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# Kinetic study of photocatalytic degradation of Ifos famide in a microchannel and simulation

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Research

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**ABSTRACT**

Among the possible solutions for the intensification of photocatalytic reactions, the use of microreactors is emerging. Such an approach was applied to the photodegradation of Ifosfamide in a flow microreactor with an immobilized catalyst composed of ZnO sensitized with CuInZnS<sub>2+x</sub> quantum dots. The conversion yield at a low flow rate was investigated. A computational flow simulation allows to determine an apparent constant rate and the diffusion coefficient towards the photocatalyst surface. Photocatalytic degradation of Ifosfamide can be fitted by the Langmuir-Hinshelwood model. The Sherwood number and the average value of Damköhler show that the photocatalytic reaction under our experimental conditions is limited by the mass transfer.

**Keywords:** Photocatalysis, Microreactor, Kinetics, Zinc Oxyde, Quantum Dots, Diffusion coefficient.

**INTRODUCTION**

Advanced oxidation processes and especially photocatalysis are of high interest for the decontamination of water containing pollutants as hospital effluents

[1,2]. Heterogeneous photocatalysts employed in a tubular reactor enable the control over reactivity and selectivity [3]. Traditionally, photochemical reactors are operated in batch mode. This configuration leads to a non-uniform distribution of UV light and a mass transfer limitation due to a slow diffusion.

Among the possible solutions for the intensification of photocatalytic reactions, the use of microreactors is emerging [4,5]. Microreactors with channel dimensions of a few hundred microns exhibit larger surface to volume ratio compared to classic reactors [6] and thus enable optimum utilization of incident radiation [7]. The observed reaction rate is at least one order of magnitude higher than that observed in conventional reactors [8].

Zinc oxide (ZnO) has been widely investigated over the last ten years as photocatalyst and has been demonstrated to be of high potential [9]. However, ZnO suffers from two main drawbacks. First, due to its wide bandgap energy ( $E_g = 3.3$  eV) [10]. ZnO can only be activated by light with a wavelength lower than ca. 380 nm to trigger the electron-hole separa-

tion. Second, ZnO exhibits a low quantum efficiency due to the fast recombination of photo generated charge carriers (electrons and holes).

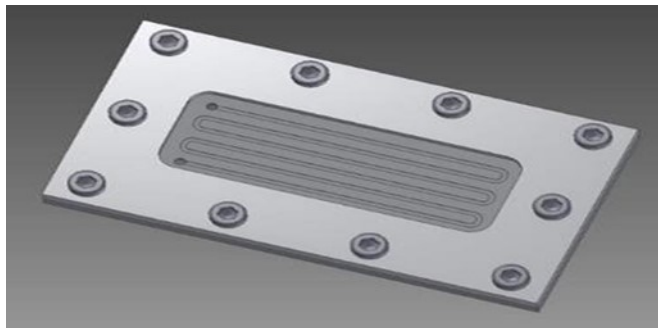
In recent reports, our group demonstrated the high photocatalytic activity either under simulated solar light or under visible light irradiation of doped- or QDs-sensitized metal oxides like ZnO or TiO<sub>2</sub> [11-14]. The developed photocatalysts were demonstrated to have a high activity for the degradation of dyes like Orange II dye or drugs like salicylic acid. Another interesting application is in the field of health with Ifosfamide (IFOS) as an anti-tumor agent [15]. It is a persistent molecule that is difficult to degrade by traditional methods and whose degradation leads to the production of intermediate products.

Photocatalysis is an interesting alternative for the degradation of IFOS but has only scarcely been investigated. OH radicals are the dominant species for the oxidation of IFOS and de-chlorination is the first step in the degradation of the photocatalytic oxidation process [16].

In this paper, the photocatalytic degradation of IFOS using a CuInZn<sub>x</sub>S<sub>2+x</sub> QDs (ZCIS) sensitized ZnO photocatalyst immobilized in a microchannel was investigated under artificial solar irradiation. Our objective is to demonstrate the interest of microreactors for the kinetic study of photocatalytic reactions. The degradation of IFOS under solar irradiation (4 mW/cm<sup>2</sup>) was carried out using a thin layer of catalyst ZnO/ZCIS deposited in a microchannel. In these conditions, it was possible to initiate and follow the photocatalytic reaction. Degradation of IFOS can be fitted by the Langmuir-Hinshelwood model and external diffusion can be characterized by Sherwood (Sh) and Damköhler (Da) numbers.

## 2. MATERIALS AND METHODS

Previously, several microchannel reactors have been tested [17-20] in order to determine the best design of microreactors. The one tested in this research is composed of a microchannel with a serpentine shape engraved in a stainless steel support. Figure 1 shows the shape of the microchannel covered by a rectangular glass.



**Figure 1.** Design of the microreactor with a serpentine microchannel shape with a length of 516 mm.

The fluid flows into the rectangular section of the channel which is 1.5 mm wide, 0.7 mm deep and 516 mm long. The characteristics of the channel geometry and flow are shown in Table 1. At a flow rate of 5 mL/h, the space time is 390s experimentally measured by the crossing time of a bubble between the inlet and the outlet of the microreactor. We have also measured the residence time distribution (RTD) to characterize the reactor. Experiments were realized with the injection of Blue Methylene tracer and absorption of the solution at the outlet. A mean residence time equal to 393s was found which is very close to the observed experimental space time. There was no stagnating zone between catalyst and flow region. Since the Reynolds number is about 1, the flow condition inside the microreactor channel is a laminar flow.

**Table 1** - Geometric and flow characteristics of the channel at 5 mL/h.

Width w (mm)	depth h (mm)	Length L (mm)	Volume V (mm <sup>3</sup> )	Space time t <sub>p</sub> (s)	Average- residence time	Mean velocity u (mm/s)	Re number
1.5	0.7	516	542	390	393	1.32	1.26

## 2.1. Photocatalytic composite

Metal oxides such as ZnO have a wide band gap (3,24 eV) which limits its absorption to the UV region. In order to improve the photocatalytic activity, we combined the ZnO semiconductor with ZCIS to form the ZnO/ZCIS composite photocatalyst. Among the good properties of this composite, there is a separation of the charges (e, h<sup>+</sup>) in order to prevent their recombination [9]. ZnO/ZCIS is an ideal candidate because of its low band gap (2.1 eV) and its low toxicity.

The ZnO/ZCIS photocatalyst was deposited at the bottom of the micro-channel reactor. The amount of deposit was determined by weighing the dried reactor before and after the deposition. Several experiments were performed to determine the minimal mass without limitation of the production of reactive sites. An optimal mass of 40 mg was found. No differences of the conversion were observed with a variation of the concentration above this value.

## 2.2. Photocatalysis experiments

IFOS solution with an initial concentration varying from 2.5 to 10 mg/L was injected by means of a syringe pump through the microreactor with a constant inlet flow rate of 5 mL/h. Even without irradiation, IFOS is adsorbed on photocatalyst when the flow is crossing the channel, which may result in a decrease of the solution concentration. To take into account this effect, we carried out the photodegradation reaction after waiting for 2 h in the dark in order to reach the adsorption equilibrium.

An artificial solar irradiation with Neon, JBL Solar Color T8 15W, was used. The intensity light was measured with a radiometer below the glass cover and just above the photocatalyst. A value of 4.0 mW cm<sup>-2</sup> was found.

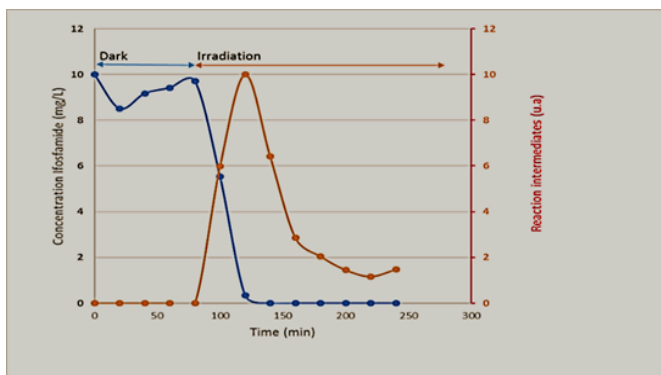
During the step of irradiation, a sample was withdrawn after regular time intervals at the outlet of the microreactor. The solution of the sample was centrifuged at 15, 000 rpm for 2 min and then injected into

an HPLC in order to determine the output concentration of IFOS and then the conversion yield. The signal at 485 nm was used to evaluate the conversion of IFOS.

## 3. EXPERIMENTAL RESULTS

### 3.1. Adsorption and degradation of Ifosfamide

Under dark conditions, the equilibrium of the adsorption of IFOS was achieved after 90 minutes. The concentration profile obtained from the kinetic experiments using the microreactor with ZnO/ZCIS and initial concentration of 10 mg/L is represented on Figure 2. After the beginning of irradiation (t = 90min), there is a fast degradation which occurs in less than 30 minutes. Nevertheless, there is a simultaneous degradation of secondary products for a longer time of 50 minutes.

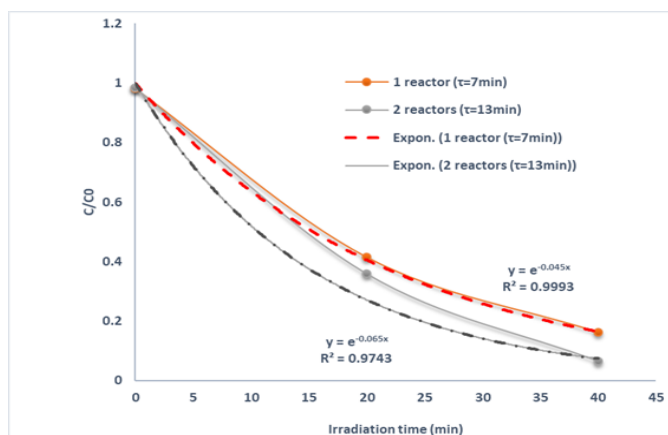


**Figure 2.** Degradation of Ifosfamide and the secondary products in microreactor under solar light ; Irradiation after 90 minutes in the dark (initial time in the dark).

### 3.2. Influence of the residence time on the conversion yield

In this case, the experiments were realized at the same flow rate of 5mL/h and C<sub>0</sub>=10mg/L; the residence time  $\tau$  is adjusted by changing the reactive volume of the microchannel. This was achieved by connecting several microreactors in series. The final phase was a development of an integrated system with several microreactors. This makes it possible to determine the improvement due to a longer reaction time. The Figure 3 illustrates the variation of the con-

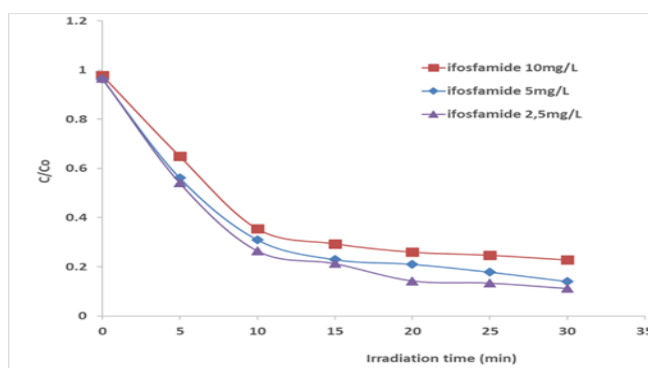
version yield as a function of irradiation time in the case of two reactors in series. After reaching the adsorption–desorption equilibrium in the dark during 90 min (Figure 1),  $C/C_0$  equals 1 and the irradiation time is started ( $t=0$ ). A gradual decrease of the concentration of Ifosfamide is observed as a function of the irradiation time (Figure 3). The efficiencies of Ifosfamide degradation by ZCis is determined quantitatively using the pseudo-first-order model  $C/C_0 = \exp(-k_{app} t)$ , where  $k_{app}$  is the apparent rate constant ( $\text{min}^{-1}$ ) and  $C_0$  and  $C$  are the concentrations of Ifosfamide at time 0 and  $t$ , respectively. The rate constants  $k_{app}$  determined for the degradation of the IFOS solution are found by R-squared values on the trend lines; the values are 0.045 ( $R^2=0.9993$ ) and 0.065  $\text{min}^{-1}$  ( $R^2=0.9743$ ) for one reactor and two reactors in series, respectively.



**Figure 3.** Photocatalytic degradation of Ifosfamide under solar irradiation (light intensity  $I=4 \text{ mW/cm}^2$ ); flow rate  $Q=5 \text{ mL/h}$ ;  $C_0=10 \text{ mg/L}$ ; Influence of the space time when connecting several microreactors in series.

### 3.3. Influence of the initial concentration of Ifosfamide

On the other hand, the effect of the initial IFOS concentration  $C_0$  on the photocatalytic degradation by varying the concentration between 2.5 to 10  $\text{mg/L}$ . The experiments have been performed with one reactor (516 mm). As can be seen in Figure 4, the initial slope of the curve is almost the same but a significant change occurs on the final conversion yield. It increases when the initial concentration of the pollutant decreases.

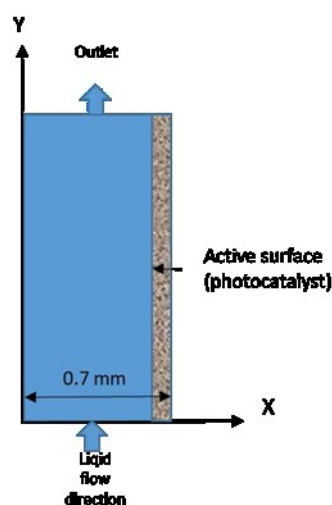


**Figure 4.** Influence of the initial concentration of Ifosfamide on the degradation; Flow rate = 5  $\text{mL/h}$ ;  $I=4 \text{ mW/cm}^2$ .

There is a decrease in the value of the rate constant until the equalization of the adsorption–desorption and kinetic rates. After 30 min of irradiation, the percentage of IFOS concentration levelled off due to the weak amount of concentration remaining in solution and equilibrium rates.

### 3.4. Modeling with COMSOL Multiphysics

In order to predict the time required for the degradation of IFOS and thus the reactor length needed, the IFOS concentration to the reactive surface  $C_s$  was calculated by the Computational fluid dynamics method. The schematic profile in the channel domain is shown in Figure 5.



**Figure 5.** Schematic profile in the channel domain and liquid flow direction.

The model takes into account the convection-diffusion coupling occurring with the flow and the deposited photocatalyst. We introduced an equation reaction rate  $r$  at the photocatalytic surface and a mean value of the outlet concentration was used to estimate the fraction of IFOS remaining ( $C/C_0$ ).

Photocatalytic degradation rate of IFOS can be fitted by the Langmuir-Hinshelwood model with the assumption that the photocatalytic degradation occurs between the substrate and the adsorbed reactive species on the photocatalytic surface.

$$r = k K C / (1 + K C) \quad (1)$$

Where  $k$  the intrinsic reaction rate constant and  $K$  the adsorption constant. By adjustment of the model with experimental values, we determined these two parameters:  $k = 2.98 \cdot 10^{-4} \text{ mol m}^{-3} \text{ s}^{-1}$  and  $K = 4000 \text{ m}^3 \text{ mol}^{-1}$ .

A Newtonian incompressible flow was adopted to simulate the aqueous solution with a laminar flow distribution in the microchannel. The surface concentration  $C_s$  is calculated by the material balance for the surface, including diffusion and a reaction rate.

$$\frac{\partial C_s}{\partial t} + (-D \nabla C_s) = r \quad (2)$$

where  $C_s$  is the surface concentration of IFOS in  $\text{mol m}^{-2}$  and  $r$  the surface reaction rate. In the model, the surface reaction rate  $r$  is considered as a negative boundary flux.

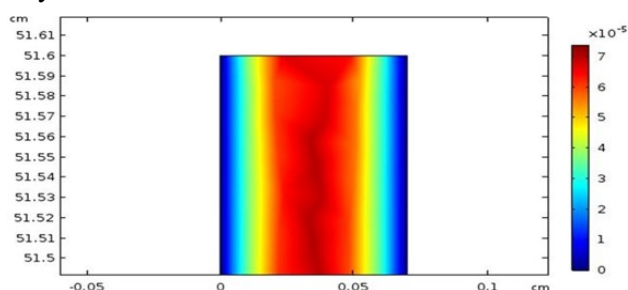
The transport in the bulk of the channel is described by a convection-diffusion equation:

$$\frac{\partial C}{\partial t} + \nabla (-D \nabla C + C u) = 0 \quad (3)$$

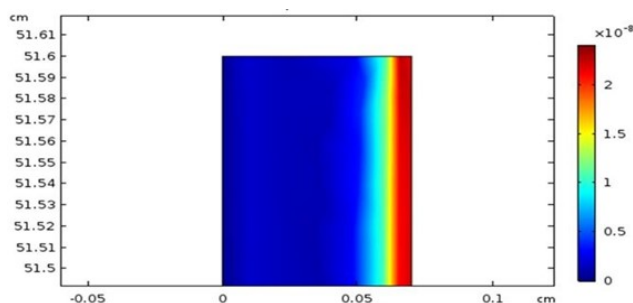
In the above equation,  $D$  denotes the isotropic diffusion coefficient of IFOS in aqueous solution, and  $u$  the flow velocity. In this case, the velocity  $u(x)$  in the perpendicular direction (width of the microchannel)

equals 0 while the velocity in the parallel direction  $u$  ( $y$ ) of the active surface (length of the microchannel) has a Gaussian distribution due to laminar flow distribution. The value of Reynolds number  $Re$  equals 1.3 which corresponds to a laminar flow.

At time  $t = 1$  minute, Figure 6 illustrates the convective flow (component  $y$ ) near the output of the microreactor, and Figure 7 shows the distribution of the diffusive flux inside the microchannel at the same time. There is clearly a concentration gradient between the reaction medium and the catalytic surface corresponding to a limiting thickness of the diffusion layer about 0.2 mm.



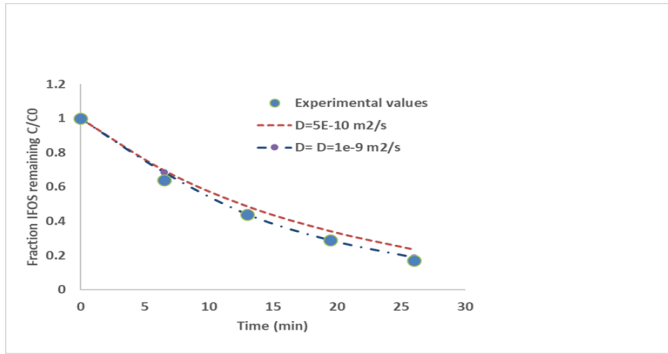
**Figure 6.** Convective flux inside the microchannel with laminar flow ( $Re=1.3$ ); flow rate  $Q=5\text{mL/h}$ .



**Figure 7.** Diffusive flux inside the microchannel and concentration gradient on the right side; flow rate  $Q=5\text{mL/h}$ .

### 3.4.1 Kinetic modeling and comparison with experimental results

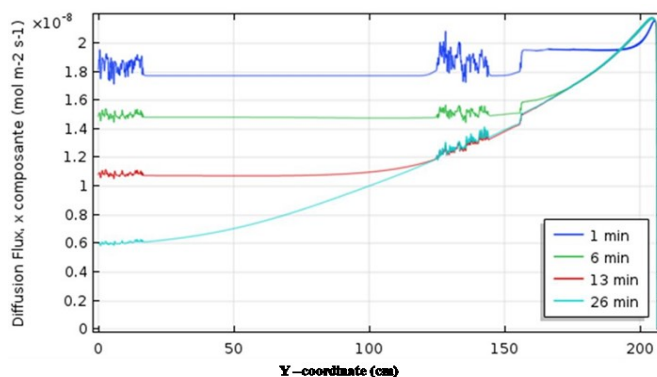
As can be seen on Figure 8, there is a good agreement between experience and model. In our conditions of 5 mL/h, 70 % of conversion was reached in 20 minutes ( $C/C_0=0.3$ ).



**Figure 8.** Comparison of experimental result and modeling based on Langmuir-Hinshelwood kinetic parameters Influence of the diffusion coefficient IFOS towards the photocatalyst; flow rate  $Q=5\text{mL/h}$ .

A small variation of the conversion rate was observed by varying the diffusion coefficient between  $5 \cdot 10^{-10}$  and  $10^{-9} \text{ m}^2 \text{ s}^{-1}$ . The adjustment of the model with the experimental results allows to evaluate the diffusion coefficient of IFOS towards the photocatalyst at about  $1 \cdot 10^{-9} \text{ m}^2 \text{ s}^{-1}$ . This value corresponds to that found in the literature for persistent pollutants like Chlorophenols. For example, Martins et al. [21] find a value of  $1 \cdot 10^{-9} \text{ m}^2 \text{ s}^{-1}$  for the self-diffusion coefficient of 2-chlorophenol in water at 298K in a bulk reactor. Figure 9 shows the diffusion flux which has decreased slightly with the reaction time and that corresponds to the change of concentration gradient over time. The diffusion flux to the catalytic wall compensates for the consumption of the reactants during the passage time in the reactor. In the case of total length of 2m and in our experimental conditions, a diffusion time is evaluated to about 250s.

**Figure 9.** Diffusive flux along the photocatalytic sur-



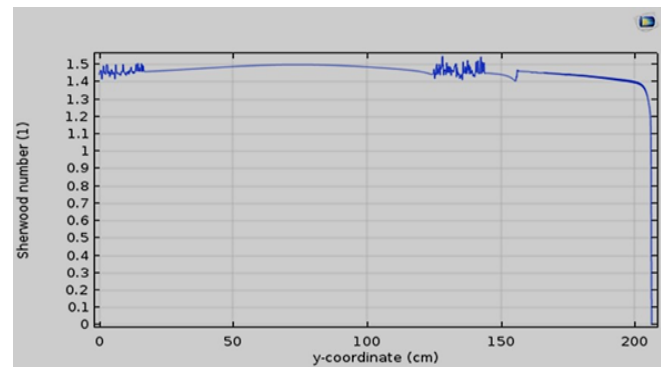
face and influence of the reaction time during the flow in the reactor.

### 3.4.2 Mass transport toward catalytic surface

The Sherwood number is a function of the hydraulic channel diameter  $d_h$ , the diffusion coefficient  $D$  and the mass transfer coefficient  $k_d$ . The hydraulic diameter and the diffusion coefficient are constant whereas the mass transfer coefficient is a function of the concentration of IFOS. As fluid flows along the channel in  $y$ -direction, the concentration gradient in  $x$ -direction due to the photocatalytic reactions leads to a mass transfer characterized by the mass transfer coefficient  $k_d = Sh \cdot d_h / D$ .

In the Figure 10, Sherwood number is plotted versus coordinate  $y$  at a given time ( $t=20 \text{ min}$ ). Based on the results of Figure 10, one can estimate the value of  $Sh$  number around 1.45 which is in agreement with previous study [20] but slightly lower than the value of 6 found previously.

**Figure 10.** Sherwood number along the catalytic sur-



face with a length of 200 cm.

From the value of Sherwood number, we deduce a mass transfer  $k_d = 7.6 \cdot 10^{-6} \text{ m s}^{-1}$ .

To evaluate the limitation by mass transfer we used the number of Damköhler  $Da$  which represents the ratio between the rate of reaction and the rate of transfer [17].

$$Da = k / (k_d \cdot a / K) \quad (4)$$

In which  $a$  is the specific area = Catalytic surface / Volume of microchannel. The value of  $Da$  around

100 shows that the rate of reaction under our experimental conditions is greater than the transfer of matter and therefore the rate of transfer limits the photocatalytic reaction.

#### 4. CONCLUSION

A complete degradation of Ifosfamide and the by-products have been achieved with the analysis of mass transfer and reaction in microchannel reactor. The adjustment of the model with the experimental results allows to evaluate the kinetics constants of the Langmuir-Hinshelwood model ( $k$ ,  $K$ ) and diffusion coefficient of Ifosfamide ( $D$ ). The model makes it possible to predict a conversion percentage of 70% in 20 minutes in a microreactor with a total length of 1.6 m. The Sherwood number and the average value of Damköhler number show that the photocatalytic reaction under our experimental conditions is limited by the mass transfer. There is clearly a concentration gradient between the reaction medium and the catalytic surface corresponding to a limiting thickness of the diffusion layer. CFD modeling allows a perfect control of the process parameters and the use of microreactors is promising for the degradation of pharmaceuticals in water.

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#### Symbols Used

$a$	$[m^{-1}]$	specific area
$C$	$[mol L^{-1}]$	IFOS concentration
$C_0$	$[mol L^{-1}]$	initial IFOS concentration
$C_s$	$[mol m^{-2}]$	IFOS surface concentration
$D$	$[m^2 s^{-1}]$	diffusion coefficient
$Da$	$[-]$	Damköhler number
$dh$	$[m]$	channel hydraulic diameter
$h$	$[m]$	channel depth
$K$	$[m^3 mol^{-1}]$	adsorption constant
$k$	$[mol m^{-3} s^{-1}]$	intrinsic reaction rate constant
$k_0$	$[min^{-1}]$	initial pseudo first order reaction rate constant
$kap$	$[min^{-1}]$	pseudo first order reaction rate constant

$kd$	$[m s^{-1}]$	mass transfer coefficient
$L$	$[m]$	channel length
$Q$	$[m^3 s^{-1}]$	volumetric flow rate
$Re$	$[-]$	Reynolds number
$t$	$[s]$	time duration
$t_0$	$[s]$	initial time: starting irradiation
$t_p$	$[s]$	space time $V/Q$
$u$	$[m s^{-1}]$	mean mass velocity
$V$	$[m^3]$	channel volume
$w$	$[m]$	channel width

#### Greek letters

$\tau$	$[s]$	average residence time
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#### Subscripts

0	initial (starting irradiation)
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#### Abbreviations

CFD	computational fluid dynamics
IFOS	Ifosfamide
OH	radicals
QDs	quantum dots
$TiO_2$	Titanium dioxide
ZCIS	$CuInZn_xS_{2+x}$ (quantum dots)
ZnO	Zinc Oxide
ZnO/ ZCIS	composite
ZnS	Zinc sulfide

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