer at polarised liquid-liquid interfaces modified with adsorbed silica nanoparticles

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Abstract

The interface between two immiscible electrolyte solutions (ITIES) can act as a scaffold for the assembly of nanometer-sized objects. Here we followed the adsorption of silica nanoparticles of 12 nm diameter at the ITIES by AC voltammetry and their interactions with methylene blue (MB^+), selected as a model ion, by cyclic voltammetry and UV-vis absorption spectroscopy. We determined the association constant, K_{ads}^{aq} , of adsorption of MB^+ onto silica nanoparticles to be $1.66 \cdot 10^5$, which indicated a strong affinity between them. This strong affinity shifted the Gibbs free energy of transfer by -8.9 kJ mol^{-1} . This is in contrast with the other two ions investigated (Eosin B and tetraethylammonium), which demonstrated low affinity for the silica nanoparticles. By combining the ability of silica to adsorb onto the ITIES and their affinity for MB^+ , we were able to accumulate MB^+ at the ITIES.

1. Introduction

The interface formed between two immiscible electrolyte solutions (ITIES) is a highly reproducible, defect-free surface that allows the assembly of a large variety of nanoscale objects [1–4]. The adsorption of nano-objects at a liquid-liquid interface is governed by their dimensions, the organic-aqueous interfacial tension and the wettability of the material [2,5–8]. The ITIES allows the application of an external electric field that can drive the adsorption of nanoparticles (Au or TiO₂) to the interface [9–11]. Carbon nanomaterials (e.g. carbon nanotubes, graphene and graphite flakes) [12–15] and gold nanoparticles [16,17] adsorbed at the ITIES improve electron transfer kinetics between electron donors and acceptors distributed on either sides of the interface [14–17]. The nano-objects assembled at the interface act as conductive rafts floating at the interface catalysing electron transfer through Fermi level equilibration [18,19]. Gold nanoparticles have also been injected onto the interface between a mixture of heptane and 1,2-dichloroethane and water to form a self-assembled monolayer, which exhibits reflectance and electrical conductance, opening the way to the development of electrovariable optical materials [20]. Despite the abundant literature on silica nanoparticles [7,21–25], their assembly at the ITIES has not been investigated yet. Nevertheless, silica sphere structures have been self-assembled at non-polarised liquid-liquid interfaces to fabricate 2 or 3D ordered structures with potential applications in optical, magnetic and electronic devices as photonic band gap materials [26]. Li et al modified the surfaces of silica spheres and then self-assembled them at the liquid-liquid interface in order to fabricate 2D/3D film structures [26]. The ability to generate low-dimensional nanostructures is by far the most popular application of self-assembled silica nanoparticles at non-polarised liquid-liquid interfaces. The film obtained at the interface after spontaneous adsorption of the silica nanoparticles has low resistance to lateral motion and so is transferrable, allowing it to be used for other applications [27]. The growing interest for the use of silica-based materials at the liquid-liquid interface is also due to their favourable adsorption kinetics and selectivity [28], which could be harnessed to the development of new sensing methods [29,30]. ITIES was modified by the electrogeneration of mesoporous silica deposits [31–35], which were the result of an electrochemically driven ion transfer followed by interfacial precipitation. Mesoporous silica was generated at the ITIES by the transfer of a cationic surfactant (e.g. trimethylhexadecylammonium) from the organic phase to the aqueous phase containing hydrolysed tetra(alkoxy)silanes, catalysing the condensation reactions of the sol-gel process. Micro-ITIES arrays were also modified with mesoporous silica channels prepared ex-situ by a Stöber-like method [36]. The vertically aligned channels obtained have demonstrated their

potential for analytical applications [37–40]. We investigated here the adsorption of silica nanoparticles onto the ITIES and their impact on the electrochemical transfer behaviour of ions (e.g. methylene blue, eosin B and tetraethylammonium) in the view of developing new analytical methods. Since silica nanoparticles themselves do not have any electrochemical signal at the ITIES they can serve as an aid to enhance the selectivity and preconcentration of target ionic species at the interface and thus their transfer across the ITIES. This ability to detect ionic species is of major benefit in both environmental monitoring and clinical diagnosis. The sensing of synthetic organic molecules, such as drug molecules and pollutants, using silica nanoparticles adsorbed at the ITIES to improve sensitivity is of significant interest.

2. Experimental section

2.1. Chemicals & Reagents

The high density organic phase electrolyte was 1,2-dichloroethane (1,2-DCE, $\geq 99\%$ Sigma-Aldrich) containing 10 mM Bis (triphenylphosphoranyldiene) ammonium tetrakis (4-chlorophenylborate) (BTPPA⁺ TPBCl⁻). The organic salt was prepared by metathesis of Bis (triphenylphosphoranyldiene) ammonium chloride (BTPPA⁺ Cl⁻, 97% Aldrich) and Potassium tetrakis (4-chlorophenylborate) (K⁺ TPBCl⁻, $\geq 98\%$ Fluka) [31]. Lithium Chloride (LiCl, $\geq 99\%$ Aldrich) was used as the aqueous phase electrolyte. Solutions of the interfacial active ions: Tetraethylammonium chloride (TEA⁺ Cl⁻, $\geq 99\%$ Fluka), Methylene Blue ($> 82\%$ Fluka) and Eosin B ($> 85\%$ Sigma-Aldrich), were prepared in the aqueous phase electrolyte. The Ludox® HS-40 colloidal silica ($\varnothing_{mean} = 12$ nm, $10.6 < \varnothing < 13.7$ nm) was purchased from Aldrich. Ferric chloride hexahydrate (FeCl₃.6H₂O, $\geq 99\%$ Fluka) was used to prepare silver/silver chloride reference electrodes.

2.2. Experimental methods

The custom-made electrochemical cell ($A_{interface} = 1.13$ cm²) consisted of 2 sets of platinum mesh and silver/silver chloride wires as counter and reference electrodes respectively for both aqueous and organic phases. Different electrochemical cells were used and their compositions were as follows:

Electrochemical cell 1:

$\text{Ag}_{(s)} | \text{AgCl}_{(s)} | 10 \text{ mM LiCl}_{(aq)} + X \text{ g L}^{-1} \text{ silica nanoparticles} || 10 \text{ mM BTPPA}^+\text{TPBCl}^-_{(org)}$ in 1,2-DCE | Saturated $\text{BTPPA}^+\text{Cl}^-$ in 10 mM LiCl in $\text{H}_2\text{O} | \text{AgCl}_{(s)} | \text{Ag}_{(s)}$;

Electrochemical cell 2:

$\text{Ag}_{(s)} | \text{AgCl}_{(s)} | 100 \text{ mM LiCl}_{(aq)} + X \text{ g L}^{-1} \text{ silica nanoparticles} || 10 \text{ mM BTPPA}^+\text{TPBCl}^-_{(org)}$ in 1,2-DCE | Saturated $\text{BTPPA}^+\text{Cl}^-$ in 10 mM LiCl in $\text{H}_2\text{O} | \text{AgCl}_{(s)} | \text{Ag}_{(s)}$;

Electrochemical cell 3:

$\text{Ag}_{(s)} | \text{AgCl}_{(s)} | 10 \text{ mM HCl}_{(aq)} + X \text{ g L}^{-1} \text{ silica nanoparticles} || 10 \text{ mM BTPPA}^+\text{TPBCl}^-_{(org)}$ in 1,2-DCE | Saturated $\text{BTPPA}^+\text{Cl}^-$ in 10 mM LiCl in $\text{H}_2\text{O} | \text{AgCl}_{(s)} | \text{Ag}_{(s)}$;

Electrochemical cell 4:

$\text{Ag}_{(s)} | \text{AgCl}_{(s)} | \text{Saturated BTPPA}^+\text{Cl}^-$ in 10 mM LiCl in $\text{H}_2\text{O} | 10\text{mM BTPPA}^+\text{TPBCl}^-_{(org)}$ in 40:60 1,2-DCE:5-nonanone (v/v) || 10mM $\text{LiCl}_{(aq)} + X \text{ g L}^{-1} \text{ silica nanoparticles} | \text{AgCl}_{(s)} | \text{Ag}_{(s)}$.

Two electrochemical techniques were used during these experiments – Alternating Current (AC) Voltammetry and Cyclic Voltammetry (CV). Cyclic voltammetry experiments were run with an Autolab 302N potentiostat using Nova 1.10.1.9 software by Metrohm Autolab B.V. The formal transfer potential of TEA^+ , $\Delta^w\phi_{\text{TEA}^+}^0 = 0.049 \text{ V}$ [41], was used to calibrate the potential window. TEA^+ was added to the aqueous phase of the electrochemical cell at the end of a series of experiments for potential calibration purposes. For CV experiments, the transfers of Li^+ and Cl^- from the aqueous phase to the DCE phase establish the positive and negative limits of the potential window, respectively. In the presence of SiO_2 , the potential was calibrated using the limits of the potential window as previous studies have shown that the transfer of both Li^+ and Cl^- is not hindered by the presence of electrogenerated silica [33]. Scans were run at a scan rate of 5 mVs^{-1} . AC voltammetry experiments were run using GPES software with an Autolab PGSTAT-12. Parameters for the AC experiments were as follows - modulation time: 0.33 s, interval time: 0.73 s, frequency: 6 Hz, step potential: 5 mV, amplitude: 5 mV RMS. The potential region was scanned at phase angles of 0° and 90° . In this frequency range, the impedance response of the cell is determined by a single RC component associated with the uncompensated resistance and the double-layer capacitance [10,42].

UV/Vis absorption spectroscopy experiments were carried out using an Agilent Technologies Cary 60 UV-Vis spectrometer. Solutions analysed were prepared in 10 mM LiCl and poured in standard plastic cuvettes with a path length of 1 cm. For the determination of the association constant and the total number of adsorption sites by UV-vis absorption spectroscopy, Solutions

of 50 μM of MB^+ in 10 mM LiCl with varying concentrations of SiO_2 nanoparticles (final silica nanoparticle concentrations were between 0.040 and 0.867 g L^{-1} , total volume was 10 mL) were poured in a stirred ultrafiltration cell (25 mm diameter, Amicon Model 8010, Millipore, Carrigtwohill, Ireland) equipped with a regenerated cellulose membrane (PLGC, 10 kDa, Millipore). The filtrate was collected and analysed by UV-vis absorption spectroscopy. Pre-concentration experiments were prepared with 200 mL of a 25 μM MB^+ solution containing 0.05 g L^{-1} of SiO_2 nanoparticles. The solutions were mixed for 5 min and then centrifuged at 10,000 rpm for 15 min. The supernatant was removed and 3 mL of a 10 mM LiCl solution was added to the nanoparticle- MB^+ precipitate, which is then used as the aqueous phase in the electrochemical cell.

3. Results and discussion

3.1 Silica nanoparticles adsorption at the ITIES

Previous studies on the behaviour of nanoparticles at the liquid-liquid interface have shown that their adsorption is the result of the presence of a promoter or a modifier that weakens the electrostatic repulsion between the nanoparticles in suspension and allows the assembly of nanoparticles at the interface [43]. An alternative to this chemically induced assembly is the application of an interfacial potential difference that causes the assembly [9]. The adsorption of silica nanoparticles was investigated here by both cyclic and AC voltammetries at the ITIES (Figure 1). After the addition of 0.5 g L^{-1} silica nanoparticles to the aqueous phase, there was no significant changes to the CV, suggesting that no ion transfer is taking place in the potential window due the presence of the silica nanoparticles.

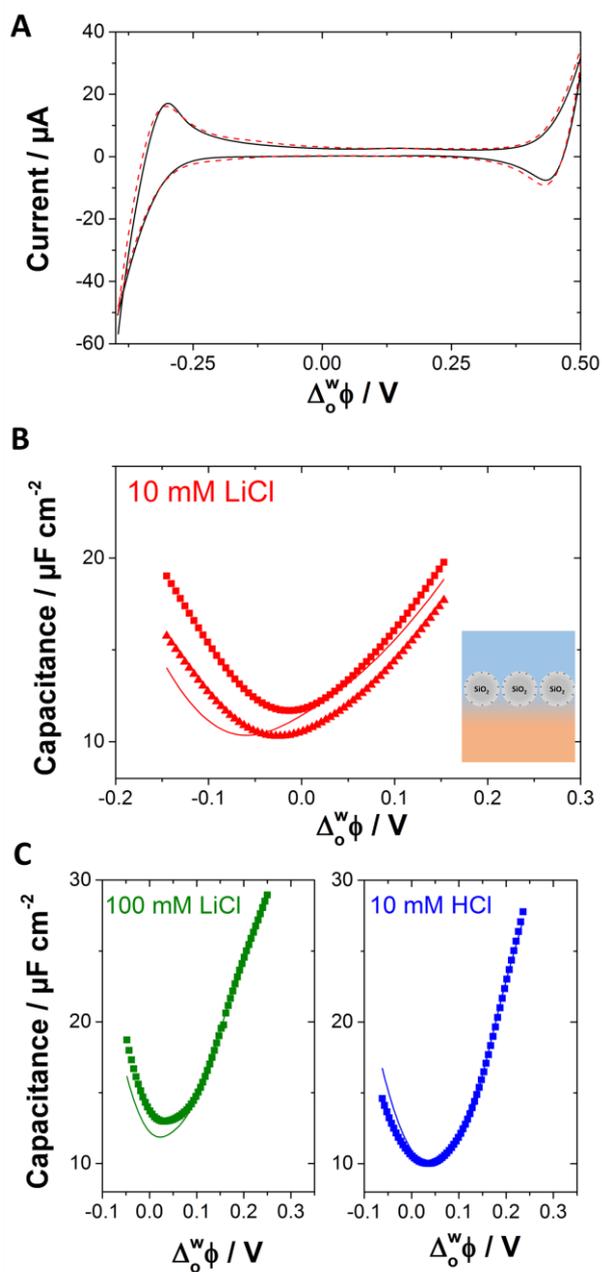


Figure 1: (A) Cyclic voltammety in the absence (black solid line) and in the presence (red dashed line) of 0.5 g L⁻¹ of silica nanoparticles in the aqueous phase. Electrochemical cell 1, $v = 5 \text{ mV s}^{-1}$; (B) Potential dependence of the interfacial capacitance for the blank cell (red solid line) and for (▲) 0.05 and (■) 0.1 g L⁻¹ of silica nanoparticles. Electrochemical cell 1. (C) Potential dependence of the interfacial capacitance for the blank cell (solid line) and for (■) 0.1 g L⁻¹ of silica nanoparticles with 100 mM LiCl (Electrochemical cell 2) or 10 mM HCl (Electrochemical cell 3) as an aqueous electrolyte solution.

Silica nanoparticles did not transfer across the interface and remained on the aqueous side of the interface throughout the potential window. AC voltammetry is more sensitive to interfacial phenomena than cyclic voltammetry and was then further used for the characterisation of the interface modified with silica nanoparticles. Even with quantities of silica nanoparticles as low as 0.05 g L^{-1} , significant changes were visible on the AC voltammograms recorded in the potential region where no background salt ion transfer was observed in cyclic voltammetry (Figure 1B). The differential capacitance was calculated from the AC voltammograms (see Experimental section). The potential of zero charge was taken as the potential of minimum of capacitance. In the absence of silica nanoparticles, it was measured at -0.06 V . Upon addition of silica nanoparticles to the aqueous phase, the capacitance value at -0.145 V increased from $14 (0 \text{ g L}^{-1})$ to $19 \mu\text{F cm}^{-2} (0.1 \text{ g L}^{-1})$ indicating that the quantity of charge located at the interface grew in the presence of silica nanoparticles. This behaviour can be attributed to the adsorption of negatively charged silica nanoparticles accumulating at the ITIES. This result is in agreement with that previously obtained for the interfacial assembly of TiO_2 nanoparticles [10] or Au nanoparticles [9] upon polarization at the liquid-liquid interface. Positively charged TiO_2 particles (at pH 3.5) caused the capacitive current to increase at positive Galvani potentials, whereas at pH 11, TiO_2 nanoparticles are negatively charged due to ionisation of the oxide surface and so the interfacial excess of the particles is increased at negative potentials. The potential value for the minimum capacitance shifted to higher potential values with the increase of silica nanoparticles amount in suspension (from -0.062 V in the absence of silica nanoparticles to -0.013 V for 0.1 g L^{-1}). This observation agrees with previous studies on the behaviour of haemoglobin at the liquid-liquid interface where positively charged haemoglobin caused a shift to lower potential values [42]. At a higher ionic strength, I , of the aqueous phase ($I = 100 \text{ mM}$), the potential for the lowest capacitance value was slightly shifted from $+0.020$ to $+0.035 \text{ V}$ when the concentration of silica nanoparticles increased from 0 to 0.1 g L^{-1} (Figure 1C). The capacitance variation was very limited due to the charge screening of higher electrolyte concentration. When the pH was decreased to 2 (10 mM HCl) no significant potential shift or increase in capacitance values was observed (Figure 1C). At this pH, silica nanoparticles are expected to be nearly neutral [44] explaining why changes in the potential of zero charge or in the capacitance were not observed. In the previous experiments, the electrochemical cell was prepared with an organic phase denser than water (the organic solvent was 1,2-dichloroethane, $d = 1.25 \text{ g mL}^{-1}$). To prove that silica nanoparticles assembled at the liquid-liquid interface, and did not just sediment because of gravity, an organic phase less dense than water was used (Figure 2). Recent studies have shown that a mixture of 5-nonanone and

1,2-dichloroethane is lighter than water and it can be used for electrochemistry at the ITIES [45]. We recorded AC voltammetry at such an electrochemical cell in the absence and in the presence of 0.5 g L^{-1} silica nanoparticles in the aqueous phase. The potential of zero charge shifted to higher potential values and the capacitance increased at -0.195 V from 26 to $32 \mu\text{F cm}^{-2}$ in the presence of silica nanoparticles. This observation confirmed that silica nanoparticles adsorbed at the ITIES.

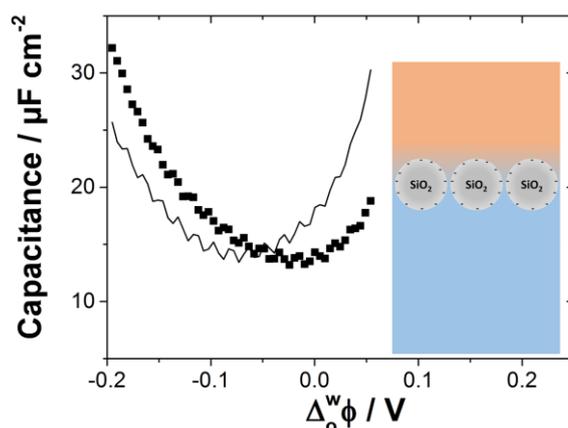


Figure 2: Potential dependence of the interfacial capacitance for the blank cell (solid line) and for (■) 0.5 g L^{-1} of silica nanoparticles with a denser aqueous phase. Electrochemical cell 4.

3.2 Silica nanoparticles interactions studies in aqueous solutions

Interactions between the negatively charged silica nanoparticles and charged organic dyes (cationic methylene blue, MB^+ ; anionic Eosin B, EB^-) were studied in a monophasic aqueous system. The organic dye was dissolved in the aqueous phase at $50 \mu\text{M}$ and a UV-vis absorption spectrum was recorded before and after a silica nanoparticle suspension was added to reach a concentration of 0.5 g L^{-1} . In the absence of silica nanoparticles, two bands were observed in the UV-vis. spectrum of MB^+ (Figure 3). The main band at $\lambda = 664 \text{ nm}$ is attributed to the $n-\pi^*$ transition whereas the shoulder at $\lambda = 614 \text{ nm}$ is attributed to the vibronic transition [46–48]. In the presence of silica nanoparticles, the UV-vis. spectrum looks quite different with the largest band at $\lambda = 590 \text{ nm}$ and a smaller at $\lambda = 664 \text{ nm}$, suggesting that MB^+ and silica nanoparticles interact with one another. This blue shift is attributed to the aggregation of MB^+ to form dimers, trimers...[49] on the silica nanoparticles surface and is in good agreement with spectra reported in the literature for MB^+ adsorbed onto silica-coated magnetic nanoparticles [50]. There was no difference between the EB^- spectra in the absence or in the presence of silica nanoparticles,

showing that there was no interaction between both negatively charged species (data not shown).

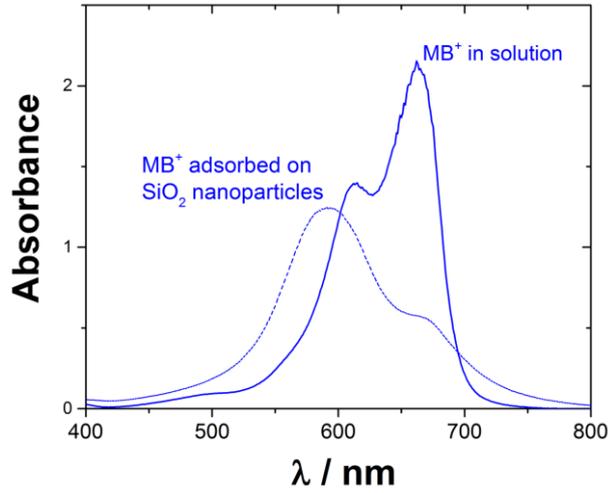
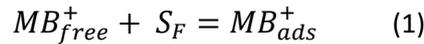


Figure 3: UV/Vis absorption spectroscopy of 50 μM MB^+ free in solution (solid line) and adsorbed on 0.5 g L^{-1} of silica nanoparticles (dashed line). Both solutions contained 10 mM LiCl.

We aimed at determining the total number of sites, S_T , available for MB^+ adsorption on silica nanoparticles and the association constant, K_{ads}^{aq} , of adsorption of MB^+ onto silica nanoparticles. The adsorption of free MB^+ in solution onto silica free sites, S_F , can be described by the following equilibrium.



$$K_{ads}^{aq} = \frac{[\text{MB}^+]_{ads}}{[\text{MB}^+]_{free}[S_F]} \quad (2)$$

The total number of sites is equal to the sum of the occupied sites, S_O , and of the free sites, S_F . Similarly, the total number of MB^+ is equal to the sum of adsorbed MB^+ and of MB^+ free in solution. Since we worked at constant volume, equations 3 and 4 are valid.

$$[S_T] = [S_O] + [S_F] \quad (3)$$

$$[\text{MB}^+]_T = [\text{MB}^+]_{ads} + [\text{MB}^+]_{free} \quad (4)$$

We assume that the number of occupied sites is equal to the number of MB^+ molecules adsorbed, equation (2) becomes:

$$K_{ads}^{aq} = \frac{[\text{MB}^+]_{ads}}{([\text{MB}^+]_T - [\text{MB}^+]_{ads})([S_T] - [\text{MB}^+]_{ads})} \quad (5)$$

Experimentally, we investigated the adsorption of MB^+ onto silica by UV-vis spectroscopy adsorption. We maintained the initial amount in solution of MB^+ constant, $[\text{MB}^+]_T = 50 \mu\text{M}$,

throughout the series of experiments and we varied the silica concentration from 0.040 to 0.867 g L⁻¹. Silica nanoparticles and MB⁺ were mixed together and the solution was then filtered through a membrane. The filtrate was collected and analysed by UV/Vis absorption spectroscopy, which allowed the determination of the amount of MB⁺ present in the filtrate, $[MB^+]_{free}$ deduced from the optical density variation at 650 nm. $[MB^+]_{ads}$ were then deduced from equation (4) and the results were plotted as the function of silica concentration and fitted with a Langmuir isotherm to determine (a) the SiO₂ surface area per MB⁺ (i.e. equivalent to the total site concentration) and (b) the adsorption constant (K_{ads}^{aq}). These two parameters were left as adjustable parameters [51]. The best fit was obtained for $K_{ads}^{aq} = 1.66 \cdot 10^5$ and $a = 1 \text{ nm}^2 / \text{MB}^+$. This represents approximately 900 adsorption sites per nanoparticle. K_{ads}^{aq} can be related to the free Gibbs energy of adsorption:

$$\Delta G_{SiO_2-MB^+}^0 = -RT \ln K_{ads}^{aq} \quad (6)$$

Where T is the temperature ($T = 293 \text{ K}$) and R is the universal gas constant ($R = 8.314 \text{ J mol}^{-1} \text{ K}^{-1}$). The value of $-29.3 \text{ kJ mol}^{-1}$ for the free Gibbs energy of adsorption testified of the spontaneous nature of the adsorption of MB⁺ onto silica nanoparticles. These results are in agreement with the existing literature, where spontaneous adsorption of MB⁺ on mesoporous silica formed as a thin film on an electrode was reported [52] with similar free Gibbs energy of adsorption. Based on the fitting of Figure 4, the surface area per MB⁺ was calculated at 1 nm^2 , which is lower than the results of the adsorption of one MB⁺ molecule for every 3 nm^2 of silica reported in the literature for a different type of silica and experimental conditions [9].

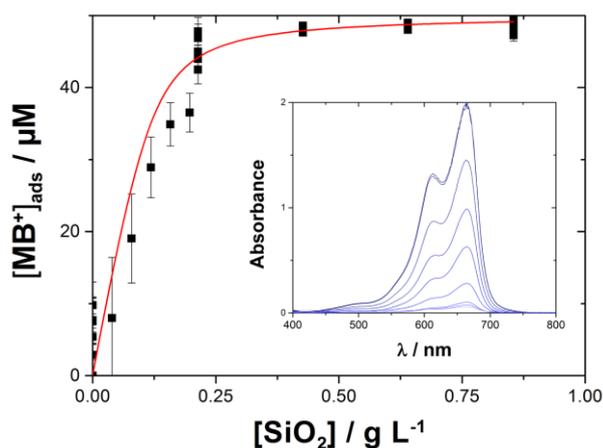


Figure 4: Amount of adsorbed MB⁺ as a function of the amount of silica nanoparticles. Experimental data represented by the black squares were fitted using a Langmuir isotherm. Inset shows the UV-vis absorption spectra of free MB⁺ in the filtrate after mixing with concentrations of silica nanoparticles (between 0.040 and 0.867 g L⁻¹) and filtration.

3.3 Ion transfer at the ITIES modified with silica nanoparticles

As mentioned in the previous section, UV-vis absorption spectroscopy experiments showed a strong affinity between the MB^+ and silica nanoparticles in an aqueous solution. Cyclic voltammetry was used to study the effect of the interaction between silica nanoparticles and ions on the ion transfer at the ITIES. In the absence of silica nanoparticles, the standard transfer potential of MB^+ was measured at $\Delta_o^w \phi_{MB^+}^{0'} = -0.223$ V (Figure 5), which was only 0.034 V larger than the open circuit potential of electrochemical cell 1 ($\Delta_o^w \phi_{ocp} = -0.257$ V). The concentration ratio of MB^+ in both phases at open circuit potential can be calculated using the Nernst equation:

$$\frac{[MB^+]_{org}}{[MB^+]_{aq}} = e^{(\Delta_o^w \phi_{ocp} - \Delta_o^w \phi_{MB^+}^{0'}) \frac{zF}{RT}} \quad (7)$$

$\Delta_o^w \phi_{MB^+}^{0'}$ is the formal transfer potential of MB^+ , z is the MB^+ charge and F is the Faraday constant. At open circuit potential, the ratio $\frac{[MB^+]_{org}}{[MB^+]_{aq}}$ was 0.26 in the absence of silica nanoparticles at the interface, confirming the spontaneous MB^+ transfer from the aqueous to the organic phase. The spontaneous transfer was observed experimentally at open circuit potential with a blue coloration of the organic solution. However, in the presence of silica nanoparticles, the transfer potential of MB^+ , $\Delta_o^w \phi_{MB^+}^{1/2}$, shifts to a potential of -0.131 V, confirming the strong affinity between silica nanoparticles and MB^+ . $\Delta_o^w \phi_{MB^+}^{1/2}$ is calculated as the half-wave potential of MB^+ transfer in the presence of silica nanoparticles adsorbed at the ITIES. At open circuit potential and in the presence of silica nanoparticles, the ratio $\frac{[MB^+]_{org}}{[MB^+]_{aq}}$, calculated with Eq (7), was 0.007. Less than 1 % of MB^+ transfers spontaneously across the interface modified with silica nanoparticles. The shift of transfer potential allows the estimation of the Gibbs free energy of ion association between the MB^+ and the silica nanoparticle adsorbed at the ITIES [53–55].

$$\Delta G_{SiO_2-MB^+}^{0,interface} = zF \left(\Delta_o^w \phi_{MB^+}^{0'} - \Delta_o^w \phi_{MB^+}^{1/2} \right) \quad (8)$$

$\Delta G_{SiO_2-MB^+}^{0,interface}$ was -8.9 kJ mol⁻¹, which is lower than the Gibbs free energy of adsorption, $\Delta G_{SiO_2-MB^+}^0$, previously determined in the aqueous phase by UV-vis spectroscopy (Figure 4). The difference between the two energies can be explained by the amphiphilic nature of MB^+

[56]. Indeed, the presence of the organic solvent at the interface lowers the adsorption constant of MB^+ onto silica nanoparticles.

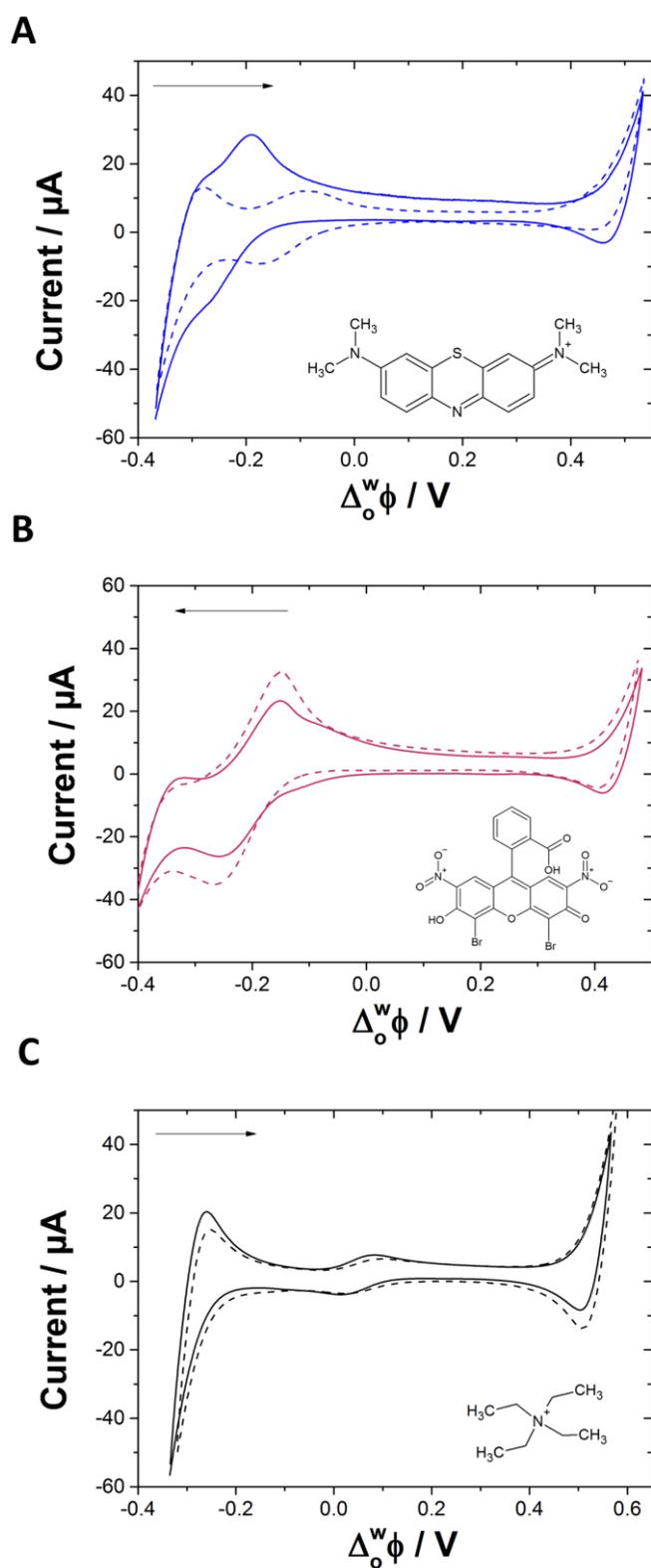


Figure 5: Cyclic voltammetry of (A) $600 \mu\text{M}$ of MB^+ , (B) $300 \mu\text{M}$ of EB^- and (C) $100 \mu\text{M}$ of TEA^+ in the absence (solid line) and in the presence (dashed line) of (A,B) 0.867 g L^{-1} and (C)

1 g L⁻¹ of silica nanoparticles in the aqueous phase. Electrochemical cell 1. $\nu = 5 \text{ mV s}^{-1}$. Arrows indicate the initial scan direction.

We also investigated the influence of the presence of silica nanoparticles on anionic EB⁻ and cationic TEA⁺. The transfer potential of both EB⁻ (+0.011 V) and TEA⁺ (+0.017 V) increased slightly after addition to the silica nanoparticles to the aqueous phase, suggesting that the transfer for these two ions is hindered by the presence of the nanoparticles. Thermodynamic values determined from the cyclic voltammograms for the three different ions are summarised in Table 1.

Table 1: Thermodynamics parameters determined experimentally

Ion	$\Delta_o^w \phi_{ion}^{0'}$ / V	ΔG_{ion}^0 / kJ mol ⁻¹	$\Delta_o^w \phi_{ion}^{1/2}$ / V	$\Delta G_{SiO_2-ion}^{0,interface}$ / kJ mol ⁻¹
MB ⁺	-0.223	-20.9	-0.131	-8.9
EB ⁻	-0.204	-19.7	-0.193	-1.1
TEA ⁺	+0.049	+4.7	+0.066	-1.6

The interfacial mechanism of MB⁺ transfer in the presence of silica nanoparticles is summarised in Figure 6. MB⁺ have shown a strong affinity constant, K_{ads}^{aq} , for the silica nanoparticles. When silica nanoparticles are adsorbed on the interface, MB⁺ transfer remained possible, provided that a sufficient excess of energy is provided: $\Delta G_{SiO_2-MB^+}^{0,interface}$. Based on this mechanism, we can harness the affinity between MB⁺ and silica nanoparticles and the adsorption of silica nanoparticles at the ITIES to accumulate MB⁺ ions at the ITIES and increase locally the MB⁺ concentration.

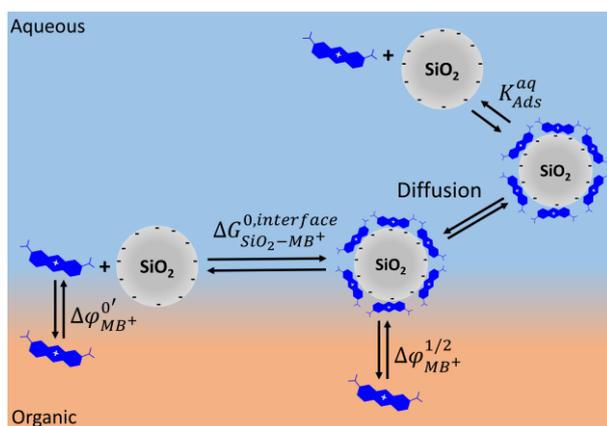


Figure 6: Schematic of the interfacial mechanism for the transfer of MB^+ in the presence of silica nanoparticles.

To demonstrate the accumulation of MB^+ at ITIES modified with silica nanoparticles, a solution of $25 \mu\text{M}$ MB^+ was mixed with 0.05 g L^{-1} of silica nanoparticles. The solution was centrifuged and the supernatant removed. The nanoparticles were then re-suspended in a 10 mM LiCl solution, which was used to prepare an electrochemical cell. Cyclic voltammograms obtained with such an electrochemical cell showed MB^+ transfer, whereas no transfer was observed for a $25 \mu\text{M}$ MB^+ solution in the absence of silica nanoparticles (Figure 7). Such an experiment demonstrated that MB^+ ions were accumulated at the ITIES with the help of the silica nanoparticles.

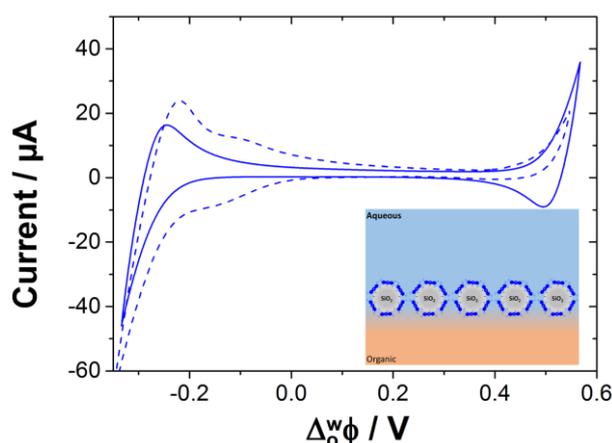


Figure 7: Cyclic voltammetry of $25 \mu\text{M}$ of MB^+ in the absence of silica nanoparticles (solid line) and in the presence of silica nanoparticles and MB^+ after centrifugation and re-suspension with 3 mL of 10 mM LiCl . Initial concentrations were $25 \mu\text{M}$ MB^+ and 0.05 g L^{-1} of silica nanoparticles. Electrochemical cell 1. $\nu = 5 \text{ mV s}^{-1}$.

Conclusions

UV-vis absorption spectroscopy and electrochemistry experiments have contributed to the determination of the mechanism of MB^+ ion transfer at the ITIES modified with silica nanoparticles. AC voltammetry experiments have shown that, upon addition of silica nanoparticles to the aqueous phase, the capacitance value increased, indicating that the quantity of charge located at the interface grew in the presence of silica nanoparticles. This behaviour can be attributed to the adsorption of negatively charged silica nanoparticles accumulating at the ITIES.

Interactions between the negatively charged silica nanoparticles and charged organic dyes (cationic methylene blue, MB^+ ; anionic Eosin B, EB^-) were studied in a monophasic aqueous system. A change in the shape of the UV-vis spectrum for MB^+ in the presence of silica nanoparticles suggested that MB^+ and silica nanoparticles interact with one another. The variation of $[\text{MB}^+]_{\text{ads}}$ was plotted as the function of silica concentration. Experimental results were fitted with a Langmuir isotherm to determine the SiO_2 surface area per MB^+ (i.e. equivalent to the total site concentration) to be $1 \text{ nm}^2 / \text{MB}^+$ and the adsorption constant ($K_{\text{ads}}^{\text{aq}}$) equal to $1.66 \cdot 10^5$, indicating a high affinity of MB^+ for the silica nanoparticle surface. The strong affinity for cationic MB^+ on the negatively charged silica nanoparticles was also confirmed by cyclic voltammetry, which monitored the transfer potential of the ions at the ITIES. In the presence of silica nanoparticles the transfer potential of MB^+ shifts to higher potential. The strong affinity of MB^+ for silica nanoparticles and their adsorption at the ITIES showed that MB^+ can be accumulated at the interface. As a result, the transfer of MB^+ in the presence of silica nanoparticle was monitored at MB^+ concentrations that were not detected in the absence of silica nanoparticles.

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