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1 Influence of activation conditions on textural
2 properties and performance of activated biochars
3 for pyrolysis vapors upgrading

4 *Christian Di Stasi^{1*}, Gianluca Greco¹, Rafael L. S. Canevesi², M. Teresa Izquierdo³, Vanessa*
5 *Fierro², Alain Celzard², Belén González¹, Joan J. Manyà¹*

6 ¹ Aragón Institute of Engineering Research (I3A), Thermochemical Processes Group,
7 University of Zaragoza, Technological College of Huesca, Crta. Cuarte s/n, 22071 Huesca,
8 Spain

9 ² Université de Lorraine, CNRS, IJL, Épinal, F-88000, France

10 ³ Instituto de Carboquímica (ICB-CSIC), Miguel Luesma Castán 4, Zaragoza, E-50018, Spain

11

12 * Corresponding author at: Aragón Institute of Engineering Research (I3A), Thermochemical
13 Processes Group, University of Zaragoza, Technological College of Huesca, Crta. Cuarte s/n,
14 22071 Huesca, Spain

15 E-mail address: christiandistasi@unizar.es.

16 **ABSTRACT**

17 The main aim of the present study is to provide a comprehensive assessment of the effects of
18 process activation conditions on the textural properties of the resulting activated carbons,
19 which were produced from wheat straw-derived biochar through chemical activation (with
20 K_2CO_3 at different pressures and mass impregnation ratios) and physical activation (with CO_2
21 at different temperatures and pressures). For chemically activated biochars, it was found that
22 specific surface area and pore size distribution were both only positively affected by
23 increasing the carbonate loading. However, physically activated biochars produced at the
24 highest pressure and lowest temperature (1.0 MPa and 700 °C) had the highest surface areas
25 and widest pore size distributions. The materials with the most appropriate textural properties
26 were then tested as catalysts for steam and dry reforming of the aqueous phase of pyrolysis
27 oil. The best catalytic performance (a total gas yield of 74% and a selectivity toward H_2 of
28 almost 40%) was observed for a physically activated biochar. This good performance was
29 ascribed to the high availability of K^0 on the catalyst surface, which could effectively
30 promote the reactions involved in the upgrading process.

31 **Keywords**

32 Biochar; chemical activation; K_2CO_3 ; physical activation; pressure; steam and dry reforming
33 of pyrolysis oil

34

35 **1. Introduction**

36 In recent years, the demand for activated carbons has increased due to their usefulness in a
37 wide range of different applications such as gas and liquid purification [1], electrochemistry
38 [2], soil remediation [3] and catalysis [4]. Thanks to their versatility to be adapted to specific
39 applications, these materials are considered promising candidates to address environmental
40 issues related to global warming and pollution [5]. Nevertheless, the current main drawback
41 is that the production of activated carbons is still partly based on fossil fuels, which do not
42 meet sustainability criteria [6,7]. Hence, much more efforts should be made to produce much
43 more carbons from biomass. However, the direct conversion of biomass feedstock into the
44 final product has a low yield. Thus, the production of activated carbons from biochar
45 produced by slow pyrolysis seems to be more appropriate in terms of scalability, while at the
46 same time these value-added materials can strengthen the value chain of existing biochar
47 production systems.

48 In general, pristine biochar has neither a well-developed surface area nor a hierarchical
49 pore size distribution, which is mainly dominated by narrow pores with a diameter (d_p) lower
50 than 0.7 nm (ultra-micropores). Hence, a subsequent activation step is required to improve
51 the textural properties of the carbon material. The activation process involves the
52 development and opening of the porosity of a char using an activation agent. Depending on
53 the agent used, the procedure can be called chemical or physical activation. Chemical
54 activation usually involves two steps. First, the biochar is impregnated with an aqueous
55 solution of the chemical activation agent, or mixed with it in the dry state, and, in a second
56 step, the blend is heated up to a given temperature at which oxidation, dehydration,
57 aromatization and crosslinking reactions, among others, occur. Although the most widely
58 used chemical reagents are KOH [8] and H₃PO₄ [9], non-hazardous and relatively cheap
59 alternative compounds such as K₂CO₃ [10,11] have recently attracted considerable attention.

60 Mai *et al.* [12] have recently reported that activation with potassium carbonate can result in
61 carbons with a high percentage of structural defects and a well-balanced porosity between
62 micro- and mesopores.

63 When activation is carried out by exposing the precursor to relatively high temperatures
64 under an oxidizing atmosphere (e.g., CO₂ [13], H₂O [14] and O₂ [15]), the process is called
65 physical activation. Through either physical or chemical activation, it is possible to tune the
66 textural features of the starting biochar by properly adjusting the activation conditions,
67 mainly temperature, type and/or concentration of activation agent, as well as pressure. To the
68 best of our knowledge, although both activation procedures are widely reported in the
69 literature, a thorough study on the effects that the activation parameters and their possible
70 interactions have on the textural properties of the resulting activated biochar, has never been
71 reported so far. Such information could be very helpful to properly establish the most
72 appropriate operating conditions to produce engineered carbon materials from biomass.

73 One of the most interesting fields of application of activated carbons is the upgrading of
74 raw pyrolysis vapors, which contain both permanent gases (e.g., CO₂, CO, CH₄, and H₂) and
75 condensable compounds. The condensable fraction, which is generally called pyrolysis oil or
76 bio-oil, is a mixture of hundreds of organic compounds (alcohols, ketones, acids, etc.) and
77 water, and its composition strictly depends on the biomass composition and the pyrolysis
78 operating conditions [16]. For biochar production systems based on slow pyrolysis,
79 downstream processes aimed at upgrading pyrolysis vapors are required to avoid undesirable
80 condensation of organic compounds and to increase simultaneously the quality of the gaseous
81 product. For this purpose, combined steam and dry reforming of pyrolysis oil is a promising
82 option [17,18], due to the presence of a relatively high amount of CO₂ and steam in the raw
83 vapor phase. Heterogeneous catalysts based on transition metals such as Ni [19,20], Co [21],
84 Pt and Rh [22] can be used to improve the overall conversion and the selectivity of the

85 products. However, the main drawback of these catalysts is that the relatively expensive
86 active phase can easily be poisoned and/or deactivated by deposition of coke [23]. One
87 possible solution to reduce the overall cost of the upgrading process is to use activated
88 biochar as catalyst. Its porous structure and inherent inorganic contents (especially K, Mg and
89 Ca) could result in a relatively good catalytic activity [24–29]. Furthermore, due to the
90 presence in the reaction system of H₂O and CO₂, the carbonaceous support, as well as coke
91 deposits, could be continuously partly gasified, thus creating new pores to avoid deactivation.
92 In addition, the spent biochar can be burned to recover energy [30] or, depending on the
93 metal loaded on the support, employed as soil conditioner [31].

94 Keeping in mind all the above, the main objective of this study was to perform a
95 preliminary investigation about the effects of several activation conditions on the textural
96 properties and the catalytic activity of the resulting activated biochars. To this end, wheat
97 straw-derived biochar was activated chemically and physically with K₂CO₃ and CO₂,
98 respectively, under different operating conditions. To objectively assess the effect of the
99 selected operation conditions (temperature and mass ratio K₂CO₃/precursor for chemical
100 activation, and temperature and pressure for CO₂ activation), we implemented two factorial
101 designs of experiments, with two factors and three replicates at the center point. The specific
102 surface areas and pore size distributions of the resulting activated biochars were then
103 evaluated. The most promising activated biochars were finally tested as catalysts for steam
104 and dry reforming of pyrolysis oil. Their performance was evaluated in terms of conversion
105 of liquid into gas, product selectivity and resistance to deactivation.

106 **2. Experimental Section**

107 This study was divided into two main steps: in the first, we carried out physical and
108 chemical activations under different process conditions in order to study their influence on
109 the pore size distribution and specific surface area of the resulting activated biochars; in the

110 second part, the most promising materials were then tested as catalysts in pyrolysis vapors
111 upgrading experiments. The methodology used in this study is summarized graphically in
112 Fig. A.1.

113 ***2.1. Biochar production***

114 The biochar used in this study was produced by slow pyrolysis of binder-free wheat straw
115 pellets (9 mm OD and 10–13 mm long). Pyrolysis was carried out at atmospheric pressure in
116 a fixed bed reactor, which was placed inside a furnace and heated up to 500 °C at an average
117 heating rate of 5 °C min⁻¹ and using N₂ as carrier gas. More details on the pyrolysis device
118 and the experimental procedure are available elsewhere [32]. The resulting biochar (i.e.,
119 “pristine”) was ground and then thoroughly sieved to obtain particle sizes in the range of
120 0.212 to 1.41 mm. The pristine biochar was characterized by proximate analysis (performed
121 in quadruplicate according to ASTM standards D3173 for moisture, D3174 for ash, and
122 D3175 for volatile matter) and ultimate analysis by means of an elemental analyzer CHN628
123 from Leco Corporation (USA).

124 ***2.2. Activation of pristine biochar***

125 *2.2.1. Chemical activation*

126 For chemical activation, the pristine biochar was first impregnated with a 1 mol L⁻¹
127 aqueous solution of K₂CO₃. Three K₂CO₃: biochar mass impregnation ratios (1:1, 2:1 and
128 3:1) were achieved by adjusting the volume of solution. The heterogeneous mixture was then
129 stirred for 2 h at 50 °C, filtered and dried overnight at 110 °C to remove the residual water.
130 Afterwards, 10 g of the impregnated samples were heated up to 700 °C at a heating rate of 10
131 °C min⁻¹, under an inert atmosphere (N₂), and at three different values of absolute pressure
132 (0.10, 0.55 and 1.00 MPa). For this purpose, a tubular fixed bed reactor (made of nickel-
133 chromium alloy UNS N06600, 28.1 mm ID and 600 mm long) placed in a vertical furnace
134 (model EVA 12/300 from Carbolite Gero, UK) was used. The relatively low activation

135 temperature was chosen to avoid the evaporation of the metallic potassium derived from the
136 decomposition of the activation agent and, also, to ensure low activation extensions and
137 clearly see the potential effects of the studied parameters. On the other hand, the range of
138 activating pressures was chosen in order to allow this experimental setup to be easily scaled
139 up, since the involved pressures were not too high (up to 1.0 MPa). A soaking time at the
140 highest temperature of 60 min was set. The pressure within the reactor was adjusted using a
141 downstream servo-controlled regulator valve. The gas hourly space velocity (GHSV) at the
142 activation temperature was estimated to be 7000 h^{-1} , considering the pressure applied and a
143 bed void factor of 0.5. Thus, the mass flow rate of the inlet gas stream (N_2) was properly
144 adjusted as a function of the selected pressure and the highest temperature to achieve the
145 aforementioned GHSV value.

146 As a last step, the carbons were rinsed to remove the unreacted reagent and other
147 impurities from their surface. In this work, two different washing procedures were adopted to
148 assess possible effects on the resulting surface area and the catalytic activity of the activated
149 carbons produced. For this purpose, among all the unwashed chemically activated biochars,
150 four of them were just washed with hot deionized water ($100 \text{ }^\circ\text{C}$), while the seven chemically
151 activated biochars adopted in the design of experiments were washed with a 0.25 mol L^{-1}
152 solution of HCl followed by hot water. Both washing procedures were carried out until
153 neutral pH. The resulting activated biochars were then dried overnight at $110 \text{ }^\circ\text{C}$. The
154 efficiency of the washing step was evaluated according to the washing yield, calculated from
155 Eq. 1, in which m_c and m_w are the masses of activated carbon before and after the washing
156 steps, respectively.

$$157 \quad Y_{wash} = \left(1 - \frac{m_c - m_w}{m_c}\right) 100 \quad (1)$$

158 2.2.2. Physical activation

159 Physically activated biochars were produced under an atmosphere of pure CO₂ at three
160 different temperatures(700, 775 and 850 °C), and at three different absolute pressures (0.10,
161 0.55, and 1.00 MPa). The selected activation temperatures are within the range commonly
162 reported in the literature, whereas the pressure values were established with the purpose to
163 compare the properties of resulting activated carbons produced under atmospheric and
164 moderate pressures. Using the same device described in the previous section, 10 g of pristine
165 biochar were heated under N₂ atmosphere, at a heating rate of 10 °C min⁻¹, until the target
166 temperature was reached. Then, the gas supply was switched from N₂ to CO₂ at a constant
167 GHSV of 7000 h⁻¹. These conditions were maintained during the time required to reach a
168 degree of burnout (η , defined as the percentage of mass loss) in the range of 30% to 60%.

169 2.2.3. Design of experiments

170 To objectively assess the effects of the activation conditions on the textural properties of
171 activated biochars, an unreplicated two-level factorial design of experiments (with two
172 factors and three replicates at the center point) was adopted for both chemical and physical
173 activations. For chemical activation, the analyzed factors were the absolute pressure and the
174 impregnation ratio of K₂CO₃ to raw biochar, whereas in the case of physical activation, the
175 temperature and the absolute pressure were the studied factors. The structure of the
176 regression model used in the statistical analysis was the following:

$$177 \quad \hat{y} = \beta_0 + \beta_i A + \beta_j B + \beta_{ij} AB \quad (2)$$

178 where A and B corresponded to the assessed factors (normalized values in the range from -1
179 to 1), whereas β_0 , β_i , and β_{ij} were the intercept, linear, and interaction coefficients,
180 respectively. The results obtained for the selected response variables (y) were analyzed using
181 the Minitab v17 statistical package. A significance level of 5% was assumed and the adjusted
182 coefficient of determination (R^2_{adj}) was taken as an indicator of the quality of the fit. Table 1
183 summarizes the adopted designs of experiments and lists the names of the activated biochars

184 produced. Briefly, activated biochar are referred as X_Y_P . X corresponds to the activation
185 procedure (CB or PB for *chemically activated biochar* and *physically activated biochar*,
186 respectively); Y is the activation temperature for PBs or the impregnation ratio for CBs; and P
187 is the activation pressure.

188 ***2.3. Catalytic pyrolysis vapors upgrading***

189 The liquid feed used here for pyrolysis vapors upgrading tests was the filtered aqueous
190 phase of the pyrolysis oil formed during the production of pristine biochar derived from
191 wheat straw. This resulting liquid product was characterized in terms of elemental
192 composition, using the same CHN analyzer as mentioned in section 2.1, and water content,
193 using a volumetric Karl-Fischer titrator Titrino plus 870 from Metrohm (Switzerland).

194 A schematic overview of the experimental device used for the reforming tests is shown in
195 Fig. A.2 (Appendix A). The reactor and furnace elements were the same as those previously
196 described in section 2.2. The reforming operating conditions were selected based on the
197 findings of a previous study [33], which aimed to establish the best process conditions to
198 minimize deactivation by coke deposition.

199 Briefly, activated biochar (2–5 g) was loaded into the reactor and heated under N_2
200 atmosphere at an absolute pressure slightly above 0.1 MPa. Once the bed temperature
201 reached the target value of 750 °C, CO_2 was added to the inlet gas stream at a partial CO_2
202 pressure of 0.02 MPa. The liquid feed was then injected into the gas stream at the inlet of the
203 reactor by means of a HPLC pump (model 521 from Analytical Scientific Instruments, USA).
204 The liquid hourly space velocity (LHSV) was kept constant and equal to 2 h^{-1} for all the
205 experiments conducted. The duration of the experiments was 60 min. The reactor outlet
206 stream, consisting of permanent gases, unreacted pyrolysis oil, and condensable side
207 products, was forced to pass through an ice-bath condensation train. The permanent gases
208 were analyzed using a dual-channel micro gas chromatograph (μ -GC 490 from Agilent,

209 USA) equipped with TCD detectors and two analytical columns (a Molsieve 5 A and a
 210 PolarPlot U). The known amount of N₂ fed was used as a tracking compound to calculate the
 211 yield of produced gas.

212 **Table 1.** Matrix of the factorial designs adopted to assess the effects of the selected factors
 213 during both chemical and physical activations. For example, a physical activation carried out
 214 at A=-1 and B=1 means that the process was performed at 700°C and 1.0 MPa

Chemical activation (thermal treatment under N ₂ up to 700 °C and 1 h soaking time)			
Level	Factors		Response Variables <i>S_{2D-NLDFT}, V_b, V_{ultra}, V_{micro}, V_{meso}</i>
	A Pressure (MPa)	B Mass ratio (K ₂ CO ₃ : raw biochar)	
Low (-1)	0.10	1:1	
Middle (0)	0.55	2:1	
High (+1)	1.00	3:1	
Runs			Activated biochar designation (all washed with an acidic solution)
1	0	0	CB_2_0.55
2	-1	+1	CB_3_0.10
3	+1	+1	CB_3_1.00
4	0	0	CB_2_0.55
5	0	0	CB_2_0.55
6	-1	-1	CB_1_0.10
7	+1	-1	CB_1_1.00
Physical activation			
Level	Factors		Response Variables <i>S_{2D-NLDFT}, V_b, V_{ultra}, V_{micro}, V_{meso}</i>
	A Temperature (°C)	B Pressure (MPa)	
Low (-1)	700	0.10	
Middle (0)	775	0.55	
High (+1)	850	1.00	
Runs			Activated biochar designation
1	-1	-1	PB_700_0.10
2	-1	+1	PB_700_1.00
3	+1	+1	PB_850_1.00
4	0	0	PB_775_0.55
5	0	0	PB_775_0.55
6	+1	-1	PB_850_0.10
7	0	0	PB_775_0.55

215

216

217 The performance of the activated biochars tested was measured in terms of total gas yield
218 (Y_g) and hydrogen yield (Y_{H_2}), as defined by Eqs. (3) and (4), respectively.

219

$$220 \quad Y_g = \frac{m_g}{m_L} 100 \quad (3)$$

$$221 \quad Y_{H_2} = \frac{F_{H_2,av}}{F_{H_2,stoi}} 100 \quad (4)$$

$$222 \quad F_{H_2,stoi} = \left(2n + \frac{m}{2} - k\right) F_{BO} \quad (5)$$

$$223 \quad S_i = \frac{n_i}{n_{tot}} 100 \quad (6)$$

224 In Eq. (3), m_g was the cumulative mass of the gas produced during the 60-min
225 experiments, whereas m_L corresponded to the total mass of the liquid fed. In Eq. (4), $F_{H_2,av}$
226 was the experimental average molar flow rate of hydrogen from minute 20 to minute 40.
227 $F_{H_2,stoi}$, which was calculated according to Eq. (5), corresponded to the stoichiometric molar
228 flow rate of hydrogen considering the contribution from steam reforming reaction. F_{BO} was
229 the molar flow rate of the dry pyrolysis oil fed. The selectivity towards specific gaseous
230 products (S_i) was calculated according to Eq. (6), in which n_{tot} was the sum of the gaseous
231 products in moles at the outlet stream.

232 **2.4. Characterization of carbon materials**

233 The textural properties of pristine and activated biochars were determined from both N₂
234 and CO₂ adsorption isotherms at -196 °C and 0 °C, respectively. Around 120 mg of sample
235 were degassed under vacuum at 150 °C. ASAP 2020 and ASAP 2420 automatic adsorption
236 analyzers (Micromeritics, USA) were used. The results obtained from the isotherms were
237 treated using the MicroActive software. We determined the BET area (A_{BET}) by application of
238 the BET model, as well as the Gurvitch volume ($V_{0.97}$). The enhanced 2D-NLDFT model [34]
239 was employed using SAIEUS software (available at www.nldft.com) to evaluate the pore size
240 distributions (PSDs) and other related parameters: surface area ($S_{2D-NLDFT}$), ultra-micropore

241 volume (V_{ultra} , $d_p < 0.7$ nm), micropore volume (V_{micro} , $d_p < 2$ nm), total pore volume (V_t) and
242 mesopore volume (V_{meso} , d_p in the range of 2 to 50 nm). The latter was calculated as the
243 difference between V_t and V_{micro} .

244 To evaluate the availability of residual potassium carbonate on the surface of chemically
245 activated biochars, Fourier-transform Infrared Spectroscopy (FTIR) analyses were conducted
246 using a Frontier MIR/NIR Spectrometer from PerkinElmer (USA).

247 Qualitative temperature-programmed desorption (TPD) analyses of pristine biochar and
248 fresh catalysts used in reforming tests were also performed to assess the thermal stability of
249 these samples. TPD measurements were carried out using a thermogravimetric analyzer
250 coupled with a mass spectrometer (STA 449 F3 and QMS 403 Aëolos Quadro from Netzsch,
251 Germany), by heating the sample up to 750 °C at a heating rate of 5 °C min⁻¹ under an Ar
252 atmosphere.

253 A FEI XL30 SFEG scanning electron microscope coupled with energy-dispersive X-ray
254 spectroscopy (Oxford Instrument EDS SDD XMAX detector) was used to characterize the
255 surface of fresh and spent activated biochars. The secondary electron images were taken with
256 an acceleration voltage of 3 or 5 kV to investigate the topology of the samples surface, while
257 an acceleration tension of 10 kV was applied to carry out the chemical mapping of the
258 materials by EDX analysis. Additionally, X-ray photoelectron spectroscopy (XPS) data of
259 spent and fresh catalysts were recorded using an ESCAPlus OMICROM system equipped
260 with a hemispherical electron energy analyzer, following the procedure detailed elsewhere
261 [35].

262 **3. Results and discussion**

263 The results of the textural characterization as well as the proximate and ultimate analyses
264 of the pristine biochar are summarized in Table A.1. The reported values of A_{BET} and S_{2D} .

265 *NLDFT* confirm that the porosity of the pristine biochar, as previously stated, was characterized
266 by a very high fraction of ultra-micropores.

267 **3.1. Chemical activation**

268 Although the use of K_2CO_3 as catalyst in coal gasification processes has been widely
269 reported in the literature [10–12,33], the mechanism explaining its interaction with the carbon
270 matrix still remains unclear. The main reactions that probably occur during chemical
271 activation of a carbonaceous material are the following:



276 The increase in specific surface area could be ascribed to the CO_2 released from carbonate
277 decomposition (Reaction 7), which can diffuse into the solid structure and subsequently react
278 with carbon, thereby stimulating the production of CO by the reverse Boudouard reaction
279 (Reaction 8) and creating vacancies in the solid structure [36]. Despite the fact that the
280 decomposition of K_2CO_3 mainly occurs at temperatures above 900 °C, its contact with the
281 carbon matrix can promote its decomposition at relatively low temperatures [37,38]. The
282 potassium oxide resulting from the decomposition of the carbonate, or the carbonate itself,
283 can also react with the carbon-based material through Reaction 9 and 10 towards CO and
284 metallic potassium [39].

285 The assessed textural properties of chemically activated biochars are listed in Table 2. Fig.
286 1 shows the normalized plots of the standardized effects for the response variables
287 considered: $S_{2D-NLDFT}$, V_t , V_{ultra} , V_{micro} , and V_{meso} . The PSDs obtained for biochars activated
288 with different impregnation ratios and the N_2 adsorption isotherms of such samples are shown

289 in Figs. A.3a and A.4, respectively. More detailed statistical outcomes for the results reported
290 in the present section can be found in Table A.2.

291

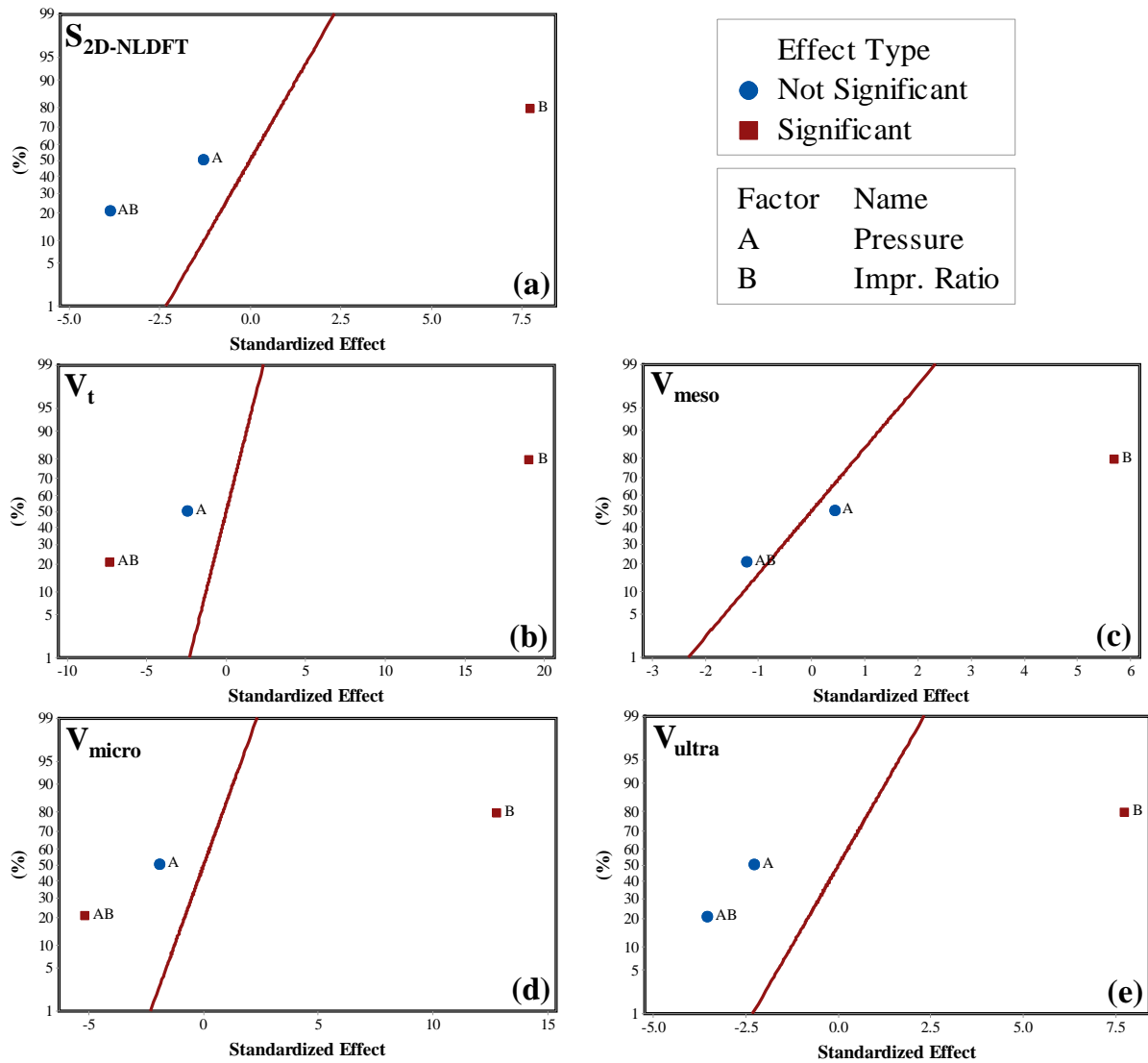
292 **Table 2.** Results of textural characterization and washing efficiency of chemically activated
 293 biochars

Material	Specific surface area		$V_{0.97}$	V_t	Pore volumes			Washing yield Y_{wash} %
	A_{BET}	$S_{2D-NLDFT}$ $m^2 g^{-1}$			V_{ultra}	V_{micro}	V_{meso}	
CB_1_0.10	170	712	0.078	0.181	0.143	0.175	0.006 (3.19 %) ^a	77.0
CB_1_0.10_W ^b	142	700	0.068	0.180	0.135	0.169	0.011 (6.14 %) ^a	77.1
CB_1_1.00	180	761	0.081	0.195	0.149	0.188	0.007 (3.82%) ^a	76.4
CB_1_1.00_W ^b	100	700	0.049	0.168	0.144	0.160	0.008 (5.01%) ^a	78.1
CB_3_0.10	413	933	0.181	0.259	0.197	0.246	0.013 (4.96%) ^a	54.3
CB_3_0.10_W ^b	368	893	0.158	0.242	0.193	0.229	0.013 (5.35%) ^a	56.0
CB_3_1.00	337	834	0.148	0.230	0.169	0.218	0.012 (5.23%) ^a	56.9
CB_3_1.00_W ^b	316	851	0.138	0.229	0.180	0.217	0.012 (5.40%) ^a	60.6
CB_2_0.55	316	900	0.137	0.238	0.180	0.228	0.009 (3.92%) ^a	58.1
CB_2_0.55	310	872	0.136	0.233	0.177	0.222	0.011 (4.63%) ^a	54.4
CB_2_0.55	307	864	0.137	0.233	0.171	0.221	0.011 (4.86%) ^a	55.5

294 ^a Calculated as $\frac{V_{meso}}{V_t} 100$

295 ^b Carbons just washed with water that were not included in the design of experiments.

296



297 **Fig. 1.** Normal plots of standardized effects ($\alpha = 0.05$) for chemically activated biochars:
 298 specific surface area (a); total pore volume (b); mesopore volume (c); micropore volume (d);
 299 ultra-micropore volume (e).

300 Regarding the results obtained in terms of specific surface area ($S_{2D-NLDFT}$), Fig. 1a clearly
 301 shows that a higher concentration of activation agent led to a significant increase in the
 302 porosity of the resulting activated biochars. In contrast, the absolute pressure proved to be
 303 irrelevant for the values of final specific surface area. Our results regarding the development
 304 of porosity when higher impregnation ratios were used confirm the relevance of K_2CO_3 as a
 305 chemical reagent for the production of tailored activated carbons. All pore volumes analyzed
 306 in the present study were also found to be strongly dependent on the amount of activating

307 agent (see Fig. 1). The chemical activation process was able to slightly broaden the original
308 pore size distribution of the pristine biochar, leading to a more hierarchical porous structure
309 with a large contribution of ultra-micropores and a slightly increased amount of mesopores.
310 Almost all the carried out N_2 isotherms required more than 70 h to be completed, thus
311 confirming the predominant narrow microporous structure of chemically activated biochars
312 (see Fig. A.3a), which hindered the diffusion of N_2 within the porosity. For both total and
313 micropore volumes, a statistically significant combined effect of pressure and impregnation
314 ratio (AB) was found. However, these effects were comparatively much weaker than that
315 observed for the main effect of the impregnation ratio (B). From a thermodynamics point of
316 view, an increased pressure shifts the equilibrium of the activation reactions to the left. The
317 fact that the textural properties of the resulting activated carbons were practically
318 independent of pressure could indicate that the chemical activation process was mainly
319 kinetically controlled.

320 It should be pointed out that the overall curvature terms for $S_{2D-NLDFT}$, V_t and V_{micro} were
321 statistically significant (p -values below 0.05, as reported in Table A.2). This indicates that, in
322 further studies, our adopted factorial design should be expanded to a central composite design
323 to be able to evaluate the pure quadratic regression coefficients and then apply Response
324 Surface Methodology (RSM) for optimization purposes.

325 As shown in Table 2, after the carbons were washed only with hot water, there was a
326 slight decrease in both the specific surface area and the micropore volume, compared to those
327 measured for carbons washed with the acidic solution. This can be explained by the fact that
328 water could not completely remove the remaining carbonate (and other chemical species)
329 from the solid surface, thus causing blockage of some micropores. By focusing on the
330 difference between the washing procedures of activated carbons at the same loadings of
331 K_2CO_3 , the results reported in Table 2 indicate that a higher activation pressure led to an

332 increased amount of product not soluble in water, i.e., to higher Y_{wash} values. This finding
333 agrees with results reported by Malekshahian *et al.* [40], who concluded that high gasification
334 pressures can limit the volatilization of potassium.

335 **3.2. Physical activation**

336 It must be pointed out that an increase in absolute pressure also implied an increase in the
337 partial pressure of the reactant (CO_2). In the literature, the effects of absolute pressure and
338 partial pressure of reactant are commonly assessed separately. Nevertheless, and as far as we
339 know, no previous studies have addressed the question of whether activation under pressure
340 can cause textural changes in the resulting physically activated biochars. Activation of
341 biomass-derived carbon had only been studied using supercritical water at high pressure [41],
342 but such conditions are too different from the present ones for allowing some comparison to
343 be done.

344 Table 3 lists the textural properties of the activated carbons produced, whereas Fig. 2
345 shows the normal plots of the standardized effects obtained for each response variable. More
346 detailed statistical outcomes are reported in Table A.3. The most relevant PSDs and N_2
347 adsorption isotherms are shown in Figs. A.3b and A.5, respectively. As previously mentioned
348 in section 2.2.2, the activation times (also shown in Table 3) were varied to obtain degrees of
349 burnout in the range of 30% to 60 %, since the gasification rate was strictly dependent on
350 reaction temperature and reactant partial pressure. As expected, the reaction rate was minimal
351 at the lowest temperature (700 °C). In fact, and whatever the pressure applied, an activation
352 time of 180 min was required to reach the desired reaction extent. For activation temperatures
353 of 775 and 850 °C, an activation time of 60 min was enough, except for the material
354 PB_850_1.00, for which this reaction time led to an almost complete gasification of the
355 sample ($\eta = 85\%$). Hence, the activation time at the highest levels of temperature and
356 pressure was finally set at 30 min to obtain a degree of burnout of 57%.

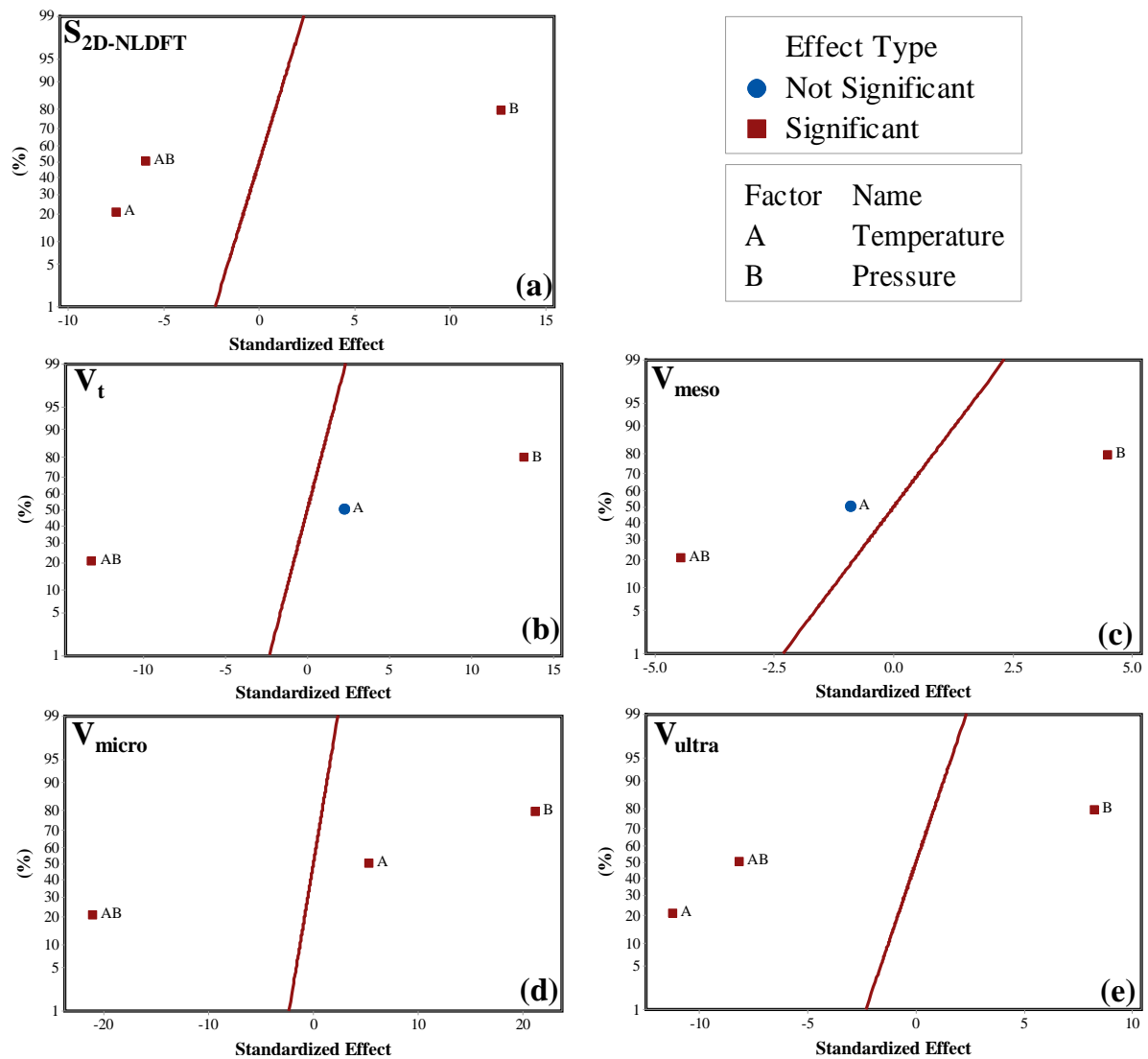
357 **Table 3.** Degrees of burnout activation times and textural properties for physically activated
 358 biochars

Material	Activation time	Burnout η %	Specific surface area		Pore volumes				
	t_{act} min		A_{BET}	$S_{2D-NLDFT}$ $m^2 g^{-1}$	$V_{0.97}$	V_t	V_{ultra}	V_{micro} $cm^3 g^{-1}$	V_{meso}
PB_700_0.10	180	30.1	552	833	0.237	0.243	0.184	0.227	0.016 (6.77%) ^c
PB_700_1.00	180	52.8	743	1008	0.333	0.333	0.226	0.301	0.032 (9.55%) ^c
PB_850_0.10	60	54.9	600	818	0.263	0.266	0.169	0.240	0.026 (9.64%) ^c
PB_850_1.00	30	56.7	688	881	0.298	0.296	0.176	0.273	0.022 (7.57%) ^c
PB_775_0.50	60	43.7	719	931	0.314	0.312	0.176	0.285	0.027 (8.62%) ^c
PB_775_0.50	60	42.2	707	936	0.309	0.306	0.171	0.282	0.025 (8.02%) ^c
PB_775_0.50	60	42.4	707	949	0.308	0.306	0.174	0.282	0.023 (7.64%) ^c
PB_650_0.10 ^d	60	12.5	68.8	482	0.029	0.126	0.100	0.124	0.001 (1.07%) ^c
PB_650_1.00 ^d	180	25.0	447	750	0.191	0.214	0.154	0.200	0.014 (6.46%) ^c

359 ^c Calculated as $\frac{V_{meso}}{V_t} 100$

360 ^d Additional carbons that were not included in the design of experiments.

361 It was clear that a high pressure accelerated carbon gasification, even at 700 °C. At this
 362 temperature, an increase in pressure from 0.1 to 1.0 MPa led to a marked increase in the
 363 burnout, from 30.1% to 52.8%, due to the higher reaction rate. In this sense, an increased
 364 CO₂ partial pressure resulted in a higher fraction of reactant adsorbed on the surface of the
 365 sample [42]. These outcomes are in agreement with the study conducted by Malekshahian *et*
 366 *al.* [42], but disagree with the results reported by Feroso *et al.* [43], who observed that
 367 gasification of heartwood was enhanced using low partial pressures of CO₂.



368 **Fig. 2.** Normal plots of standardized effects ($\alpha = 0.05$) for physically activated biochars:
 369 specific surface area (a); total pore volume (b); mesopore volume (c); micropore volume (d);
 370 ultra-micropore volume (e).

371 The specific surface areas reported in Table 3 were in line with expectations, given the
 372 present activation conditions. Even at the lowest temperatures (700 °C) it was possible to
 373 obtain relatively high surface areas, even greater than those measured for activated biochars
 374 produced at higher temperatures. As can be seen in Fig. 2a, the specific surface area was
 375 positively influenced by the absolute pressure and negatively affected by the activation
 376 temperature. Especially at 1.0 MPa, relatively low temperatures and longer activation times
 377 allowed the biochar to be gasified more homogeneously, thus resulting in a more developed

378 porous structure. Conversely, higher temperatures led to faster reaction rates, which can
379 result in a more severe enlargement of micropores and the subsequent loss of surface area.

380 Fig. 2b shows that the total pore volume increased significantly when the pressure was
381 raised from 0.1 to 1.0 MPa. In general, the total pore volumes for physically activated
382 biochars, which were not significantly affected by the activation temperature, were higher
383 than those measured for chemically activated carbons. As can be seen when comparing the
384 data reported in Tables 2 and 3, physical activation led to porous carbons with slightly higher
385 ultra-micropore volumes. An inspection of Figs. 2d, 2e and A.3b reveals that high
386 temperatures can lead to some widening of the narrowest micropores, since the activation
387 temperature had a negative effect on the ultra-micropore volumes and a positive one on the
388 micropore volumes. With regards to the volume of mesopores, which was mainly affected by
389 pressure (see Fig. 2c), it should be noticed that physical activation under pressure resulted in
390 more hierarchical porous structures with higher contributions of mesopores (see values of
391 V_{meso} and relative percentages in Table 3). The statistically significant overall curvature terms
392 reported in Table A.3 for $S_{2D-NLDFT}$, V_t , V_{meso} , and V_{micro} response variables also suggest that a
393 central composite design could be required for optimization purposes.

394 Given the significant effects of pressure on the textural properties of carbons activated at
395 700–850 °C, we decided to perform two additional physical activations at 650 °C, leading to
396 the materials called PB_650_0.10 and PB_650_1.00. The results obtained, which are also
397 reported in Table 3, seem to confirm that activation with CO₂ under pressure is a very
398 interesting way to produce biomass-derived porous carbons with high specific surface area
399 and wide pore size distributions (including relatively high mesopore volumes), even at
400 relatively low temperatures.

401

402 3.3. *Catalytic activity*

403 Due to their relatively high specific surface area, both PB_700_1.00 and CB_3_0.10
404 activated biochars (one for each activation procedure) were selected as catalysts and tested
405 during the upgrading process of the aqueous phase of a real pyrolysis oil.

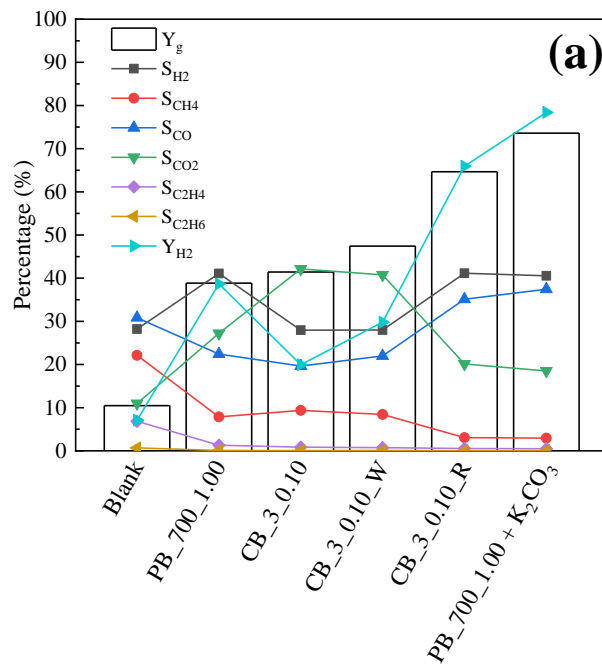
406 Elemental and moisture analyses of the aqueous phase of the pyrolysis oil revealed that the
407 liquid sample had an average chemical formula of $C_5H_{6.7}O_2$ and a water content of 80 wt. %.
408 Since the liquid was fed to the reformer without providing additional water, the steam to
409 carbon molar ratio, S:C, was 4:1.

410 It is generally assumed that deactivation and/or instability of the catalyst can be attributed
411 to two different phenomena: (1) the deposition of carbonaceous material (coke) on the
412 surface of the catalyst, which clogs part of the available active sites; and (2) the extent of
413 gasification reactions (both with steam and CO_2) of the carbon-based catalyst, which leads to
414 a loss of mass and a subsequent increase in LHSV (i.e., less contact time).

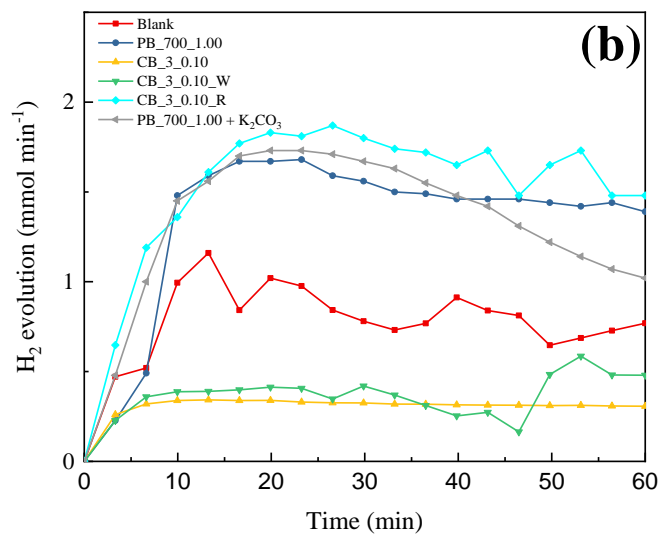
415 The results obtained in terms of total gas yield, selectivity towards specific gaseous
416 species, hydrogen yield and hydrogen release over time are displayed in Fig. 3a. It should be
417 noted that the poor results measured for the blank test (i.e., empty reactor) suggested that the
418 reactor wall did not play a crucial catalytic role. The physically activated biochar
419 (PB_700_1.00) exhibited performances comparable to those obtained in a previous study
420 [33]. Using this material, which showed a good stability during the whole experiment (see
421 Fig. 3b), a total gas yield of about 40% was obtained. Nevertheless, its relatively low
422 hydrogen yield indicates that the extent of steam reforming was modest.

423 Regarding the performance of the chemically activated biochar (CB_3_0.10), a similar
424 total gas yield was measured, compared to the physically activated one. However, selectivity
425 to gaseous products were markedly different, leading to a decrease in the production of H_2
426 (the hydrogen yield was even lower than that of PB_700_1.00) and an increase in that of

427 CO₂. This result could suggest that the decomposition of pyrolysis oil instead of steam and/or
 428 dry reforming, was the main process involved.



429



430

431 **Fig. 3.** Results obtained from combined steam and dry reforming of pyrolysis oil at 750°C:
 432 total and H₂ yields (Y_g and Y_{H₂}, respectively) as well as selectivities toward a given gaseous
 433 product, S_i (a); and evolution of the hydrogen production rate (b).

434 As can be deduced from Fig. 3a, similar results were also obtained for the chemically
 435 activated biochar that was washed with just water (CB_3_0.10_W). The differences in the

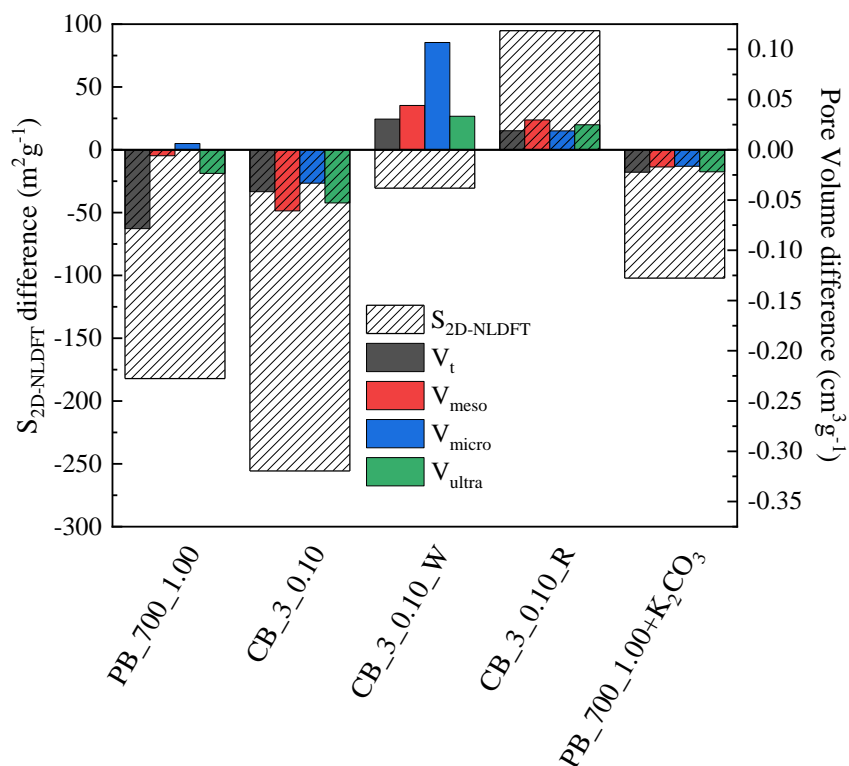
436 performance between the physically and chemically activated biochars could be related to
437 their textural properties. In this respect, the physically activated biochar exhibited a higher
438 specific surface area and higher volumes of micro- and mesopores, thus providing more
439 catalytic active sites of inherent alkali or alkaline earth metal species (AAEMs). The catalytic
440 role of these species (especially K, Mg and Ca) in both steam reforming and gasification of
441 carbon has been widely reported [25,44–46]. Furthermore, the more hierarchical pore size
442 distribution of PB_700_1.00, with relatively high contributions of the mesopores, could
443 shorten the diffusion path to reach the active sites.

444 Fig. 3b clearly shows that H₂ release rates fluctuated over time for most of biochars tested.
445 There are at least two possible explanations for this: First, the coke deposits, which gradually
446 covered the catalyst surface, could have led to large fluctuations in the extent of the reactions
447 involved. Second, the structural modification of the activated carbons (the process
448 temperature was higher than the activation temperature) could also lead to unsteady hydrogen
449 flow rates at the outlet. It should be noted that the reaction system studied was certainly
450 complex, due to the numerous competitive reactions that can affect the yields of the different
451 gaseous products. Nevertheless, it can be assumed that methane is mainly produced by the
452 cracking of the heaviest fraction of the pyrolysis oil, and that its yield will be correlated to the
453 amount of coke produced [47]. Similarly, Fig. 3a reveals that the selectivity towards methane
454 was almost constant for PB_700_1.00, CB_3_0.10, and CB_3_0.10_W materials. Therefore,
455 it can be concluded that the coke production in the process was quite similar for the above-
456 mentioned carbons, suggesting that some structural modifications in the carbons could be
457 accounted for the unsteady production of hydrogen. These modifications could be ascribed to
458 the reverse Boudouard reaction [48,49], steam gasification and, in the case of the chemical
459 activated biochars, potassium-catalyzed gasification. To support this argument, Fig. 4 shows
460 the textural properties of fresh and spent activated biochars. Contrary to the loss of porosity

461 observed for spent PB_700_1.00 and CB_3_0.10 catalysts, the spent CB_3_0.1_W material
462 showed an increase in pore volumes, especially for micropores. In addition, the results from
463 TPD measurements, which are summarized graphically in Figs. 5 and A.6, indicate that the
464 water-washed chemically activated carbon exhibited a larger mass loss (with an increased
465 release of CO₂ and CO, as shown in Fig. A.6) than that of the acid-washed carbon, thus
466 confirming a greater extent of carbon gasification. This fact could be attributed to the residual
467 amount of activation reagent available on the surface of the catalyst CB_3_0.10_W, which
468 might promote further biochar gasification during the experiment, as recently reported by
469 Wang *et al.* [11].

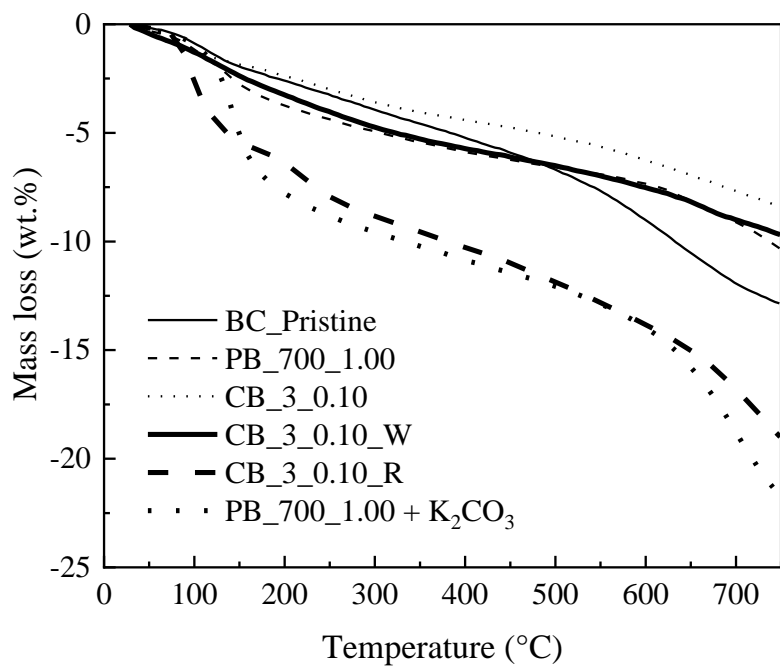
470 To understand better the effect of the washing procedure on the performance of
471 chemically activated biochars, an unwashed (raw) material (CB_3_0.10_R) was also tested as
472 catalyst. In this case, the total gas yield was 65% with enhanced selectivity towards both H₂
473 and CO (see Fig. 3a), thus indicating a higher extent of the steam reforming reactions
474 (leading to a Y_{H_2} of 66%). Fig. 6 displays the FTIR spectra obtained for fresh and spent
475 chemically activated biochars as well as for pure K₂CO₃. In the case of acid and water-
476 washed samples, no evident K₂CO₃ content was detected on the surface, indicating the
477 effectiveness of both washing procedures. On the other hand, most of the chemical activating
478 agent available on the surface of the fresh CB_3_0.10_R catalyst disappeared after the
479 catalytic test. The potassium carbonate (and other chemical species derived from its partial
480 decomposition) available on the surface of the catalyst at the beginning of the reforming test
481 (reductive environment) could progressively be reduced to metallic potassium (K⁰), leading
482 to a further promotion of reforming reactions. Furthermore, a process temperature relatively
483 close to the volatilization point of potassium can enhance the mobility and reactivity of the
484 metal [50]. Generally speaking, activated biochars have relatively abundant oxygen and
485 nitrogen-containing functional groups, which are not present in the materials shown in Fig. 6.

486 The reason behind this could be the relatively high activation temperatures, since the majority
 487 of functional groups decompose below 800 °C [51]. Furthermore, the relatively high amount
 488 of K_2CO_3 loaded on the analyzed samples resulted in a very large peak, which may hide other
 489 peaks of interest, such as those related to oxygenated functionalities.



490
 491 **Fig. 4.** Differences in textural properties between the fresh and spent activated biochars used
 492 as catalysts in steam and dry reforming tests.

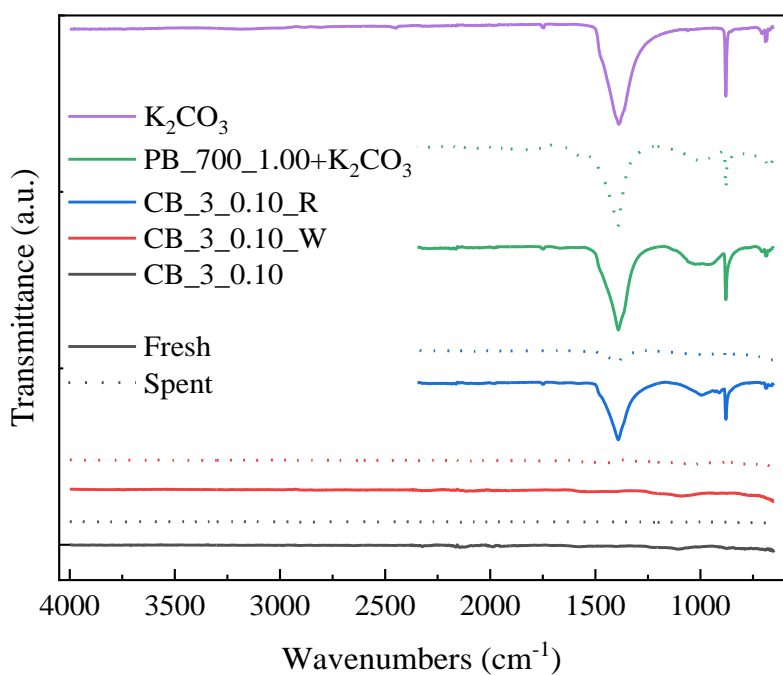
493



494

495 **Fig. 5.** Results from TPD characterization (percentage of mass loss vs. temperature) for

496 pristine and activated biochars.



497

498 **Fig. 6.** Comparison of the FTIR spectra obtained for pure K_2CO_3 and some activated biochars

499 tested in the present study.

500 The results from TPD measurements (see Fig. 5) showed that the mass loss under an inert
501 atmosphere was much higher for the unwashed catalyst. In addition, the release of CO during
502 TPD also increased for the material CB_3_0.10_R (see Fig. A.6c). The increase in both mass
503 loss and CO production when heating in N₂ up to 750 °C could be due to the reaction of
504 K₂CO₃ with carbon, which results in an additional release of CO, especially at temperatures
505 near 700 °C and above, and the formation of catalytically active COK and CK complexes
506 [52]. Hence, a certain gasification of carbon during the catalytic upgrading tests should be
507 expected. In line with this, Fig.4 shows that the spent CB_3_0.10_R exhibited more
508 developed textural properties than the fresh one (for instance, the specific surface area
509 increased from 29 m² g⁻¹ to 124 m² g⁻¹), thus confirming the extent of gasification reactions.
510 Despite the loss of catalyst through carbon gasification, which results in a higher LHSV, the
511 production rate of H₂ for this catalyst was relatively stable during the experiment. This could
512 be explained by the fact that the carbon gasification catalyzed by K₂CO₃ contributes to
513 creating new pores and then counterbalancing the deposition of coke, which induces
514 micropore blockage and active sites coverage.

515 Based on the findings discussed above, it seems reasonable to assume that an activated
516 biochar having a wider hierarchical pore size distribution and some availability of K₂CO₃ on
517 its surface appears as an excellent candidate for catalytic pyrolysis vapors upgrading. With
518 this in mind, a new biochar-derived material, designated as PB_700_1.00+K₂CO₃, was
519 prepared by impregnating the material PB_700_1.00 with K₂CO₃ at a mass ratio of carbonate
520 to precursor of 3:1. It should be expected that the better textural properties of the starting
521 physically activated biochar could lead to a more homogeneous dispersion of the active phase
522 on the carbon support, thus resulting in an improved performance of the catalyst.

523 The results obtained for the catalyst PB_700_1.00+K₂CO₃ confirmed our expectation. The
524 total gas yield reached a maximum value of 74%, whereas the hydrogen yield was the highest

525 reported in the present study (78%). From the FTIR spectra given in Fig. 6, it can be seen that
526 a fraction of K_2CO_3 was still present on the catalyst surface after running the test.

527 XPS analyses were carried out for both PB_700_1.00 and PB_700_1.00+ K_2CO_3 materials
528 to get additional insight into the role of K_2CO_3 during the pyrolysis vapors upgrading
529 process. From the summarized XPS results given in Table 4, it can be observed that the
530 atomic composition of PB_700_1.00 remained relatively constant after the catalytic test, even
531 though the simultaneous increase of the peaks related to CII and OII seemed to indicate the
532 formation of phenolic groups on the catalysts. The comparison of the atomic compositions
533 between the fresh and spent PB_700_1.00+ K_2CO_3 materials indicates that the extent of
534 carbon gasification reactions was much higher than that of the PB_700_1.00 material.
535 Furthermore, from the spectra displayed in Fig. A.7, an additional peak at 283 eV, which is
536 attributed to KC_x carbides, was observed for the spent PB_700_1.00+ K_2CO_3 . On the other
537 hand, the $K2p_{3/2}$ region presented only a contribution and suggested the presence of
538 potassium on the surface as K_2O or K_2O_2 , according to Li *et al.*[53]. This fact was also
539 confirmed by the shift to lower binding energies of the OII and OIII peaks (see Fig. A.7). The
540 marked difference in the potassium content between fresh and spent PB_700_1.00+ K_2CO_3
541 was probably due to the migration of the potassium-based compounds from the internal pores
542 to the external surface, which can lead to the formation of a coating layer composed of K_2O ,
543 as confirmed by SEM-EDX images reported in Fig. 7. The presence of potassium oxides on
544 the surface of the catalyst can be ascribed to (i) the direct decomposition of the carbonate
545 (Reaction 7) and/or (ii) the formation of metallic potassium. The latter can migrate from the
546 inner structure of the catalyst to its surface during the course of the pyrolysis vapors
547 upgrading tests and finally be oxidized when the catalyst has been removed from the reactor.
548 This fact can also justify the significant decrease in the specific surface area reported in Fig. 4
549 for the spent PB_700_1.00+ K_2CO_3 catalyst (from 130 to $27.7\text{ m}^2\text{g}^{-1}$).

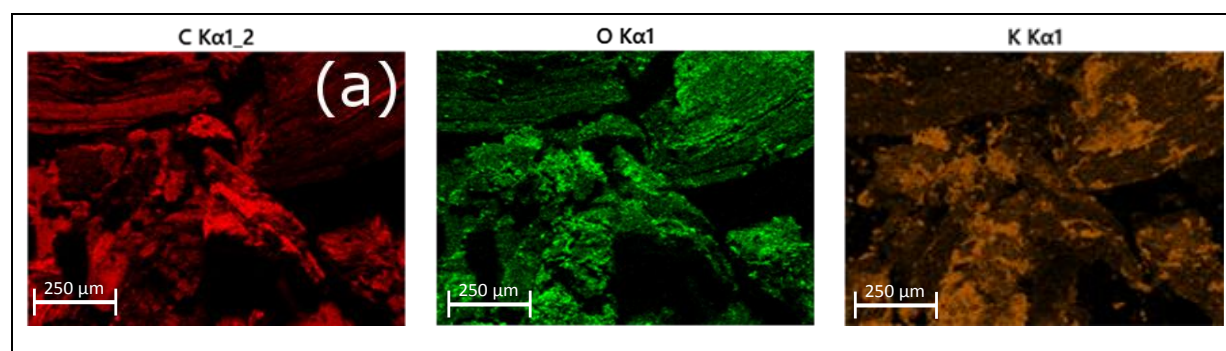
550 From the SEM-EDX analyses, it can also be deduced that a relatively homogeneous
 551 dispersion of K_2CO_3 was achieved for the fresh catalyst. Nevertheless, both fresh and spent
 552 catalysts exhibited some clusters and snail shell-like agglomerates, which could be attributed
 553 to the relatively high load of carbonate. However, no visible surface carbon layers was found
 554 on the superficial K_2O of the spent catalyst, thus suggesting that the extent of the carbon
 555 gasification was probably sufficient to prevent the deposition of coke on the surface.

556 **Table 4.** Surface composition measured by XPS and peak contributions of C1s, O1s and
 557 K2p.

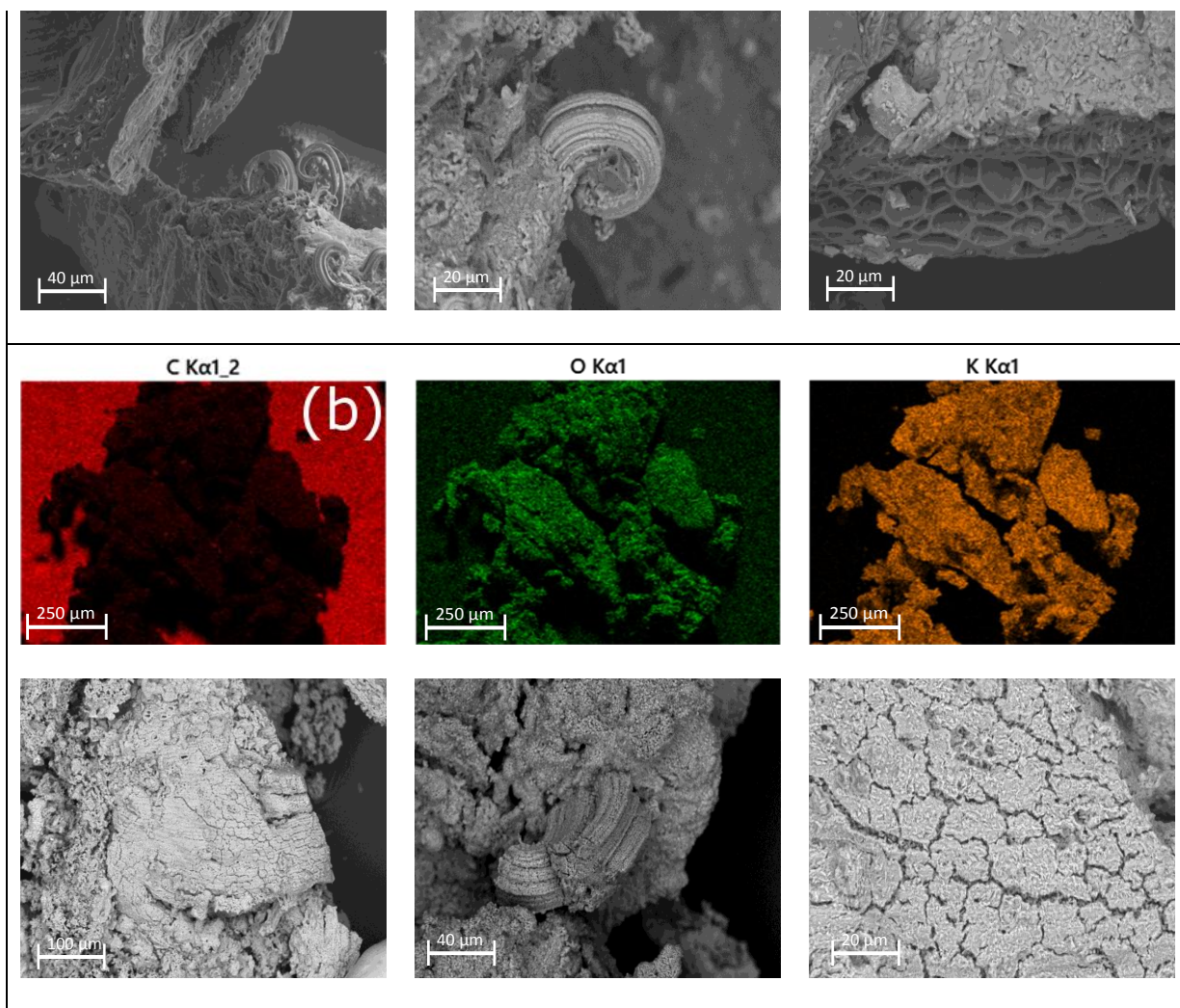
Sample	Surface concentration (at. %)			Binding energies (eV) and relative peak areas (%) ^e								
	C	O	K	CI	CII	CIII	CIV	Carbide	OI	OII	OIII	K2p3/2 region
PB_700_1.00 (Fresh)	87.3	11.3	1.4	284.4 (68.5)	285.7 (23.8)	287.6 (6.3)	289.2 (1.3)		531.1 (30.7)	532.8 (55.4)	534.5 (13.9)	293.3 (100)
PB_700_1.00 (Spent)	82.8	15.1	2.1	284.5 (62.7)	285.5 (31.3)	287.6 (4.1)	289.4 (1.9)		531.3 (20.0)	532.9 (60.2)	534.5 (19.8)	293.7 (100)
PB_700_1.00+ K_2CO_3 (Fresh)	72.6	21.6	5.8	284.6 (47.3)	285.8 (40.6)	287.6 (11.0)	289.3 (1.1)		531.3 (12.0)	532.7 (50.5)	534.5 (37.5)	293.7 (100)
PB_700_1.00+ K_2CO_3 (Spent)	42.0	27.8	30.2	284.5 (64.2)			288.7 (22.5)	283.1 (22.5)	531.1 (57.9)			292.8 (100)

558 ^e C1s binding energies in carbon materials: **CI**, hydrocarbons, aromatics, aliphatics (284.5 eV); **CII**, single bond
 559 C-O associated to alcohols, phenols, carboxyls (286 eV); **CIII**, double bond C=O in carbonyl, quinone (287.5
 560 eV); **CIV**, carboxyl and carbonate groups (288.7 eV).

561 O1s binding energies in carbon materials: **OI**, C=O quinone type groups (around 531 eV); **OII**, C-OH phenol
 562 groups and/or C-O-C ether groups (532.5 eV), **OIII**, chemisorbed oxygen (COOH carboxylic groups) and/or



563 water (535 eV).



564 **Fig. 7.** SEM-EDX images of fresh (a) and spent (b) PB_700_1.00+K₂CO₃ catalyst.

565 **4. Conclusions**

566 From the results obtained in this study, the following conclusions can be drawn:

- 567 1. The chemical activation of the pristine biochar led to carbonaceous solids with a
 568 relatively high specific surface area and well-developed microporous structure.
 569 The most important activation parameter was found to be the amount of activation
 570 agent, which accounted for the development of the porous structure. Activation
 571 pressure, on the other hand, was not significant for the outcomes of the procedure.
 572 Regarding the washing efficiency, acid washing resulted to be effective in
 573 cleaning the surface of the samples, whereas washing with water was not

574 sufficient to remove completely the residual reactants, partially clogging the
575 porosity of the samples.

576 2. Physical activation was found to be more sensitive to the process parameters. It
577 was possible to obtain acceptable values of specific surface area even at low
578 temperature (650 °C) by increasing the activation pressure to 1.00 MPa. A
579 negative effect of the activation temperature on the results was also observed. In
580 fact, the surface area decreased with the increase in the activation temperature,
581 probably due to the widening of the smaller pores by a too rapid (and thus less
582 controlled) gasification.

583 3. The most promising material produced by physical or chemical activation
584 (PB_700_1.00 and CB_3_0.10) were tested as catalysts in the upgrading of the
585 aqueous fraction of a real pyrolysis oil. Both activated biochars, which were
586 initially characterized by a well-developed porosity, did not exhibit good
587 performance, showing a huge decrease in the specific surface area, probably due
588 to coke deposition. Nevertheless, the addition of K_2CO_3 to the best physically
589 activated biochar boosted the overall performance towards the production of more
590 hydrogen, as a consequence of the enhanced reforming of pyrolysis oil. This
591 finding can be related to the availability of K^0 , which was formed from the
592 decomposition of the K_2CO_3 previously deposited on the activated biochar.

593

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604 **Appendix A. Supplementary data**

605 Further details on characterization of pristine and activated biochars. Statistical outcomes
606 from the adopted designs of experiments.

607 **Nomenclature**

608	A_{BET}	Brunauer-Emmett-Teller area ($\text{m}^2 \text{g}^{-1}$)
609	d_p	Pore diameter (nm)
610	F_{BO}	Molar flow rate of dry bio oil (mol min^{-1})
611	$F_{H2 Av}$	Experimental average H_2 molar flow rate (mol min^{-1})
612	$F_{H2 Stoi}$	H_2 stoichiometric molar flow rate (mol min^{-1})
613	m_0	Initial mass of biochar before activation (g)
614	m_c	Biochar mass before the washing step (g)
615	m_f	Final mass of biochar after activation (g)

616	m_g	Total mass of produced gas during the upgrading process (g)
617	m_l	Mass of liquid fed into the upgrading reactor (g)
618	m_w	Biochar mass after the washing step (g)
619	n_i	Produced amount of a given gaseous specie i (mol)
620	n_{tot}	Total amount of produced gas (mol)
621	$S_{2D-NLDFT}$	2D-NLDFT specific surface area ($m^2 g^{-1}$)
622	S_i	Selectivity toward a given gaseous specie i (%)
623	$V_{0.97}$	Gurvitch pore volume ($cm^3 g^{-1}$)
624	V_{meso}	Volume of mesopores ($cm^3 g^{-1}$)
625	V_{micro}	Volume of micropores ($cm^3 g^{-1}$)
626	V_t	Total pore volume ($cm^3 g^{-1}$)
627	V_{ultra}	Volume of ultra-micropores ($cm^3 g^{-1}$)
628	Y_g	Total gas yield (%)
629	Y_{H2}	Hydrogen yield (%)
630	Y_{wash}	Washing yield (%)
631	η	Degree of burnout (%)
632	<i>Acronyms</i>	
633	FTIR	Fourier-Transform Infrared spectroscopy
634	GHSV	Gas hourly space velocity
635	LHSV	Liquid hourly space velocity

636	PSD	Pore size distribution
637	RSM	Response surface methodology
638	S:C	Steam to carbon molar ratio
639	STP	Standard temperature and pressure
640	TPD	Temperature-programmed desorption
641	XPS	X-ray photoelectron spectroscopy
642	μ -GC	Micro gas chromatograph
643		

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