

Effects of ethylene glycol saturation protocols on XRD patterns: a critical review and discussion

Régine Mosser-Ruck, Karine Devineau, Delphine Charpentier, Michel Cathelineau

► **To cite this version:**

Régine Mosser-Ruck, Karine Devineau, Delphine Charpentier, Michel Cathelineau. Effects of ethylene glycol saturation protocols on XRD patterns: a critical review and discussion. *Clays and Clay Minerals*, Clay Minerals Society, 2005, 53 (6), pp.631-638. 10.1346/CCMN.2005.0530609 . hal-01876678

HAL Id: hal-01876678

<https://hal.univ-lorraine.fr/hal-01876678>

Submitted on 18 Sep 2018

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

EFFECTS OF ETHYLENE GLYCOL SATURATION PROTOCOLS ON XRD PATTERNS: A CRITICAL REVIEW AND DISCUSSION

REGINE MOSSER-RUCK^{1,*}, KARINE DEVINEAU², DELPHINE CHARPENTIER¹ AND MICHEL CATHELINÉAU¹

¹ UMR CNRS 7566 G2R, UHP, BP 239, 54506 Vandœuvre-lès-Nancy, France

² UMR CNRS 7569 LEM, INPL, BP 40, 54501 Vandœuvre-lès-Nancy, France

Abstract—The study of the transformation of smectite to illite, chlorite or vermiculite *via* interstratified clay minerals needs precise qualitative and quantitative determinations of the different layers in the mixed-layer clays and is generally based on X-ray diffraction (XRD) patterns after specific treatments of the clay samples. Saturation with K or Mg followed by ethylene glycol (EG) solvation are classical methods used to identify high-charge smectite and vermiculite. These procedures have been applied to two experimental clays, one composed of smectite layers and the second, a mixture of vermiculite and smectite layers. Different methods of glycolation (EG vapor or liquid EG) produce significant differences in the XRD patterns. Comparison with literature data indicates that K-saturated, high-charge smectite ($\sim 0.8 < \text{total charge} < 1/\text{unit-cell}$) and Mg-vermiculite (whatever its charge) do not expand in ethylene glycol vapor (d values $\sim 14\text{--}15$ Å). Expansion to 17 Å in liquid ethylene glycol occurs for Mg-vermiculite with a total charge of $< \sim 1.2/\text{unit-cell}$ and for K-saturated, high-charge smectite, when the tetrahedral charge is $< \sim 0.7/\text{unit-cell}$. This study shows that: (1) procedures of glycolation need to be standardized; (2) the use of saturation protocols using both liquid ethylene glycol and ethylene glycol vapor yields useful additional information about the distribution of charges in clay minerals.

Key Words—Ethylene Glycol Solvation, High-charge Smectite, K Saturation, Mg Saturation, Vermiculite.

INTRODUCTION

The transformation of smectites under natural or experimental conditions has been studied intensely in order to describe the processes controlling the nature of clays during diagenesis, weathering or anthropogenic activities such as radioactive waste disposal. In most studies, the characterization of clay layers is carried out by XRD on air-dried and glycol-saturated oriented preparations, sometimes after saturation with different cations (K, Mg, Li, *etc.*, see Bouchet *et al.*, 2000 for a compilation). The main objective of glycolation is to identify the nature of the non-swelling clays (illite or chlorite), to discriminate the swelling clays (di- or trioctahedral high-charge or low-charge smectite, vermiculite) and to quantify the nature, degree and ordering of the mixed layering. For a long time, the complexes formed by ethylene glycol or glycerol with clays have received attention because of their use in the identification and characterization of smectites and mixed-layer clays, in particular illite-smectite mixed-layer clays (MacEwan, 1944, 1946; Bradley, 1945; Brindley, 1966). Saturation with glycerol or ethylene glycol produces similar effects but some differences are noticed such as higher d values with glycerol due to the larger size of the molecule, and a variable number of layers of

organic molecules related to the exchangeable cation (reviewed by Bouchet *et al.*, 2000). The expandability of these complexes varies not only with the nature of the mineral, but also with the magnitude and source of the charge in the silicate layers, the number and kind of the exchangeable cations, the particle size and the relative humidity (Walker 1957, 1958; Brindley, 1966; Schultz, 1969; Harward *et al.*, 1969; Środoń, 1980; Sato *et al.*, 1992). In general, it is established that vermiculites form a single-layer complex and smectites a double-layer complex. Beidellite, in comparison with montmorillonite, exhibits differences in expansion upon solvation which are consistent with stronger ionic attraction for tetrahedral than for octahedral charge sites (Harward and Brindley, 1965). The physical state of the solvating agent (liquid, vapor) also yields differences in the rate of expansion (Harward and Brindley, 1965, in Brindley, 1966; Suquet *et al.*, 1975). As a preliminary conclusion, methods of saturation have changed continuously since the middle of 20th century. A careful reading of the literature (summarized below) has shown that several differences and inconsistencies characterize the methods of saturation with polyalcohols (ethylene glycol and glycerol) and that no international standardization of these procedures is available. Therefore, an experimental study was conducted in order to evaluate the relative effects produced by distinct procedures of saturation on the expandability of two clays. The effects of the ethylene glycol saturation method (vapor or liquid) on the swelling was tested for a K-saturated smectite and a Mg-saturated vermiculite.

* E-mail address of corresponding author:
regine.ruck@g2r.uhp-nancy.fr
DOI: 10.1346/CCMN.2005.0530609

Saturation with liquid polyalcohol

Liquid glycerol or liquid ethylene glycol was used in the pioneer works (Bradley 1945; MacEwan, 1948). At that time, air-dried powdered samples were packed in glass capillary tubes, then the sealed ends of the tubes were broken off and held in hot (110°C) water-free glycerol until glycerol rose by capillarity to the top of the sample (e.g. Barshad, 1950; Reynolds, 1965). In the case of a thin oriented layer of air-dried clay on a glass slide, a few drops of the organic liquid were placed at the edge of the clay sample and allowed to diffuse into it (e.g. Harward *et al.*, 1969; Brindley and Ertem, 1971), or a fine spray (e.g. Inoue *et al.*, 1989) was used so that clay becomes visibly moist. Solvation was also realized by pressing the dry preparation upside down against absorbent paper wetted with ethylene glycol and allowing the composite to remain overnight in this position (e.g. Środoń, 1980). Ethylene glycol liquid was sometimes applied directly with a brush (Lanson and Besson, 1992) and excess ethylene glycol was removed just before recording the XRD pattern by pressing the preparation on an absorbent paper. But direct wetting of the slide with the polyalcohols can lead to mechanical disturbances, in particular to the change of particle orientation (Brunton, 1955; Kunze, 1955; Brown and Farrow, 1956).

Saturation with polyalcohol vapor

Another procedure was then recommended that involved exposing a clay-covered slide to organic vapors in a closed vessel over a heated bath maintained at 60–65°C for ethylene glycol (Brunton, 1955; Kunze, 1955) and 100°C for glycerol (Brown and Farrow, 1956). Brunton (1955) observed that most clay samples glycolated satisfactorily within 1 h in the vapor at equilibrium with a bath of liquid ethylene glycol at 60°C, but Brindley (1966) asserted that at least an overnight period in ethylene glycol or glycerol vapors was necessary to establish optimum conditions of solvation. According to Holtzapffel (1985), the saturation in a polyalcohol vapor atmosphere at room temperature should be complete after 12 h.

In the literature, preparation methods of clay samples before an XRD analysis are very different or poorly described. In a few examples, it is stated only that samples are ethylene glycol solvated: “in vapor at 60°C” (Inoue *et al.*, 1989); “at 60°C for at least 6 h” (Howard 1981), “at least 12 h at ~50°C” (Bish and Aronson, 1993); only “overnight” (Huang *et al.*, 1993); “overnight at 60°C” (Whitney and Velde, 1993); or “at room temperature for 24 h” (Price and McDowell, 1993). No international standard exists concerning the procedure for saturation with polyalcohol vapor.

Effects of K saturation and layer charge of smectite on swelling with polyalcohol

The K saturation of montmorillonite differs with the site of the charge within a layer structure (Schultz,

1969). A charge in the tetrahedral sheet near the interlayer cation will strongly attract and tends to fix K⁺ ions, making the layer non-expandable, whereas a layer charge mainly in the octahedral sheet far from the exchangeable cation tends to let the layer expand. Other authors (see review in Schultz, 1969) report that expansion after K saturation is related more to the total amount of net layer charge than to the position of the charge within the layer. According to Schultz (1969), the formation of a double-layer ethylene glycol complex (*d* value = 17 Å) occurs if the negative layer charge of the smectite is 0.85 or less per unit-cell (O₂₀ (OH)₄). According to Sato *et al.* (1992), K-smectites (montmorillonite and beidellite) with layer charges of 0.68–0.94 (per unit-cell) form 17 Å bi-layer complexes and K-smectites with layer charge of 1.12–1.17 (per unit-cell) form 14 Å monolayer complexes.

In recent works (Bouchet *et al.*, 1988; Righi *et al.*, 1997; Beaufort *et al.*, 2001; Guillaume *et al.*, 2004; Mosser-Ruck and Cathelineau 2004), the identification of high-charge smectite in mixed-layer clays is accomplished using a K-saturated sample heated at 110°C overnight and then solvated by ethylene glycol vapor. This procedure is often used and cited in the literature but authors sometimes do not refer to the source, or they refer to sources in which the procedure described is slightly different. Bouchet *et al.* (1988) refer to Środoń (1980), Howard and Roy (1985) and Środoń *et al.* (1986). Righi *et al.* (1997) cite Schultz (1969) and Sato *et al.* (1992). Certainly, all these works (Schultz 1969; Środoń 1980; Howard and Roy 1985; Środoń *et al.*, 1986; Sato *et al.*, 1992) examine the behavior of expandable smectite layers upon K saturation to give an estimate of the magnitude of the layer charge but, curiously, none of them deals with heating overnight at 110°C on K-saturated samples before saturation with a polyalcohol. The reason why heating overnight at 110°C has been introduced to the protocol by some authors (Bouchet *et al.* 1988; Righi *et al.*, 1997) is not explained.

Furthermore, these studies do not use the same solvating agent for the preparation of samples. For example, in Sato *et al.* (1992), solvations with glycerol and ethylene glycol are accomplished by the vapor-pressure methods of Brunton (1955) and Brown and Farrow (1956). In Środoń (1980) and Środoń *et al.* (1986) liquid ethylene glycol is used by pressing the preparation against absorbent paper overnight and Schultz (1969) uses K-saturated clay slides heated at 300°C for 30 min and treated overnight in ethylene glycol vapor at 60°C.

Mg-saturation and distinction between smectite and vermiculite

Harward and Brindley (1965) show that synthetic beidellites and montmorillonite, saturated with Mg, and exposed to glycerol vapor are characterized by a spacing

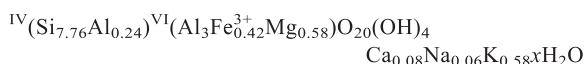
of 14.3 Å and 17.8–17.9 Å, respectively; exposed to liquid glycerol, a 17.8–17.9 Å spacing is obtained in both cases. The reason why a one-layer glycerol complex is formed in beidellite by glycerol vapor condensation but a two-layer complex in contact with excess liquid is unclear. According to Harward *et al.* (1969), the Mg-saturated montmorillonites and beidellites generally expand to the equivalent two-layer complex with ethylene glycol vapor (16.8–16.9 Å). Mg-saturated saponite, in the presence of liquid ethylene glycol or liquid glycerol has a similar behavior to that of the other smectites, *i.e.* swelling from 16.7 Å in ethylene glycol to 18.1 Å in glycerol (Suquet *et al.*, 1975). For Mg-saturated vermiculites, Harward *et al.* (1969) concluded that whatever the solvating procedure (ethylene glycol or glycerol, vapor or liquid), *d* values always range between 14.1 and 14.3 Å. These results disagree with those of Walker (1957, 1958) who studied the expansion properties upon solvation of five vermiculites characterized by different layer charge, from 1.2 to 1.6 (in equivalents per O₂₀(OH)₄ unit of structure). That author noticed that the rate and the degree of expansion of the five Mg-vermiculites saturated with liquid ethylene glycol are related to the layer charge of the vermiculite. The basal *d* value is ~14.3 Å for Mg-vermiculite with high layer-charge (>1.4 in equivalents per O₂₀(OH)₄), 16.3 Å if the layer charge is <1.2, and intermediate, ~15.2 Å when the layer charge is between 1.2 and 1.4. Since none of these Mg-saturated vermiculites expands beyond 14.5 Å when solvated with liquid glycerol, Walker (1957, 1958) recommends this solvation as a procedure to distinguish between vermiculite and smectite.

MATERIALS, METHODS AND ANALYTICAL TECHNIQUES

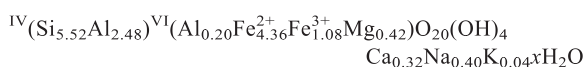
Materials

The effects of the state of ethylene glycol on the XRD pattern have been tested on two products.

Material 'A' was obtained experimentally by Mosser-Ruck and Cathelineau (2004). Its mean formula, based on O₂₀(OH)₄, is:



The second material (referred to here as material 'B') is a mixture of vermiculite and smectite. It was obtained experimentally by Charpentier *et al.* (2004). The characterization of vermiculite layers was obtained by analyses of particles (TEM analyses) and Mg saturation of material B before recording the XRD pattern. The structural formulae of the vermiculite in material B (before Mg saturation) is:



EG saturation protocols

Material A. A preliminary Ca saturation of material A followed by solvation with ethylene glycol vapor revealed that it was composed of smectite layers. A second saturation with K was carried out to determine the presence of high-charge smectite layers in material A and to test the effect of both heating to 110°C overnight and the state of ethylene glycol (vapor or liquid) following four different saturation protocols: (1) K saturation + 110°C (overnight) + 24 h-room temperature-EG vapor solvation; (2) K saturation + 24 h-room temperature-EG vapor solvation; (3) K saturation + 24 h-65°C-EG vapor solvation; (4) K saturation + room temperature-EG liquid solvation (a few drops of the organic liquid were placed at the edge of the clay film until the slide was visibly moist).

Material B. The effect of the state of ethylene glycol (vapor and liquid) has been tested following two different saturation protocols: (1) Mg saturation + 24 h-room temperature-EG vapor solvation; (2) Mg saturation + room temperature-liquid EG solvation.

Analytical techniques

XRD. The XRD patterns were recorded using a Bruker[®] D8 diffractometer, with CoK α radiation. The patterns were recorded from 3 to 40°2 θ , with a step scan of 0.035°2 θ and time per step of 3 s.

The XRD patterns were recorded under similar conditions of relative humidity (42–46%) and room temperature (22.6–25.8°C).

Analyses of material A and material B were carried out by electron microprobe (EMPA) and transmission electron microscopy (TEM) and were reported by Mosser-Ruck and Cathelineau (2004) and by Charpentier *et al.* (2004).

RESULTS

The four protocols carried out on K-saturated material A result in rather significant differences in the XRD patterns (Figure 1). If the K-saturated material A is solvated in EG vapor at 65°C (the widely used method in literature), the *d*₀₀₁ value is ~15 Å. The same result is obtained when it is solvated in EG vapor at room temperature over 24 h. Heating at 110°C overnight before EG vapor solvation led to a *d*₀₀₁ spacing of ~13 Å. When material A is saturated with liquid ethylene glycol, a higher *d* value is observed (17 Å). The presence of a weak reflection at 10.09 Å (only on the XRD of the K-110°C-EG vapor and K-EG vapor samples) is also noticed and could be assigned to non-swelling 'mica-like' crystals (Figure 1).

The XRD patterns presented in Figure 2 show that Mg-vermiculite in material B does not expand after EG vapor saturation (14.26 Å on AD and on EG patterns) whereas Mg-smectite expands (14.26 Å on AD pattern

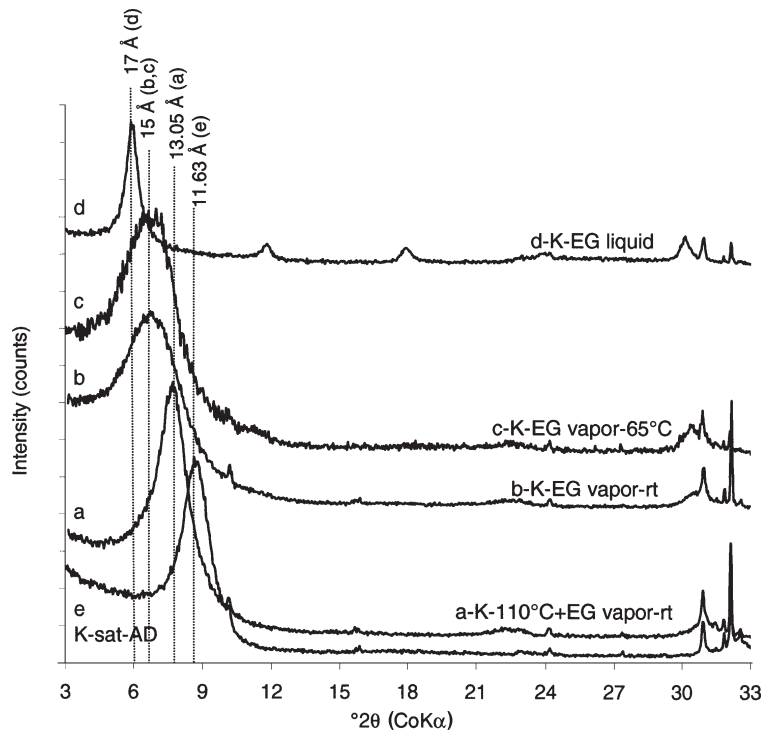


Figure 1. XRD patterns of the K-saturated material noted 'A': (a) heating overnight at 110°C and solvated with ethylene glycol vapor at room temperature (K-110°C-EG vapor-rt); (b) solvated with ethylene glycol vapor at room temperature (K-EG vapor-rt); (c) solvated with ethylene glycol vapor at 65°C (K-EG vapor-65°C); (d) solvated with liquid ethylene glycol; (e) air-dried (K-sat-AD). Index of reflections corresponds to d_{hkl} in Å.

and 16.92 Å on EG pattern). The d value at 9.04–9.08 Å is attributed to an associated mineral (mordenite). Solvation with liquid ethylene glycol was also carried

out and shows only one single reflection (d value = 16.92 Å). Mg-vermiculite layers in material B expand in liquid ethylene glycol but not in ethylene glycol vapor.

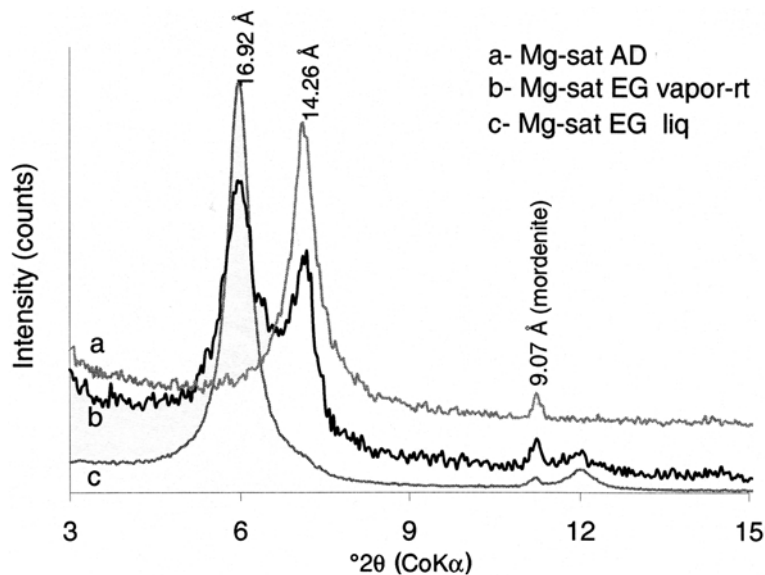


Figure 2. XRD patterns of the air-dried Mg-saturated material B (Mg-sat AD), solvated with ethylene glycol vapor for 24 h at room temperature (Mg-sat EG vapor-rt) and solvated with ethylene glycol liquid (Mg-sat EG liq). Index of reflections corresponds to d_{hkl} in Å.

DISCUSSION

Effects of heating at 110°C and the state of ethylene glycol on the XRD pattern of K-saturated clay samples

In our test, the variations in d values could be due to several factors: (1) Glass-slide preparations can quickly lose some of the interlayer ethylene glycol as already suggested by Kunze (1955). In our case, XRD patterns were recorded just after the ethylene glycol vapor saturation, so the hypothesis of an EG evaporation can be ruled out. Dyal and Hendricks (1952) demonstrated that evaporation of EG is not significant in K-saturated smectites in the 2 h following saturation. (2) Walker (1958) specifies that the initial hydration state of interlayer cations has an influence on swelling after glycolation. K-saturated smectites (~ 0.8 charge/unit-cell) collapse in a dry atmosphere to 10.3 Å (Sayegh *et al.*, 1965, Bérend *et al.*, 1995). Consequently, glycol retention depends strongly on whether or not the clay has been dried prior to glycol solvation. Air-dried K-saturated montmorillonite was found to retain ~ 1.5 times as much EG as the dried samples (Dyal and Hendricks, 1952). The low glycol retention seen in dried

K-montmorillonite was considered to be related to K fixation. The role of water in the formation of complexes is predominant and heating perhaps deteriorates the ability to re-expand both in organic liquids and in water itself. Traces of water would be essential to the formation of complexes (Brindley, 1966). These reasons could explain the difference in d values observed when our sample was dried overnight at 110°C before the EG saturation. (3) Variations in d values are also related to differences in orientation and packing of the organic molecules (Brindley, 1966). The simple geometry of the ethylene glycol molecule may facilitate different degrees of close-packing as the layer charge of clay minerals increases; single or double-layer complexes of EG can form and are related to the net layer charge of the clay minerals (Brindley, 1966). In our case, the material studied is the same for the four ethylene glycol treatments. Consequently, it is not possible to invoke a variation of the layer charge to explain the different expansions observed on the XRD patterns. In Table 1, literature data show clearly that the swelling of K-saturated smectite with ethylene glycol is dependent on both the total charge and its distribution in the

Table 1. Mean d_{001} (Å) values of ethylene glycol (EG) solvation complex of K-saturated smectites.

Reference	Sample	Solvation method	Basal spacings (Å)
This study	K-saturated material 'A' Total charge = 0.8/unit-cell	EG vapor (r.t.)	15 (10.09)* * 'mica-like' crystal
		EG vapor (65°C)	15 (10.05)* * 'mica like' crystal
		110°C, 24 h before EG vapor (r.t.) Liquid EG (r.t.)	13.04 (10.09)* * 'mica-like' crystal 17
† Brindley, 1966	† K-saturated smectites (total charge = 0.7)	Liquid EG	16.9 to 17.1
Schultz, 1969	K-saturated smectites (83 samples)	300°C, 1/2 h before EG vapor (60°C)	17 if total layer charge < 0.85/unit-cell 13–16 if >0.85/unit-cell
Weir, 1960 in Suquet <i>et al.</i> , 1975	K-saturated beidellite [#] (Black Jack Mine-Idaho, total layer charge = 0.96/unit- cell, tetrahedral charge = 0.96)	Liquid EG	14.7
Suquet <i>et al.</i> , 1975	K-saturated saponite (total charge = 0.92/unit-cell)	Liquid EG	16.5
Glaeser – personal communication in Suquet <i>et al.</i> , 1975	K-saturated beidellite [§] (Rupsroth- Germany, total charge = 1.17, tetrahedral charge = 0.75)	Liquid EG	14
Sato <i>et al.</i> , 1992	K-saturated smectites (10 samples)	EG vapor (65°C)	16.9–17.34 if total layer is ~ 0.68 – 0.94 /unit-cell and tetrahedral charge < 0.5 /unit-cell 14 if total charge is 1.12 – 1.17 /unit-cell and tetrahedral charge > 0.7 /unit-cell

r.t.: solvated at room temperature; †: data compiled from different authors. [#]Layer charge mentioned by Schultz (1969). [§]Layer charge mentioned by Sato *et al.* (1992).

different crystallographic sites. Moreover, the behavior of a further sample of the same K-saturated smectite would be different depending on the state of the ethylene glycol used to solvate it. K-saturated smectites characterized by a total layer charge > 1 (per unit-cell) do not expand in ethylene glycol vapor at 60–65°C (Sato *et al.*, 1992) or in liquid ethylene glycol (Weir, 1960; Glaeser, pers. comm., in Suquet *et al.*, 1975). If the total charge is between 0.85 and 1, K-saturated smectites generally do not expand in ethylene glycol vapor (Schultz, 1969) except if tetrahedral charge is very low < 0.5 (Sato *et al.*, 1992), and they expand to 16.9–17 Å in liquid ethylene glycol (Suquet *et al.*, 1975). This contrasting behavior could be related to the amount of the tetrahedral charge of the K-saturated smectite. Expansion in liquid ethylene glycol could be possible for K-saturated high-charge smectite ($\sim 0.8 < \text{total charge} < 1/\text{unit-cell}$) only if tetrahedral charge is < 0.7 , referring to the results of Glaeser (in Suquet *et al.*, 1975). This would require that certain layers admit liquid EG but not EG vapor. Comparison of our results with the data in Table 1 suggests that material A could be a mixed-layered clay composed of both low-charge and high-charge smectite layers ($\sim 0.8 < \text{total charge} < \sim 1/\text{unit-cell}$) because of a basal spacing of 15 Å when the sample is saturated in EG vapor. But the expandability (17 Å) observed when it is saturated with liquid EG can be due to the presence of high-charge smectite with tetrahedral charge $< 0.7/\text{unit-cell}$.

Effect of the state of ethylene glycol on the XRD pattern of Mg-saturated clay sample

Our results show that the contrasting behavior of Mg-vermiculite in material B is dependent on its layer charge, confirming the literature data (Table 2). Displacement of the basal reflection towards higher

spacings (16.3–17 Å) when Mg-vermiculite is solvated in liquid ethylene glycol, is noticed only for low-charge vermiculite (< 1.2 per $\text{O}_{20}(\text{OH})_4$), such as that from Young River in Australia (Walker, 1957, 1958). This is also typically the case for the vermiculite in material B. By contrast, after solvation in ethylene glycol vapor, the basal reflection always occurs at ~ 14.2 Å whatever the layer charge of vermiculite.

CONCLUSIONS

This study has shown that differences and inconsistencies exist in literature concerning the methods of saturation with polyalcohols (ethylene glycol and glycerol). No international standards for these procedures are available at the present time.

Based on XRD results obtained after different saturation protocols of a K-saturated smectite and a Mg-vermiculite, our study shows that: (1) treatment with EG vapor at $\sim 60^\circ\text{C}$ or at room temperature leads to the same saturation ratios if duration of solvation at room temperature is at least 24 h; (2) heating to 110°C overnight could induce either a partial or a complete lack of re-expansion of K-saturated smectite layers in ethylene glycol vapor; (3) liquid ethylene glycol leads to the expansion to 17 Å of K-saturated, high-charge smectite layers when they have a low tetrahedral charge ($< \sim 0.7/\text{unit-cell}$), and of low-charge Mg-saturated vermiculite ($< 1.2/\text{unit-cell}$).

The various behaviors of K-saturated, high-charge smectite and Mg-saturated vermiculite facing the different treatments confirm the strong dependency of expandability to the saturation procedure with polyalcohol and the difficulty of correlating expandability with the mineralogy and the type of layering in mixed-layered clays. The present study is restricted to two

Table 2. Mean values of d_{001} (Å) of ethylene glycol (EG) solvation complex of Mg-saturated vermiculites.

Reference	Sample	Solvation method	Basal spacings (Å)
This study	Mg-saturated material 'B' vermiculite Total layer charge = 1.16	EG vapor (r.t.)	14.26
		Liquid EG (r.t.)	16.92
Walker (1958)	Mg-vermiculite	Liquid EG	14.3 if layer charge is > 1.4 per $\text{O}_{20}(\text{OH})_4$ 16.3 if layer charge is 1.2 or less 15.2 for intermediate layer charge (1.3)
† Brindley (1966) *vermiculite from Young River (Australia)	† Mg-vermiculite	Liquid EG	14.3; 16.3* (*total layer charge = 1.2)
Harward <i>et al.</i> (1969)	Mg-vermiculites (12 samples) layer charge > 1.3	EG vapor (60–65°C)	14.22
		Liquid EG	14.25

r.t.: solvated at room temperature; †: data compiled from different authors.

samples obtained from experiments. To confirm the conclusions, well-known reference clays should be necessary to characterize fully the inferred processes using various polyalcohol saturation protocols.

ACKNOWLEDGMENTS

We are very grateful to Dr J. Cuadros (National History Museum, London) and Dr R. Ferrell (Associate Editor) for reviewing the manuscript and for their constructive suggestions.

REFERENCES

- Barshad, I. (1950) The effect of interlayer cations on the expansion of the mica type crystal lattice. *American Mineralogist*, **35**, 225–238.
- Beaufort, D., Berger, G., Lachapagne, J.C. and Meunier, A. (2001) An experimental alteration of montmorillonite to a di+trioctahedral smectite assemblage at 100 and 200°C. *Clay Minerals*, **36**, 211–225.
- Bérend, I., Cases, J.M., Francois, M., Uriot, J.P., Michot, L., Masion, A. and Thomas, F. (1995) Mechanisms of adsorption and desorption of water vapor by homoionic montmorillonites: 2. The Li⁺, Na⁺, Rb⁺, and Cs⁺-exchanged forms. *Clays and Clay Minerals*, **43**, 324–336.
- Bish, D.L. and Aronson, J.L. (1993) Paleogeothermal and paleohydrologic conditions in silicic tuff from Yucca Mountain, Nevada. *Clays and Clay Minerals*, **41**, 148–161.
- Bouchet, A., Proust, D., Meunier, A. and Beaufort, D. (1988) High-charge to low-charge smectite reaction in hydrothermal alteration processes. *Clay Minerals*, **23**, 133–146.
- Bouchet, A., Meunier, A. and Sardini, P. (2000) Minéraux argileux: structure cristalline, identification par diffraction des rayons X. *Bulletin du Centre de Recherche Elf Exploration et Production*, **23**, 136 pp.
- Bradley, W.F. (1945) Molecular association between montmorillonite and some polyfunctional organic liquids. *Journal of the American Chemical Society*, **67**, 975–981.
- Brindley, G.W. (1966) Ethylene glycol and glycerol complexes of smectite and vermiculites. *Clay Minerals*, **6**, 237–259.
- Brindley, G.W. and Ertem, G. (1971) Preparation and solvation properties of some variable charge montmorillonites. *Clays and Clay Minerals*, **19**, 399–404.
- Brown, G. and Farrow, R. (1956) Introduction of glycerol into flakes aggregates by vapour pressure. *Clay Minerals Bulletin*, **3**, 44–45.
- Brunton, G. (1955) Vapour glycolation. *American Mineralogist*, **40**, 124–126.
- Charpentier, D., Devineau, K., Mosser-Ruck, R., Cathelineau, M. and Villiéras, F. (2004) Bentonite-iron interactions under alkaline condition: an experimental approach (*submitted to Applied Clay Science*).
- Dyal, R.S. and Hendricks, S.B. (1952) Formation of mixed layer minerals by potassium fixation in montmorillonite. *Proceedings of the Soil Science Society of America*, **16**, 45–48.
- Guillaume, D., Neaman, A., Cathelineau, M., Mosser-Ruck, R., Peiffert, C., Abdelmoula, M., Dubessy, J., Villiéras, F. and Michau, N. (2004) Experimental study of the transformation of smectite at 80 and 300°C in the presence of Fe oxides. *Clay Minerals*, **39**, 17–34.
- Harward, M.E. and Brindley, G.W. (1965) Swelling properties of synthetic smectite in relation to lattice substitutions. *Clays and Clay Minerals*, **13**, 209–222.
- Harward, M.E., Carstea, D.D. and Sayegh, A.H. (1969) Properties of vermiculite and smectites: Expansion and collapse. *Clays and Clay Minerals*, **16**, 437–447.
- Holtzapffel, T. (1985) Les minéraux argileux. Préparation. Analyse diffractométrique et détermination. *Société Géologique du Nord*, **12**, 136 pp.
- Howard, J.J. (1981) Lithium and potassium saturation of illite/smectite clays from interlaminated shales and sandstones. *Clays and Clay Minerals*, **29**, 136–142.
- Howard, J.J. and Roy, D.M. (1985) Development of layer charge and kinetics of experimental smectite alteration. *Clays and Clay Minerals*, **33**, 81–88.
- Huang, W.L., Longo, J.M. and Pevear, D.R. (1993) An experimental derived kinetic model for smectite-to-illite conversion and its use as a geothermometer. *Clays and Clay Minerals*, **41**, 162–177.
- Inoue, A., Bouchet, A., Velde, B. and Meunier, A. (1989) Convenient technique for estimating smectite layer percentage in randomly interstratified illite/smectite minerals. *Clays and Clay Minerals*, **37**, 227–234.
- Kunze, G.W. (1955) Anomalies in the ethylene glycol solvation technique used in X-ray diffraction. *Clays and Clay Minerals*, **3**, 83–93.
- Lanson, B. and Besson, G. (1992) Characterization of the end of smectite-to-illite transformation: Decomposition of X-ray patterns. *Clays and Clay Minerals*, **40**, 40–52.
- MacEwan, D.M.C. (1944) Identification of the montmorillonite group of minerals by X-rays. *Nature*, **154**, 577–578.
- MacEwan, D.M.C. (1946) The identification and estimation of the montmorillonite group of minerals, with special reference to soil clays. *Journal of the Society of Chemical Industry*, **65**, 298–304.
- MacEwan, D.M.C. (1948) Complexes of clays with organic compounds. I. Complex formation between montmorillonite and halloysite and certain organic liquids. *Transactions of the Faraday Society*, **44**, 349–367.
- Mosser-Ruck, R. and Cathelineau, M. (2004) Experimental transformation of Na,Ca-smectite under basic conditions at 150°C. *Applied Clay Science*, **26**, 259–273.
- Price, K.L. and McDowell, S.D. (1993) Illite/smectite geothermometry of the Proterozoic Oronto group, mid-continent rift system. *Clays and Clay Minerals*, **41**, 134–147.
- Reynolds, R.C. (1965) An X-ray study of an ethylene glycol-montmorillonite complex. *American Mineralogist*, **50**, 990–1001.
- Righi, D., Räisänen, M.L. and Gillot, F. (1997) Clay mineral transformation in podzolized tills in central Finland. *Clay Minerals*, **32**, 531–544.
- Sato, T., Watanabe, T. and Otsuka, R. (1992) Effects of layer charge, charge location, and energy change on expansion properties of dioctahedral smectites. *Clays and Clay Minerals*, **40**, 103–113.
- Sayegh, A.H., Harward, M.E. and Knox, E.G. (1965) Humidity and temperature interaction with respect to K-saturated expanding clay minerals. *American Mineralogist*, **50**, 490–495.
- Schultz, L.G. (1969) Lithium and potassium absorption, dehydroxylation temperature, and structural water content of aluminous smectites. *Clays and Clay Minerals*, **17**, 115–149.
- Środoń, J. (1980) Precise identification of illite/smectite interstratification by X-ray powder diffraction. *Clay and Clay Minerals*, **28**, 401–411.
- Środoń, J., Morgan, D.J., Eslinger, E.V., Eberl, D.D. and Karlinger, M.R. (1986) Chemistry of illite/smectite and end-member illite. *Clays and Clay Minerals*, **34**, 368–378.
- Suqet, H., De la Calle, C. and Pezerat, H. (1975) Swelling and structural organization of saponite. *Clays and Clay Minerals*, **23**, 1–9.
- Walker, G.F. (1957) On the differentiation of vermiculite and smectites in clays. *Clay Minerals Bulletin*, **3**, 154–163.

- Walker, G.F. (1958) Reactions of expanding lattice clay minerals with glycerol and ethylene glycol. *Clay Minerals Bulletin*, **3**, 302–313.
- Weir, A.H. (1960) Relationship between physical properties, structure and composition of smectite. Thesis, London University.
- Whitney, G. and Velde, B. (1993) Changes in particle morphology during illitization: An experimental study. *Clays and Clay Minerals*, **41**, 209–218.

(Received 5 November 2004; revised 15 June 2005; Ms. 979; A.E. Ray E. Ferrell, Jr.)