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1 Article

2 Laser Cutting Coupled with Electro-exfoliation to Prepare Ver- 3 satile Planar Graphene Electrodes for Energy Storage

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10 **Abstract:** The study of planar energy storage devices characterized by low-cost, high capacity, and
11 satisfactory flexibility, is becoming a valuable research hotspot. Graphene, monolayer sp² hybrid
12 carbon atoms with a large surface area, always act as its active component yet with a tension be-
13 tween high conductivity and easy to implement. Although the difficult-to-assemble graphene can
14 easily achieve planar assemblies in its highly oxidized form (GO), the undesired conductivity, even
15 after proper reduction, still restricts its further applications. Here, a facile “Top-down” method has
16 been proposed to prepare the graphene planar electrode via in-situ electro-exfoliation of graphite
17 supported on a piece of laser-cutting patterned scotch tape. Detailed characterizations have been
18 performed to study its physiochemical property evolution during electro-exfoliation. The obtained
19 flexible graphene planar electrodes show a decent energy storage performance, e.g., 40.8 mF cm⁻² at
20 a current density of 0.5 mA cm⁻² and a 81% capacity retention at a current density of 8 mA cm⁻² for
21 the optimized sample G-240. Their high conductivity also makes them possible to couple with
22 other redox-active materials by electrodeposition to improve their performance, e.g.,
23 ferrocene-functionalized mesoporous silica film (Fc-MS), MnO₂, and polyaniline (PANI). The
24 highest capacity is achieved with the PANI functionalized sample, which achieves a 22 times ca-
25 pacity increased. In a word, the versatility, practicality, and adaptability of the protocol to prepare
26 the planar graphene electrode proposed in this work make it a potential candidate to meet the
27 continuously growing energy storage demands.

28 **Keywords:** pseudocapacitive materials; graphene; planar electrode.

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30 1. Introduction

31 The last few decades have witnessed the continuously surging demands of flexible
32 and wearable microelectronics in various fields, such as biomedical sensors, wearable
33 products, and sports assist devices, etc [1,2]. Therefore, beyond higher energy and power
34 densities, these cutting-edge applications put forward more rigorous requirements for
35 energy storage systems, including ease of miniaturization, robust mechanical properties,
36 as well as high flexibility [3]. Generally, conventional energy storage devices, mostly
37 composed of vertical sandwich structures with two electrodes, one separator membrane,
38 and filled with liquid electrolyte, are too difficult to be fabricated into flexible mi-
39 cro-devices due to the complexity of such configuration [4]. In contrast, in-plane
40 interdigitated devices, comprising flatly aligned but spatially separated energy storage
41 materials with electrolyte coating inside directly, have gained more and more research
42 interest in recent years. This idiosyncratic design offers several merits over the traditional
43 sandwich structure in terms of the emerging application scenarios, such as:

- the distance between the adjacent working and coun-

ter electrodes could be minimized to some extent through advanced patterning techniques, which could reduce the internal resistance of the devices and improve their power output [5];

- the monolayer nature of the in-plane structure could largely increase the volume energy density by decreasing the thickness of devices, and also reduce short-circuit risks triggered by the external forces;
- the simplified architecture also makes it easier to be incorporated into electronic devices, and even suitable as the wearable power source when utilizing superior flexible supporting materials as substrates that can bear the long-term twisting and folding distortions.

Featured with a large surface area ($\sim 2630 \text{ m}^2 \text{ g}^{-1}$), excellent electrical conductivity ($\sim 104 \text{ S cm}^{-1}$), and robust mechanical properties (Young's modulus of $\sim 1 \text{ TP}$) [6], graphene is widely used as the active material for flexible in-plane energy storage devices when integrated on various flexible substrates, such as polyethylene terephthalate (PET), polydimethylsiloxane (PDMS), and textiles, for instance [7–9]. Instead of directly utilizing graphene, graphene oxide is mostly used as the starting material to assemble these devices owing to its abundant functional groups, hydrophilic property, and high surface adhesion energy [10]. However, after the reduction processes, the products, so-called reduced graphene oxide, cannot recover the pristine graphene structure either in thermal, chemical, or electrochemical ways; therefore they always suffer from relatively low conductivities [11]. Besides, one more technical difficulty to fabricate such devices lies in precisely patterning graphene on substrates to make full use of the surface space while avoiding the short-circuit problem. Several possible strategies have been developed in recent years, including chemically coating graphene on pre-patterned current collectors [12], laser writing on in-situ reducing graphene oxide [13], plasma etching [14], and inkjet printing [15]. Nevertheless, these methods are still facing some drawbacks: dependence on sophisticated facilities, complicated preparation processes, utilization of organic solvent, and limited graphene-loading amount. As such, developing a low-cost and facile method to assemble high-quality graphene on flexible substrates in a dense and patterned way is pivotal to promoting the development of this kind of device.

Electro-exfoliation is an efficient protocol to produce various types of two dimensional materials from their bulk counterparts in a scalable fashion [16]. This method can also be used to exfoliate bulk graphite into mono/few layers through a coupling of electric field driven sulfate ions intercalation and anodic water oxidation processes [17]. The flakes produced with such a method always show a high C/O ratio (~ 20) over other "Top-down" chemical methods (e.g., ~ 2 of Hummers method) [18], indicative of a small number of defects; as reported previously only the edge of the flakes can be oxidized [19]. Therefore, their properties, particularly electrical conductivity, should be more similar to graphene than graphene oxide [19], making them good candidates for energy storage electrodes. However, the as-formed graphene flakes are more likely to be peeled into the solution phase due to the intense bubbles produced by the constant water decomposition and the lack of binding sites at the graphite electrode. Although an "in-situ exfoliate and deposit" method has been developed to assemble vertically aligned graphene on conductive substrate [20], the graphene density of the final electrode is still relatively low, leading to a limited areal capacitance. If one tears a piece of graphite film with scotch tape to form a composite, the sticky tape surface would prevent the flakes from stripping during the electro-exfoliation process, and finally creates a scotch tape confined densely packed graphene electrode, and this is what we plan to investigate in the present work. Besides, the underlying scotch substrate also has several advantages, as discussed hereafter. The inherent characteristics of the scotch tape permit this electrode a high flexibility and robust mechanical nature, which makes them suitable for increasingly complex application scenarios. The pattern of the scotch tape can even be finely edited with a laser-cutting technique, which could be a low-cost, time saving, and versatile way to fabricate the in-plane planar devices. More than the substrate, the high conductivity and

99 large surface area of the graphene layer also make it easy to further composite with other
100 energy storage active materials. As a proof of concept, we will evaluate here some modi-
101 fiers that can be easily electrodeposited, such as, manganese dioxide (MnO_2), polyaniline
102 (PANI), and redox molecules (i.e., ferrocene moieties) functionalized mesoporous silica
103 films. Overall, a complete protocol to fabricate the planar graphene and/or its composite
104 energy storage electrodes is proposed here, and the obtained electrodes not only show
105 decent energy storage abilities but also permit adequate mechanical properties, which
106 may have the opportunity to be used in the next-generation flatly flexible energy storage
107 devices.

108 2. Results and discussions

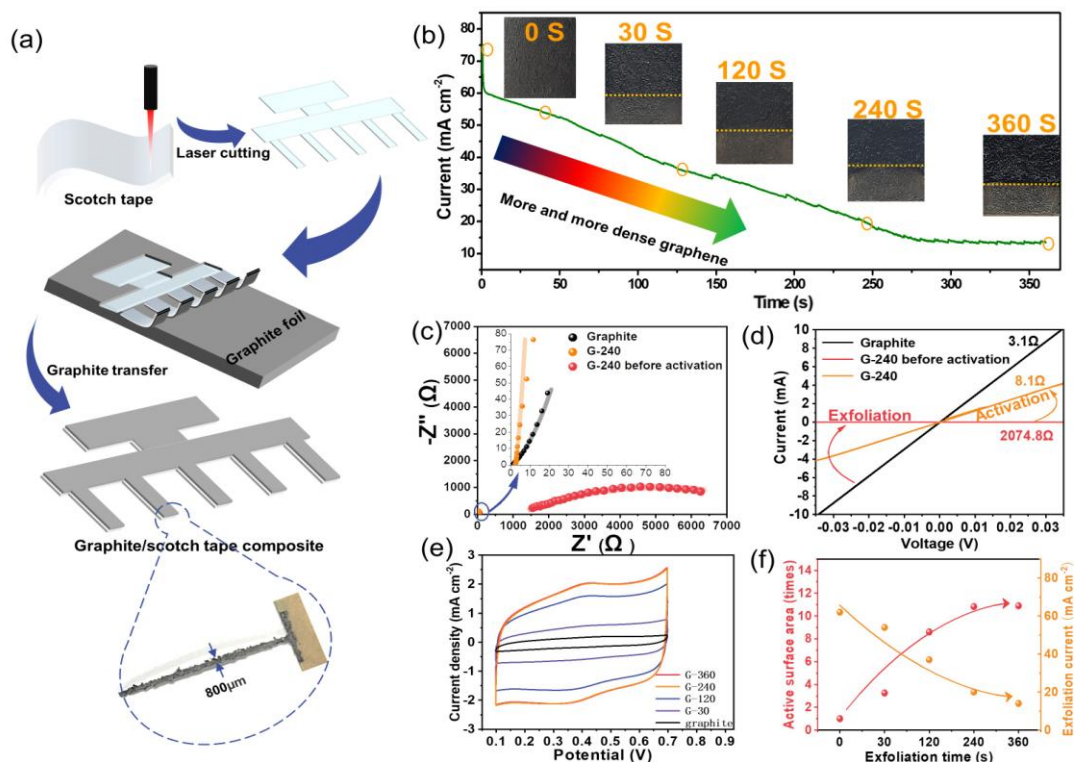
109 2.1. Preparation of the planar graphite electrodes

110 The electrode preparation process is illustrated in Figure 1a, and as one can see, the
111 laser first easily engraves the scotch tape to obtain different designed patterns, which
112 can be further used to peel off a piece of graphite film from a graphite foil via its sticky
113 surface. The resolution of this technique can reach here $800\ \mu\text{m}$, which suits the minia-
114 turization of the next-generation energy storage devices. Please also note that there is
115 still a room for optimization, and higher resolution might be achieved with the ad-
116 vanced laser generator.

117 This graphite-scotch tape composite undergoes a following electro-exfoliation pro-
118 cess to fabricate the final flatly flexible graphene electrodes. Figure 1b gives the current
119 profile of the electro-exfoliation process along with the corresponding digital photos of
120 the electrode at a certain time. One can see that after a fast drop from $75\ \text{mA cm}^{-2}$ to $60\ \text{mA cm}^{-2}$
121 within a few seconds, the exfoliation current gradually goes down from $60\ \text{mA cm}^{-2}$
122 to $10\ \text{mA cm}^{-2}$ until the exfoliation time reaches ca. 250 s, and then faces a steady re-
123 gion after that. The initial sharp fall of the current is probably due to the formation of an
124 electrical double-layer process, which could normally finish the charging within a few
125 seconds. Afterward, the anodic current mainly comes from the sulfate ions inserted into
126 the graphite interlayers and the oxygen evolution reaction [21], both of which help to
127 expand and separate the graphite layer into mono/few layers of graphene. The contin-
128 uously decreased current indicates the consumption of the graphite reactant and the in-
129 creasing resistance of the graphite/scotch electrode because the inserted sulfate ions in-
130 side the exfoliated graphene could deteriorate the conductivity, as confirmed by the later
131 activation experiments. The following current-steady region could be a signal that all
132 graphite on scotch tape has been converted into graphene. The residue current could be
133 a balance of the inserted sulfate ions between the electrical driving force and the ther-
134 modynamic distribution. Different from the almost constant current of directly elec-
135 tro-exfoliating graphite foil (Figure S1a), the gradually decreased current and the almost
136 clear final solution (Figure S1b) indicate the scotch tape confinement effect towards its
137 inside graphite, and the following characterizations will also prove that. The digital
138 photos of the electrode with distinct exfoliation time show the color of the exfoliated
139 material, the part above the yellow line gets darker with prolonging the exfoliation, in-
140 dicating a structural change.

141 Taking the 240 s exfoliation sample (G-240) as an example, the conductivity change
142 has been studied by both the AC impedance (Figure 1c) and the DC I-V curve (Figure
143 1d) to prove the need of the activation process at the potential of 0V (v.s. Ag/AgCl, 3 M).
144 From the Nyquist plot of the impedance spectra, two kinds of shape can be observed,
145 and obvious semicircle can be observed for the G-240 before activation sample, which
146 may come from the charge transfer between graphite and the intercalated sulfate ions.
147 Therefore, two equivalent circuits have been used to fit them. An extra R_{ct} element in
148 parallel to the constant phase element (CPE) has been added in the equivalent circuit for
149 graphite and G-240 sample. There corresponding circuits and detailed fitting results are
150 given in Table S1. As can be seen, the resistance of the graphite/scotch electrode is ~ 0.7

Ω , while its resistance surges to as high as 1,145 Ω after the electro-exfoliation process, which might be due to deterioration from the intercalated sulfate ions. Fortunately, the intercalated sulfate ions can be reversibly repelled out to solution by applying a 0V potential outside for another 240 s, and the current of this activation process is shown in Figure S2. A strong cathodic current density ranging from -35 mA cm^{-2} to -1 mA cm^{-2} can be first observed at the beginning of 40 s, which may come from the fast move-out of the negatively charged sulfate ions located at the surface/near-surface of graphene. For the period from the 40 s to 240 s, the current delivers a slowly decreased trend from -1 mA cm^{-2} to -0.25 mA cm^{-2} , and the much smaller but relatively steady current may reflect the leaching of sulfate ions from the lattice of few layers of graphene and un-exfoliated graphite. After the activation, most sulfate ions are removed from the graphene electrode, and the conductivity of G-240 recovers back to 2.21 Ω . The similar resistance change of this process can also be confirmed by the DC I-V test, where two platinum clamps with 0.5 cm in width well contact the targeted sample in a distance of 2 cm. The



resistances of graphite electrode, G-240, and G-240 before activation are respective 3.1 Ω , 8.1 Ω , and 20,740.8 Ω in our test condition.

Figure 1. The preparation processes and some features of the planar graphene electrodes. (a) The illustration of the preparation process with an editable pattern; (b) the current measured during the preparation process (current density variation with time of exfoliation), and corresponding digital photos of the electrodes at a certain exfoliation times; (c,d) the impedance spectra (c) and the I-V curves (d) of the graphite electrode and the graphene electrode with 240s exfoliation time respectively before and after activation; (e) the CV curves of the graphene electrodes prepared with different exfoliation times at 50 mV s^{-1} in 1M H_2SO_4 electrolyte (*vs.* Ag/AgCl 3 M); (f) the electroactive surface area and the exfoliation current for different exfoliation times obtained from (b) and (e), respectively.

Besides, the relation between the exfoliation time and the electroactive area has also been analyzed by the cyclic voltammetry (CV) method. (Figure 1e) With prolonging the exfoliation time, the corresponding sample shows the current response steadily increases to 11 times until 240 s, and then reaches a stable state afterward. (Please note the active area here is not the actual area of the electrode but only reflects the increased times of the area) While the exfoliation current gives an opposite trend, it decreases first and then almost stays stable. This phenomenon not only proves the feasibility of using elec-

tro-exfoliation to prepare large surface area graphene electrodes but also demonstrates the exfoliation current could be used as an indicator to evaluate the density of graphene. The G-240 is therefore the optimized sample, and its related physiochemical properties as well as energy storage abilities are further evaluated in following.

2.2. Physico-chemical characterization of the planar graphite electrodes

The proposed graphene formation process during the electro-exfoliation is illustrated in Figure 2a. The intercalation of sulfate ions inside the interlayer of graphite would weaken the interaction between the adjacent graphene layers and the simultaneous oxygen gas evolution will tear them into graphene. The XRD patterns variation with different exfoliation time (Figure S3) prove this assumption to some extent, as the decreased intensity and the widening half-peak width of the (002) peak with extending the exfoliation time indicates the insertion of sulfate ions will affect the integrity of the graphite lattice. Besides, from our observation, long-time exfoliation will also lead to a stronger peak at ca.17 θ , which might be due to the formation of the sulfate intercalated graphite [22].

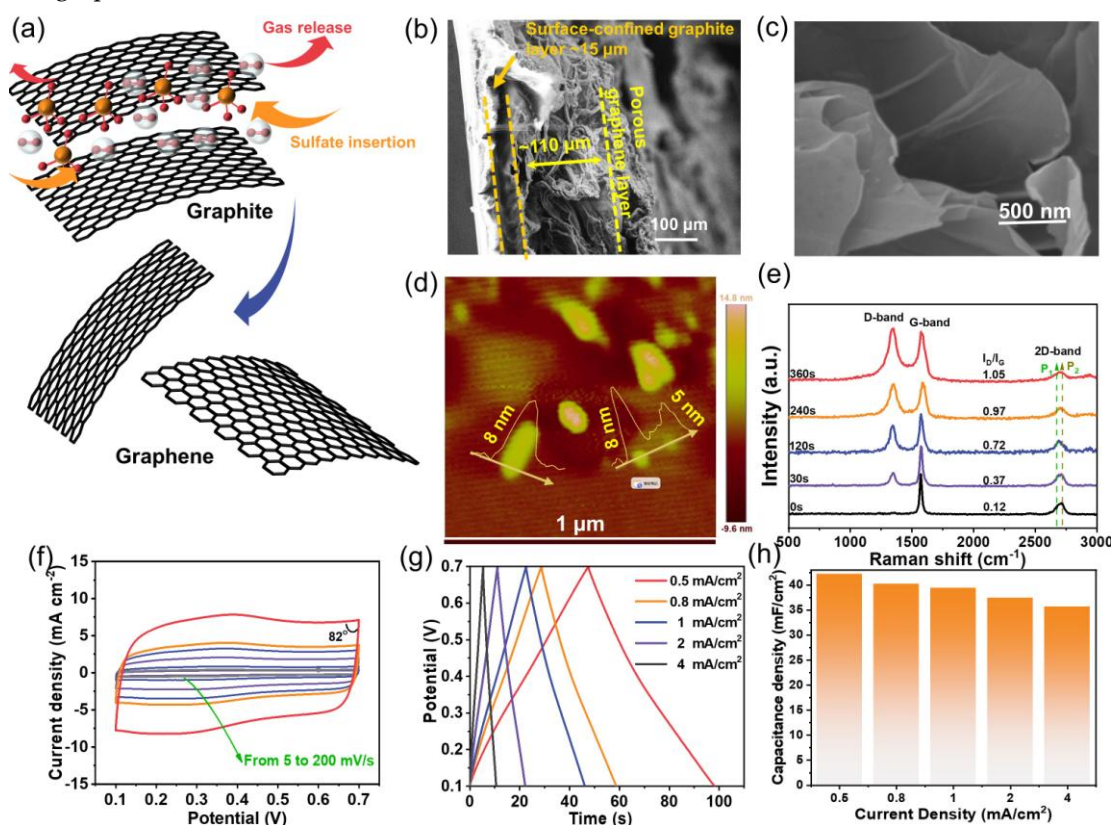


Figure 2. The detailed physiochemical properties of the prepared electrodes (especially G-240) and the energy storage abilities of G-240. (a) the illustration of the conversion from graphite to graphene during the electro-exfoliation process; (b,c) the respective cross-section (b) and high-resolution (c) SEM images of the G-240; (d) the AFM image of the flakes from G-240; (e) the Raman spectra of the sets of graphene electrodes; (f) the CV curves of G-240 from the scan rates of 5 to 200 mV s^{-1} in the solution of 1M H_2SO_4 (vs. Ag/AgCl 3 M); (g,h) the charge-discharge curves and corresponding capacitance of G-240 at the current densities from 0.5 to 8 mA cm^{-2} , respectively.

The microstructure of the graphite/scotch electrode and G-240 electrode have been characterized by scanning electron microscopy (SEM). The cross-section view of the original graphite/scotch tape electrode (Figure S4) shows a compact graphite layer with a thickness of $\sim 50 \mu\text{m}$, and the underlying brightest region that arose from the charge accumulation effect is the insulating scotch tape substrate. After the exfoliation, the cross-section of G-240 (Figure 2b) shows that the material becomes heterogeneous. It is

211 composed of two distinct layers: the highly porous top layer with the thickness swelling
212 to ca.110 μm due to the oxygen bubbles impact, and the layer near the scotch tape with
213 the thickness ca.15 μm , which seems not being affected by the exfoliation process and
214 still manifests the pristine graphite structure due to the scotch-tape confinement effect.
215 Further, the high-resolution SEM image (Figure 2c) of the top layer also reveals it is
216 composed of few layers graphene, which is also confirmed by the AFM result (Figure
217 2d) with a thickness of 5~8 nm and the almost transparent flakes from the TEM image
218 (Figure S5). The anisotropic structure feature of G-240 could benefit its energy storage
219 application, notably because the highly porous graphene top layer will not only possess
220 continuous channels for counter ions diffusion and but also provide large surface area
221 for the energy storage or even compositing with other active materials. In addition, Ra-
222 man spectra (Figure 2e) also demonstrate the electro-exfoliation will also alter the carbon
223 hybridization status, as the ratio of D-band (sp^3 hybridization)/G-band (sp^2 hybridiza-
224 tion) increases with extending the exfoliation time. The gradually increased D-band
225 might arise from the surface oxidation of the carbon basal plane, and/or the edge sites of
226 the newly formed graphene [23]. The 2D-band, sensitive to the layer of graphene, has
227 also been observed for different samples. For monolayer graphene, its intensity should
228 be much higher than that of G band. However, this is not the case for our samples, and
229 similarly low intensities 2D bands have been observed for all samples. This can be as-
230 cribed to the composite nature of our samples, as the graphite layer next to the scotch
231 tape cannot be exfoliated. Despite this, their noticeable shape change can be observed for
232 different exfoliation time. The shoulder peak of graphite at the lower Raman shift (P_1)
233 gradually increases with extending exfoliation time, indicative of the formation of
234 few-layers graphene [24]. While the intensity of P_2 peak gradually decrease, suggesting
235 the reduced ratio of graphite in the composites.

236 As mentioned above, the G-240 is the most “cost-effective” sample and therefore its
237 energy storage properties have been evaluated in detail. The CV curves of G-240 are
238 recorded at various scan rates from 5 mV s^{-1} to 200 mV s^{-1} in 1M H_2SO_4 aqueous solution.
239 The corresponding results (Figure 2f) show a rectangular shape, despite a pair of very
240 small redox peaks at ca. 0.35 V coming from oxygen-containing functional groups [25]
241 (to clarify this pair of peaks, the CV curve at a low scan rate of 1 mV/s is given in Figure
242 S6), no matter at the relatively low scan rate of 5 mV s^{-1} or at the high scan rate of 200
243 mV s^{-1} . This behavior is a double-layer type process, where the charges are accumulated
244 on the surface of conductive surface, and the counter ions are physically adsorbed on-
245 side the surface of the electrode by the potential driven force at the meantime. There is
246 no charge could transfer across the interface during the formation of the double-layer,
247 and therefore its behavior can be described as a capacitor in the equivalent circuit. Be-
248 sides, the well-maintained rectangular shape and the nearly 90° turning at the switching
249 potential (even 82° at the scan rate of 200 mV s^{-1}) further demonstrate the negligible in-
250 ternal resistance and fast ions diffusion process during charge-discharge processes. The
251 energy storage ability of G-240 is assessed with the galvanostatic charge-discharge tech-
252 nique (Figure 2g) with current densities from 0.5 mA cm^{-2} to 8 mA cm^{-2} . Despite a slight
253 distortion at ca. 0.35 V, the symmetrically triangular charge-discharge curves show the
254 pure electrical double-layer energy storage process and high coulombic efficiency. Based
255 on these curves, the G-240 shows a high-rate performance with 40.8 mF cm^{-2} at a current
256 density of 0.5 mA cm^{-2} (49 s charge/discharge time) and 81% (33.0 mF cm^{-2}) retention at a
257 high current density of 8 mA cm^{-2} (2 s charge/discharge time). Unsurprisingly, the G-240
258 delivers a robust cycling performance (Figure S7). Besides, there is no obvious difference
259 between the impedance of the original G-240 and that after its long-term cycling, which
260 indicates its stable structure (Figure S8).

261 2.3. Modification of the planar graphite electrodes to get pseudocapacitive materials

262 Rather than using them alone, as aforementioned, the large surface area and good
263 electrical conductivity of graphene electrodes make it possible to further turn them into

nanocomposites with associating other redox-active materials in order to improve their energy storage performance. Facile electrodeposition methods have been adopted for the modification of G-240, respectively with ferrocene functionalized mesoporous silica film (Fc-MS), manganese dioxide (MnO_2), and polyaniline (PANI), as illustrated in Figure 3a.

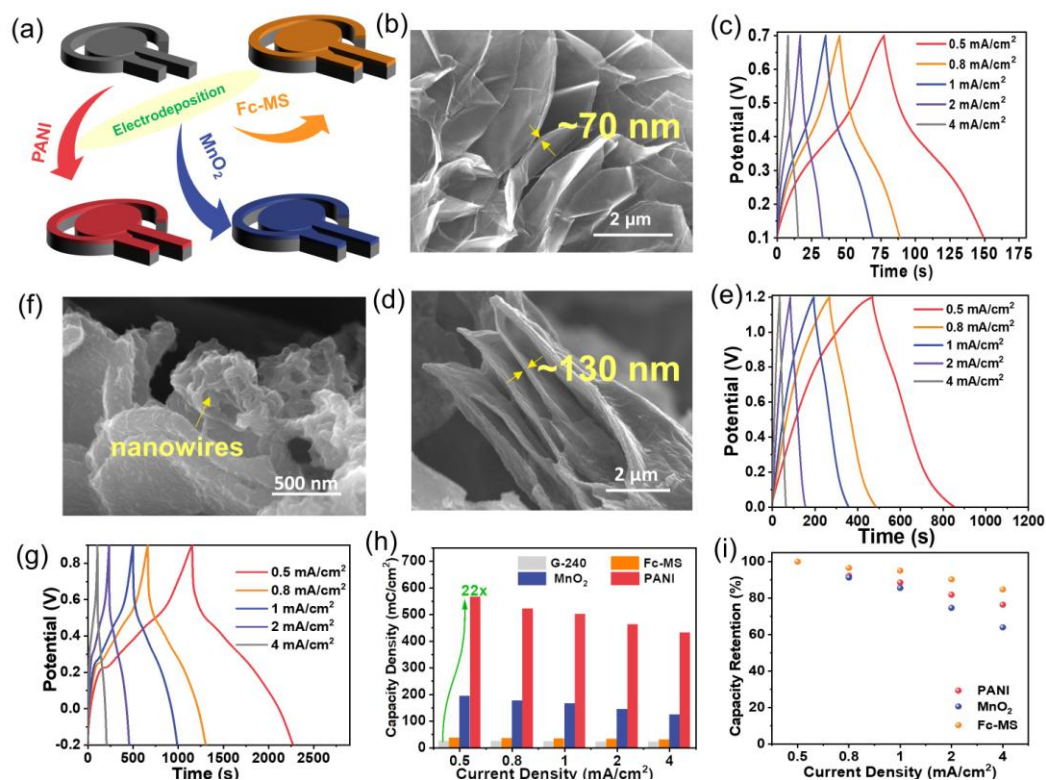


Figure 3. The preparation of the binary composites between G-240 and Fc-MS, MnO_2 or PANI by electrodeposition methods. (a) The illustration of the composition between graphene and corresponding redox active materials; (b–g) the typical SEM images (b,d,f) and energy storage abilities (c,e,g) of G-240 functionalized with Fc-MS, MnO_2 or PANI, respectively; (h,i) the calculated capacity density (h) and rate performance (i) of corresponding materials.

The corresponding SEM images (Figure 3b,3d, & 3f) and elements mapping results (Figure S9) reveal that all three redox-active components have been well integrated on the surface of graphene, yet in different forms. The Fc-MS and MnO_2 cover the graphene nanosheets in the film with a thickness of ~ 70 nm (similar to our previous report [26]) in the former case, and 130 nm in the latter case. In contrast, the electropolymerized PANI was found to independently grow on the graphene surface in a nanowires form. The loading mass of PANI and MnO_2 on the G-240 are 1.9 mg/cm^2 , and 3.1 mg/cm^2 . Regarding the Fc-MS functionalized sample, the loading of the ferrocene molecules is in a trace amount, which is below the detection limit of the balance. But our previous work has proved that almost all ferrocene molecules are redox active [26]. Please note that not limited to the three components given here, these results demonstrate the prepared graphene electrode can act as a versatile substrate to manufacture composites with extensive nanomaterials. The energy storage performance of these three nanocomposites have been further evaluated with chronopotentiometry, and the corresponding curves are given in Figure 3c, 3e, & 3g. Different from that of graphene electrodes, the shapes of all curves are distorted, no longer triangular shape, due to the redox reactions at a certain potential, which will finally lead to a potential-dependent capacitance. Although their shapes are still not the same as that of battery-type electrodes, coulomb should be a better unit to evaluate their performance, instead of farad [25]. One can see, at any current densities, much longer charge/discharge time can be obtained for the prepared composites than that of the G-240, which means the much more charges can be stored

295 inside (Figure 3h&3i). The maximal capacity increases to ca. 22 times from 24.5 mC cm⁻²
296 for G-240 to 565.5 mC cm⁻² for PANI-G at the current density of 0.5 mA cm⁻². For sure,
297 different composites show different energy storage behaviors, and better performance
298 could be achieved with further optimizations. All samples show more than 85% capacity
299 retention after 1,000 consecutive cycles at a scan rate of 100 mV s⁻¹ (Figure S10). The im-
300 pedance data of MnO₂ and PANI samples before and after long-term cycling have been
301 enriched in Figure S11. Their spectra are composed a semicircle and long tail, indicative
302 of a charge transfer and a capacitor-like behavior. Besides, there is no obvious change
303 between pristine sample and that after cycling, which indicates their stable structure.
304 Concerning the impedance of Fc-MS functionalized sample, one can refer to our previ-
305 ous work [26,27].

306 The energy storage of the three composites involves charge transfer processes
307 across the electrode interfaces, which is much more complicated than the pristine G-240
308 material. Therefore, the identification of the rate-determining step should be meaningful
309 to improve their energy storage characteristics. The CV curves of the composites can
310 maintain their originally intrinsic shapes only in a relatively slow scan rate range from
311 mV s⁻¹ to 20 mV s⁻¹ (Figure 4a, 4d, and 4g). As the nature of the R-C circuit, serious po-
312 larization and the resistance-controlled behavior will occur in “high scan rates”, which is
313 due to the increased surface-capacitance as well as the resistance from the solution, elec-
314 trode materials, and charge transfer processes. The definition of “high scan rates” is
315 however different for different samples (Figure S12). As shown, obvious distortion occur
316 at a scan rate of 50 mV s⁻¹ for the PANI sample, but Fc-MS sample can still maintain its
317 original shape at the scan rate of 200 mV s⁻¹, which means the Fc-MS could be used for
318 power-type energy storage devices, as discussed in our previous work in detail, yet for
319 non-flexible electrodes [26,27]. The rate-determining step of energy storage process in
320 the low scan rates domain has also been analyzed by fitting the CV curves with the fol-
321 lowing equation (1):

$$i(V) = k_1v + k_2v^{0.5} \quad (1)$$

322 where k_1v and $k_2v^{0.5}$ represent the surface-controlled current and diffusion-controlled
323 current respectively. The fitting results (Figure 4b, 4e, and 4h) reveal that 57%, 77%, and
324 90% of total current come from the surface-controlled current at the scan rate of 10 mV
325 s⁻¹ for PANI, MnO₂, and Fc-MS respectively. This means that the ions diffusion resis-
326 tances are also different during the energy storage process, and that the non-existing
327 solid-state ions diffusion for the Fc-MS sample gives a much higher surface-controlled
328 current ratio, even similar to that of G-240 electrode at 95 % (Figure S13). Similar trends
329 can also be seen from fitting the redox peak currents vs. potential scan rates with the
330 equation (2):

$$i = av^b \quad (2)$$

331 where the empirical b value is an indicator of the type of reaction, ranging from 0.5 to 1
332 (0.5 for a diffusion-controlled behavior, and 1 for a surface-controlled charge transfer).
333 The largest b value of Fc-MS (~0.91) demonstrates again a major surface-controlled redox
334 process, in contrast to the diffusion-controlled behavior of PANI sample (b value of
335 ~0.58).

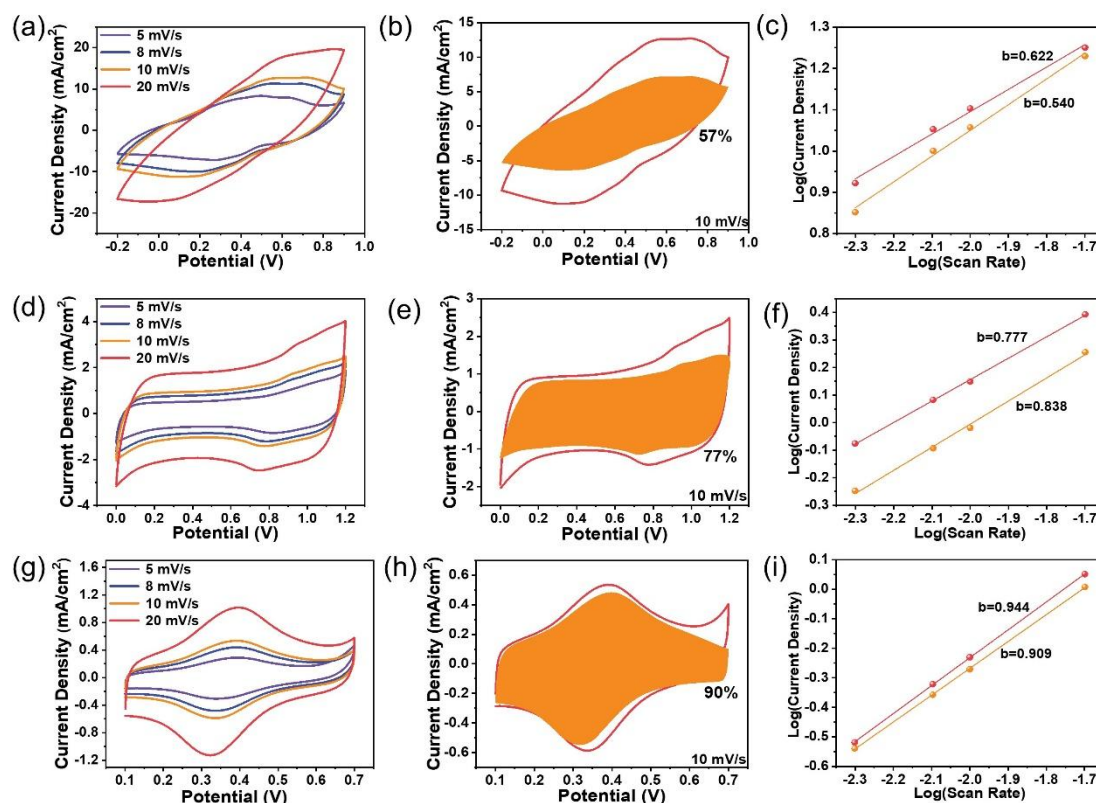


Figure 4. (a–i) The kinetic analysis of the energy storage behaviors of Fc-MS, MnO₂, and PANI: (a,d,g) the cyclic voltammetry of the three samples from 5 to 20 mV s⁻¹; (b,e,h) the calculated surface-controlled current from the respective CV curves (*vs.* Ag/AgCl 3 M); (c,f,i) the respective fitting slopes of the peak currents at different scan rates on a logarithmic scale.

2.4. Evaluation of a symmetrical device

Two pieces of PANI functionalized graphene electrodes have also been assembled into a symmetrical device to evaluate the possible practical application of the flexible graphene-based nanocomposite electrodes (Figure 5a–5d). The same acidic electrolyte as that above in a three-electrode configuration has been adopted to run the test for the assembled device. Figure 5a shows a group of CV curves from a scan rate of 5 mV s⁻¹ to 200 mV s⁻¹ in a potential window from -0.2 V to 0.9 V. A pair of broad redox peaks appear at ca. 0.3 V at the scan rate of 5 mV s⁻¹, which is consistent with that observed with the three-electrode system (Figure 4g), although being characterized by somewhat larger anodic-to-cathodic peak-to-peak separation. Indeed, the peak-polarization caused by the resistance is observed, especially with fastening the scan rate, but the rather symmetrical curves are still noticeable, even at the high scan rate of 200 mV s⁻¹, indicative of an expected rate performance. The areal capacitance of the device is examined by chronopotentiometry with applying various current densities ranging from 0.5 mA cm⁻² to 4 mA cm⁻² in a same potential window as CV tests, as plotted in Figure 5b. There is no obvious plateau, despite a little distortion, even at a small current density of mA cm⁻², which demonstrates the pseudocapacitive nature of the device. Its derived capacitances are given in Figure 5c, and one can see a high areal capacitance of 79.8 mF cm⁻² at a current density of 0.5 mA cm⁻², and a 51.3% capacity retention at a current density of 4 mA cm⁻². Its Ragone plot is given in Figure 5d for its easy comparison with other similar film-type devices [28–32]. The present one can deliver a robust cycling performance, reaching a capacity retention of 78% after 1,000 consecutive cycles at a scan rate of 100 mV s⁻¹ (Figure S14). Following this conceptual demonstration, future works could be di-

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rected to replacing this ‘classically’ deposited PANI by PANI nanowire arrays exhibiting ultra-fast electrochemical responses [33].

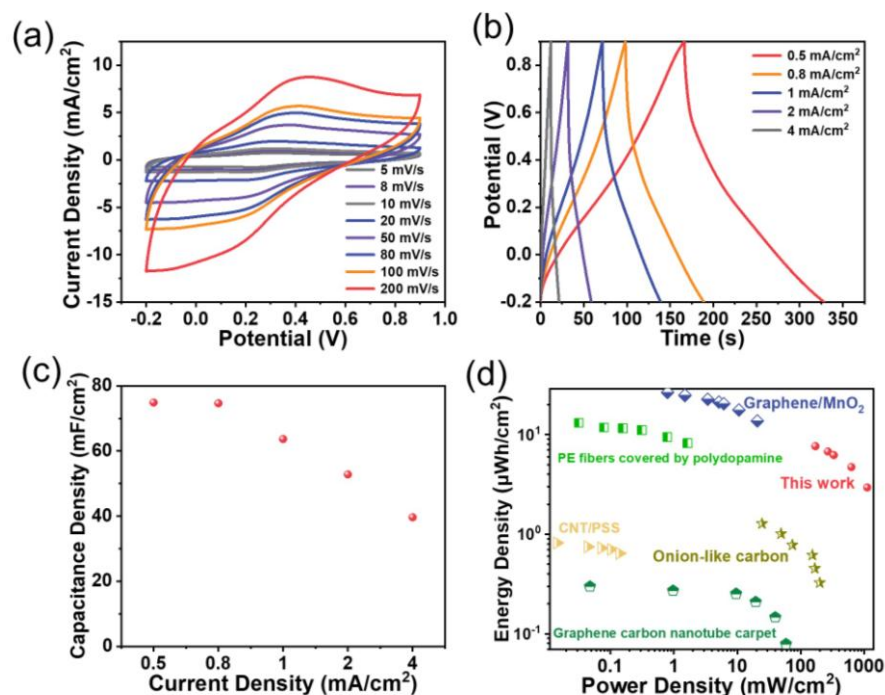


Figure 5. The energy storage performance of the PANI planar device in a symmetrical configuration: (a) the CV curves recorded at scan rates ranging from 5 to 200 mV s⁻¹ in the potential window extending from -0.2 V to +0.9V; (b) the charge-discharge curves at current densities from 0.5 to 4 mA cm⁻²; (c) the corresponding capacitance calculated from the charge-discharge curves; (d) the Ragone plot of this device and some reported benchmark devices: Graphene carbon nanotube carpet [28]; CNT/PSS [29]; Onion-like carbon [30]; PE fibers covered by polydopamine [31]; Graphene/MnO₂ [32].

3. Materials and Methods

Reagents. Tetraethoxysilane (TEOS, 98%), 3-chloropropyltrimethoxysilane (Cl-PTES, 97%), and graphite were purchased from Alfa Aesar, and cetyltrimethylammonium bromide (CTAB, 99%) was purchased from Acros. Acetonitrile (ACN, 99%), dimethylformamide (DMF, 99%), and cyclohexane (99%) were obtained from Merck. Ethynylferrocene (97%), lithium perchlorate (95%), lithium chloride (95%), copper sulfate (99%), sodium diethyldithiocarbamate trihydrate, and HCl (37%) were purchased from Sigma-Aldrich. Sodium nitrate (99%), ascorbic acid (97%), tetrabutylammonium bromide (99%), and poly(vinyl alcohol) (PVA, MW = 72,000 g/mol) originated from Fluka. Graphite foils and scotch tape were purchased from Shanghai Carbon Co. Ltd, and 3M, respectively. Sodium sulfate anhydrous (99%), manganese sulfate (95%), concentrated sulfuric acid (98%), and aniline (99%) are from Meryer, Kaitong, Kernal, and Macklin, respectively. All compounds were used directly without further purification. The (3-azidopropyl) triethoxysilane (AzPTES) was synthesized with the protocol reported previously [34].

Apparatus. Transmission electron microscopy (TEM) investigation was performed with a JEOL ARM 200F – Cold FEG TEM/STEM equipped with a GIF Quantum ER. Scanning electron microscopy (SEM) analyses were carried out with a JEOL JSM-840 or Zeiss SUPRA 55 apparatus equipped with an energy-dispersive X-ray (EDX) microanalyzer and the high-resolution SEM micrographs have been obtained with the model JSM-IT500HR apparatus (JEOL). Atomic force microscopy (AFM) analysis was done with Asylum JEOL JSPM-5200. Raman spectroscopy measurements were carried out with a RenishawinVia spectrometer with a green light laser (532 nm). X-ray diffrac-

tion (XRD) pattern was obtained using the Nanoviewer from Rigaku (CuK α radiation) at a scanning rate of 5°/min.

Electrochemistry. Electrochemical experiments (including exfoliation of the graphite foil, electrodeposition of the silica film, PANI, MnO₂, cyclic voltammetry, and charge-discharge measurements) were performed with Autolab PGSTAT-300, μ Autolab, or Biologic SP-150 potentiostat. The Ag/AgCl (3.5 M KCl) electrode and platinum plate (1 cm \times 1 cm) served as the reference electrode and counter electrode, respectively. An aqueous solution containing 1 M H₂SO₄ was used as the electrolyte for the performance test of the graphene samples, and PANI/Fc-MS functionalized samples, while 1 M Na₂SO₄ electrolyte was used for investigating the MnO₂ functionalized sample.

Preparation of the scotch tape-supported graphene samples (G-X, where X represents the exfoliation time). The laser cutting technique was first used to pattern the scotch tape to a certain design. Subsequently, the as-patterned scotch tape was pasted on the surface graphite foil followed by a peeling-off process to hand over a piece of graphite on the scotch tape. After that, similar to the electro-exfoliation method reported elsewhere [17], the electro-exfoliation process was further performed in 0.1 M Na₂SO₄ aqueous solution with a two-electrode system, where a piece of graphite coated scotch tape (graphite/scotch electrode) and platinum mesh (2.5 cm \times 2.5 cm) were used as the working and counter electrode, respectively. The exfoliation was achieved by applying a voltage of 4 V for 240 s with a distance of 5 cm between the two electrodes. The obtained intermediate is highly resistive. To recover its conductivity, an activation process has been used via applying 0 V on the intermediate for 240 s in a three-electrode system with Ag/AgCl as the reference to remove the inserted sulfate ions. Finally, the as-prepared electrode, labeled as G-240, was thoroughly washed with deionized water and dried at room temperature overnight. Besides, different exfoliation times of 30 s, 120 s, and 360 s were also conducted to investigate the structural evolution, and the corresponding products were labeled as G-30, G-60, G-120, and G-360, respectively.

Preparation of the graphene-supported ferrocene functionalized silica thin film electrode (Fc-MS-G). The preparation protocol is similar to what we reported before [26], except a piece of G-240 was used as the working electrode in the present study. The finally obtained composite is labelled as Fc-MS-G for simplicity.

Preparation of the graphene-supported MnO₂ electrode (MnO₂-G). Cyclic voltammetry was used to deposit MnO₂ on the graphene surface with a three-electrode configuration, including a piece of G-240 as the working electrode, a platinum plate counter electrode, and Ag/AgCl (3 M) as reference electrode. The deposition was carried out within a potential window ranging from 0 V to +1.2 V at a scan rate of 10 mV s⁻¹ in 0.1 M MnSO₄. After 40 consecutive cycles and thorough cleaning with water, one can obtain the MnO₂/graphene composite, labelled as MnO₂-G for simplicity.

Preparation of the graphene-supported PANI electrode (PANI-G). The same configuration of the MnO₂ preparation, but with a different electrolyte solution consisting of a mixture of 0.1 M H₂SO₄ and 0.01 M aniline, was used to prepare the graphene-supported PANI sample. The deposition process was achieved via 40 consecutive cycles in a potential window extending from -0.2 V to +1.0 V at a scan rate of 10 mV s⁻¹. After thoroughly cleaning with water, one can obtain the PANI/graphene composite, labelled as PANI-G for simplicity.

4. Conclusions

In this work, a laser-cutting coupled with electro-exfoliation protocol has been proposed to prepare planar graphene-based energy storage electrodes with a resolution of 800 μ m. Compared to the graphene electrodes reported elsewhere, the electrodes prepared with this “Top-down” method well solve the tension between high conductivity and easy processability. The optimized G-240 graphene sample shows the rate performance of 40.8 mF cm⁻² at a current density of 0.5 mA cm⁻² and 81% (33.0 mF cm⁻²) retention at a high current density of 8 mA cm⁻². The capacity can be further improved via

the formation of composites with other redox-active components, e.g., Fc-MS, MnO₂, and PANI in this work. The maximum capacity can increase 22 times to 565.5 mC cm⁻² for the PANI functionalized sample, for instance. Thanks to the flexibility of the scotch tape substrate, a flexible symmetric pseudo-capacitor with a pair of PANI functionalized graphene electrodes has been assembled, showing a capacitance of 79.8 mF cm⁻². Such satisfactory results prove the protocol of this work in offering a new idea for the construction of flexible planar energy storage devices.

Supplementary Materials: The following supporting information can be downloaded at: www.mdpi.com/xxx/s1, Figure S1: (a) a typical exfoliation current for a piece of graphite foil; (b) Respective photographs of the final solution that exfoliates graphite-scotch tape and a piece of graphite under a same condition; Figure S2: the activation current profile of the G-240 at the potential of 0V for 240s; Figure S3: the XRD pattern of the set of graphite samples exfoliated at various times; Figure S4: the cross-section SEM view of the original graphite/scotch tape electrode; Figure S5: the TEM image of the G-240; Figure S6: the CV curve of the G-240 at a scan rate of 1 mV s⁻¹ in 1M H₂SO₄ with a reference of Ag/AgCl (3 M); Figure S7: the cycling performance of G-240 at a scan rate of 100 mV s⁻¹; Figure S8: the impedance spectra of G-240 and G-240 after long-term cycling at OCP in the frequency range from 100 kHz to 0.1 Hz; Figure S9: the corresponding elements mapping results of (a) MnO₂, (b) PANI, and (c) Fc-MS functionalized samples to show their homogeneous distribution; Figure S10: the cycling performance of Fc-MS, MnO₂, and PANI at a scan rate of 100 mV s⁻¹; Figure S11: the impedance spectra of MnO₂ (a) and PANI (b) before and after long-term cycling at OCP in the frequency range from 100 kHz to 0.1 Hz; Figure S12: the CV curves of Fc-MS (a), MnO₂ (b), and PANI (c) at scan rates ranging from 20 mV s⁻¹ to 200 mV s⁻¹; Figure S13: the surface-controlled current ratio of G-240 at a scan rate of 10 mV s⁻¹; Figure S14: the cycling performance of the PANI symmetrical device; Table S1 the impedance fitting parameters of the graphite, G-240, and G-240 before activation.

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